Thermodynamic data at 298.15 K

Substance	State	$\Delta H_f^\ominus$ / kJ mol^{-1}	$\Delta G_f^\ominus$ / kJ mol^{-1}	$S^\ominus$ / J K^{-1} mol^{-1}
Ag	s	0	0	42.6
Ag_2O	s	−31.0	−11.2	121.3
CH_3OH	g	−200.7	−162.0	239.8
C_2H_4	g	52.3	68.2	219.6
H_2O	l	−285.8	−237.1	69.9
H_2O	g	−241.8	−228.5	188.8
OH^-	aq	−230.0	−157.2	−10.8
Zn	s	0	0	41.6
ZnO	s	−348.3	−318.3	43.6

Standard electrode potentials at 298.15 K

Electrode reaction	$E^\ominus$/V
$H^+(aq) + e = \frac{1}{2}H_2(g)$	0.00
$\frac{1}{2}O_2(g) + 2H^+(aq) + 2e = H_2O(l)$	+1.23
$Fe^{2+}(aq) + 2e = Fe(s)$	−0.46
$Zn^{2+}(aq) + 2e = Zn(s)$	−0.76