

**MARK SCHEME for the May/June 2011 question paper  
for the guidance of teachers**

**5070 CHEMISTRY**

**5070/21**

Paper 2 (Theory), maximum raw mark 75

This mark scheme is published as an aid to teachers and candidates, to indicate the requirements of the examination. It shows the basis on which Examiners were instructed to award marks. It does not indicate the details of the discussions that took place at an Examiners' meeting before marking began, which would have considered the acceptability of alternative answers.

Mark schemes must be read in conjunction with the question papers and the report on the examination.

- Cambridge will not enter into discussions or correspondence in connection with these mark schemes.

Cambridge is publishing the mark schemes for the May/June 2011 question papers for most IGCSE, GCE Advanced Level and Advanced Subsidiary Level syllabuses and some Ordinary Level syllabuses.



<b>Page 2</b>	<b>Mark Scheme: Teachers' version</b>	<b>Syllabus</b>	<b>Paper</b>
	<b>GCE O LEVEL – May/June 2011</b>	<b>5070</b>	<b>21</b>

**A1 Allow** correct name but formula takes precedence

(a)  $V_2O_5$  (1) [1]

(b)  $ZnSO_4$  (1) [1]

(c)  $AgI$  (1) [1]

(d)  $CF_3Cl_3$  (1) [1]

(e)  $(NH_4)_2SO_4 / ZnSO_4$  (1) [1]

(f)  $CH_4$  (1) [1]

(g)  $(NH_4)_2SO_4$  (1) [1]

**[Total: 7]**

**A2 (a)** sulfur dioxide (1)

**Allow**  $SO_2$

[1]

(b) copper(II) sulfate (1)

**Allow**  $CuSO_4$

[1]

(c)  $H^+ + OH^- \rightarrow H_2O$  (1)

**Ignore** state symbols

[1]

(d) (i) Copper(II) hydroxide (1)

**Allow**  $Cu(OH)_2$

[1]

(ii)  $Cu^{2+}(aq) + 2OH^-(aq) \rightarrow Cu(OH)_2(s)$

Balanced equation (1)

Correct state symbols (1)

[2]

(e) Mol ratio  $Cu:O = \frac{79.9}{64} : \frac{20.1}{16} / 1.25 : 1.26$  (1)

$CuO$  (1)

[2]

**[Total: 8]**

Page 3	Mark Scheme: Teachers' version	Syllabus	Paper
	GCE O LEVEL – May/June 2011	5070	21

- A3 (a) (i)** same number of electrons / same number of protons / same electronic arrangement of electrons / both have 92 electrons / both have 92 protons (1) [1]
- (ii)** different number of neutrons / uranium-238 has three more neutrons (1) [1]
- (b) (i)**  $\text{UO}_2 + 4\text{HF} \rightarrow \text{UF}_4 + 2\text{H}_2\text{O}$  (1) [1]
- (ii)**  $\text{UF}_4 + 2\text{Mg} \rightarrow \text{U} + 2\text{MgF}_2$  (1) [1]
- (iii)** reaction involving gain of electrons / reaction involving decrease in oxidation number (1)  
**Allow** a reaction involving the loss of oxygen / gain of hydrogen [1]
- (iv)**  $M_r$  of  $\text{UO}_2 = 270$  (1)  
Moles of  $\text{UO}_2 = 3704$  (1) **Allow** ecf from wrong  $M_r$   
Mass of uranium = 0.881 tonnes (1) **Allow** ecf from wrong moles  
Correct answer scores **all three** marks
- OR**  
Alternative approach using percentage composition  
 $M_r$  of  $\text{UO}_2 = 270$  (1)  
% of U = 88.1% (1) **Allow** ecf from wrong  $M_r$   
Mass of uranium = 0.881 tonnes (1) **Allow** ecf from wrong percentage [3]
- (c)** between magnesium and copper (1) [1]

**[Total: 9]**

Page 4	Mark Scheme: Teachers' version	Syllabus	Paper
	GCE O LEVEL – May/June 2011	5070	21

- A4 (a)** All covalent bond pairs shown (1)  
Rest of structure correct (1)  
**Ignore** inner shell electrons of oxygen [2]
- (b)** Must be a **comparison** in both marking points  
Particles in a gas are moving faster than particles in a liquid (1)  
Particles in a gas are further apart than those in a liquid (1) [2]
- (c)** Particles in pure hydrogen peroxide are more crowded / closer together / more particles per unit volume / particles are more concentrated (1)  
So more collisions per second / increased collision frequency / collisions more often / more chance of collision / collisions more likely (1) [2]
- (d) (i)**  $\text{Fe}^{2+} \rightarrow \text{Fe}^{3+} + \text{e}^-$  (1)  
**Allow**  $\text{Fe}^{2+} - \text{e}^- \rightarrow \text{Fe}^{3+}$   
**Allow** e instead of  $\text{e}^-$  [1]
- (ii)** Add sodium hydroxide (solution) / (aqueous) Ammonia / add (aqueous) hydroxide ions (1)  
Should be a brown-rust ppt (1) [2]
- (e)** (Colour change of  $\text{KMnO}_4$  shows) it is a reducing agent / it can be oxidised (1)  
(Colour change of KI shows) it is an oxidising agent / it can be reduced (1) [2]
- [Total: 11]**

- A5 (a)** 78–79 % (1) [1]
- (b)** Fractional distillation (1)  
of liquid air / liquefy air (1)  
because (the components of air have) different boiling points (1) [3]
- (c)** Idea that carbon cycle involves photosynthesis and respiration (1)  
Photosynthesis decreases carbon dioxide and increases oxygen / green plants change carbon dioxide into oxygen (1)  
**And**  
**any two from**  
Respiration increases carbon dioxide and decreases oxygen (1)  
Combustion increases carbon dioxide and decreases oxygen (1)  
Decomposition (of living things) increases carbon dioxide (1) [4]
- (d)** Used in flue-gas desulfurisation / removal of sulfur dioxide from gaseous emissions of power station / absorbs the sulfur dioxide / neutralises (acidic) sulfur dioxide (1)  
Added to lakes to neutralise acidic water (1) [2]
- [Total: 10]**

Page 5	Mark Scheme: Teachers' version	Syllabus	Paper
	GCE O LEVEL – May/June 2011	5070	21

- B6 (a)** Calcium nitrate solution contains ions / AW (1)  
 Pentane only contains molecules / pentane is a covalent compound / pentane does not contain ions (1) [2]
- (b)** Sodium and chlorine (1)  
**Allow** Na and  $\text{Cl}_2$  [1]
- (c)** Hydrogen, chlorine (and sodium hydroxide) (1)  
**Allow**  $\text{H}_2$ ,  $\text{Cl}_2$  (and NaOH) [1]
- (d)** Electrolyte is aluminium oxide (dissolved in cryolite) / alumina (1)  
 Graphite electrodes / Carbon electrodes (1) [2]
- (e) (i)** Gets plated with copper (1)  
 $\text{Cu}^{2+} + 2\text{e}^- \rightarrow \text{Cu}$  (1) [2]
- (ii)** 1.21 (g) [1]
- (iii)** 1.75 (g) [1]
- [Total: 10]**

- B7 (a)** Propanol / propan-1-ol / propan-2-ol (1) [1]
- (b)**  $\text{CH}_3\text{CH}_2\text{CH}_2\text{CH}_2\text{OH}$  /  $\text{CH}_3\text{CH}_2\text{CHOHCH}_3$  (1)  
 Only contains (C—C) single bonds (1)  
**Allow** there are no (carbon-carbon) double bonds [2]
- (c)**  $\text{C}_7\text{H}_{16}\text{O}$  (1)  
**Allow**  $\text{C}_7\text{H}_{15}\text{OH}$  [1]
- (d) (i)**  $\text{CH}_3\text{COOC}_2\text{H}_5$  (1) [1]
- (ii)** Solvent (1)  
**Allow** flavouring / perfume [1]
- (e)**  $\text{C}_6\text{H}_{12}\text{O}_6 \rightarrow 2\text{C}_2\text{H}_5\text{OH} + 2\text{CO}_2$  (1)  
 Use of yeast (1)  
 Any temperature or range of temperature within 20–40 °C / absence of oxygen / anaerobic conditions / presence of water / Fractional distillation (to separate ethanol) (1)  
**Ignore** incorrect reactants this has been assessed by the equation [3]
- (f)** Ethene /  $\text{C}_2\text{H}_4$  (1) [1]

**[Total: 10]**

Page 6	Mark Scheme: Teachers' version	Syllabus	Paper
	GCE O LEVEL – May/June 2011	5070	21

- B8 (a) (i)** Position of equilibrium moves to the right (1)  
**Allow** make more CH<sub>3</sub>COOH  
 Because the reaction is exothermic / to release energy (1) This mark is dependant on the position of equilibrium moves to the right [2]
- (ii)** Reaction is faster / activation energy is very high (1) [1]
- (b)** Labelled products to the right and below reactants (1)  
 Correct labelled activation energy for the forward reaction (1)  
**Allow** double headed arrow head / arrow without any heads  
**Not** arrow in wrong direction  
 Correct labelled enthalpy change (1)  
**Not** arrow in wrong direction / double headed arrow  
**Note** – arrows do not have to start exactly at reactant level and finish exactly at product or maximum of curve  
**Maximum** of **two** marks for an error carried forward for a reaction that is endothermic i.e. enthalpy change mark and activation energy [3]
- (c)** Lowers the activation energy (1)  
**Allow** more effective collisions / more successful collisions [1]
- (d)** Maximum moles that can be made is 10 / limiting reactant is the carbon monoxide (1)  
 98% (1) [2]
- (e)** CH<sub>3</sub>CO<sub>2</sub>NH<sub>4</sub> (1) [1]
- [Total: 10]**

- B9 (a)** Only partially dissociates / does not completely ionise (1) [1]
- (b)** Use universal indicator (1)  
 Idea that the different colours indicate different pH values / match colour against a colour chart (1)  
**Allow** this mark even for an incorrect indicator [2]
- (c)** Moles of sulfamic acid =  $\frac{0.105}{97} / 0.00107$  (1)  
 Moles of KOH =  $\frac{10.8}{1000} \times 0.100 / 0.00108$  (1)  
 so reacts with one mole (1) [3]
- (d) (i)**  $\text{Mg} + 2\text{SO}_3\text{NH}_3 \rightarrow \text{Mg}(\text{SO}_3\text{NH}_2)_2 + \text{H}_2$  (1) [1]
- (ii)**  $\text{CaCO}_3 + 2\text{SO}_3\text{NH}_3 \rightarrow \text{Ca}(\text{SO}_3\text{NH}_2)_2 + \text{H}_2\text{O} + \text{CO}_2$  (1)  
 Forms carbon dioxide / bubbles (1)  
**Allow** carbon dioxide from the equation [2]
- (e)** Nitrogen (1) [1]
- [Total: 10]**