

UNIVERSITY OF CAMBRIDGE INTERNATIONAL EXAMINATIONS International General Certificate of Secondary Education

CO-ORDINATED SCIENCES

0654/01

Paper 1 Multiple Choice

May/June 2009

45 minutes

Additional Materials: Multiple Choice Answer Sheet

Soft clean eraser

Soft pencil (type B or HB is recommended)

READ THESE INSTRUCTIONS FIRST

Write in soft pencil.

Do not use staples, paper clips, highlighters, glue or correction fluid.

Write your name, Centre number and candidate number on the Answer Sheet in the spaces provided unless this has been done for you.

There are **forty** questions on this paper. Answer **all** questions. For each question there are four possible answers **A**, **B**, **C** and **D**.

Choose the one you consider correct and record your choice in soft pencil on the separate Answer Sheet.

Read the instructions on the Answer Sheet very carefully.

Each correct answer will score one mark. A mark will not be deducted for a wrong answer.

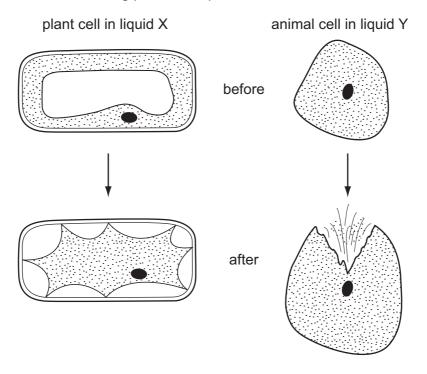
Any rough working should be done in this booklet.

A copy of the Periodic Table is printed on page 20.

This document consists of **18** printed pages and **2** blank pages.



- 1 Which characteristic is shown by members of the same species?
 - A They all live in the same place.
 - **B** They are all identical in appearance.
 - **C** They breed with each other to produce fertile offspring.
 - **D** They cannot form clones.
- 2 The diagram shows a plant cell before and after being placed in liquid X for 30 minutes, and an animal cell before and after being placed in liquid Y for 30 minutes.

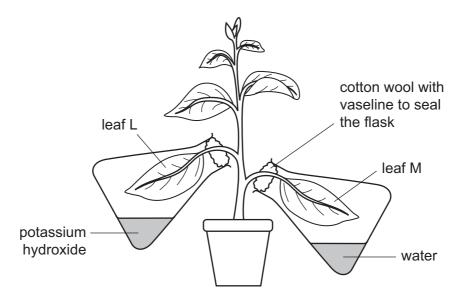


What describes liquids X and Y?

	Х	Y		
Α	concentrated solution	concentrated solution		
В	concentrated solution	pure water		
С	pure water	concentrated solution		
D	pure water	pure water		

3 The diagram shows an experiment to investigate photosynthesis.

Potassium hydroxide absorbs carbon dioxide.

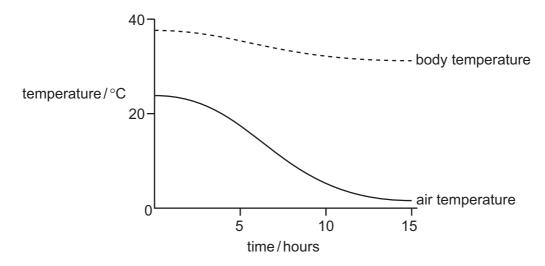


After standing in sunlight for 10 hours, leaf L contained no starch but leaf M contained a lot of starch.

What does this show?

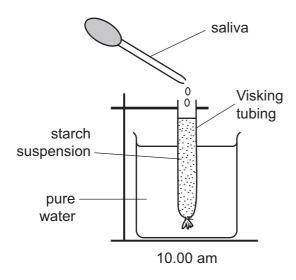
- A A leaf cannot make starch in a sealed flask.
- **B** A leaf cannot make starch without carbon dioxide.
- **C** A leaf cannot make starch without light.
- **D** A leaf cannot make starch without oxygen.
- **4** Between which structures are the pleural membranes found?
 - A bronchi and bronchioles
 - B diaphragm and ribs
 - C larynx and trachea
 - **D** lungs and intercostal muscles

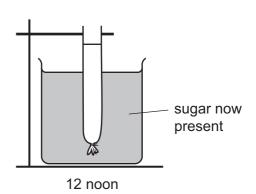
5 The graph shows how a person's body temperature changes with changing air temperature.



Which process provides the energy for maintaining the body temperature as shown in the graph?

- A breathing
- **B** digestion
- **C** excretion
- **D** respiration
- 6 The diagram shows a piece of Visking tubing (partially permeable) containing starch suspension, held in a beaker of pure water. Saliva, containing the enzyme amylase, is added to the starch and left for two hours.



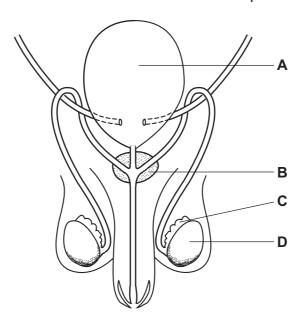


What does the experiment show?

- **A** Amylase is a solvent for starch.
- **B** Saliva passes through the Visking tubing.
- **C** Starch can be changed to sugar.
- **D** Starch is soluble in pure water.

- 7 Which substance is needed in the diet so that red blood cells can carry oxygen?
 - A calcium
 - **B** iron
 - C vitamin C
 - D vitamin D
- 8 What is most likely to happen after a person eats a meal high in protein?
 - A The amount of water in the blood would decrease.
 - **B** The concentration of urea in the urine would increase.
 - **C** The level of insulin in the blood would increase.
 - **D** The temperature of the body would decrease.
- **9** In which part of a seed is the micropyle found?
 - A cotyledon
 - **B** plumule
 - C radicle
 - **D** testa
- **10** The diagram shows the male reproductive system.

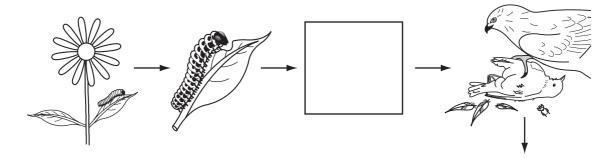
In which structure are the hormones that control adolescence produced?



11 Black coat colour in mice is dominant to white coat colour. A pure-bred black mouse mates with a white mouse.

What colour are the offspring?

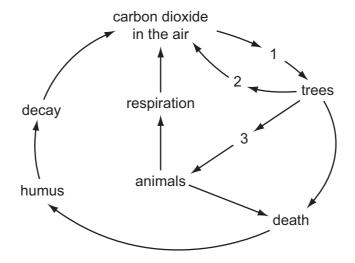
- A black only
- B black and white
- **C** grey
- **D** white only
- **12** The diagram shows a food chain.



What does the empty box represent?

- A consumer
- **B** decomposer
- C photosynthesis
- **D** producer

13 The diagram shows part of the carbon cycle in a forest. The numbers represent different processes.



Which of these processes is reduced in rate as a result of deforestation?

- A 1 only
- B 1 and 2 only
- C 2 and 3 only
- **D** 1, 2 and 3

14 The element phosphorus burns in air, as shown.

$$4P + 5O_2 \rightarrow P_4O_{10}$$

What does the formula P₄O₁₀ show?

- A a mixture of atoms of two elements
- **B** a mixture of molecules of two elements
- C a molecule of a compound
- D an atom of a compound

15 Which types of oxide are formed by magnesium and sulfur?

	magnesium	sulfur		
A acidic		acidic		
B acidic		basic		
С	basic	acidic		
D	basic	basic		

- 16 Which process produces molecules with longer chains?
 - A combustion of hydrocarbon
 - **B** cracking
 - C fractional distillation of crude oil
 - **D** polymerisation
- **17** Proteins consist of long chains of1..... molecules and always contain the elements carbon, hydrogen, nitrogen and2.......

Which words correctly complete gaps 1 and 2?

	1	2		
Α	amino acid	oxygen		
В	amino acid	sulfur		
С	glucose	oxygen		
D	glucose	sulfur		

18 The equation for the extraction of a metal from its oxide can be written as shown.

metal oxide + carbon → metal + carbon dioxide

Which statements about this reaction are correct?

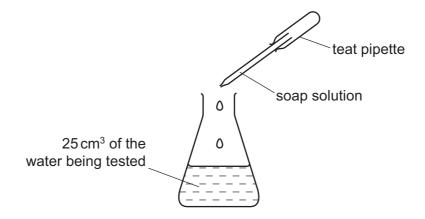
	the metal oxide is reduced	the carbon is oxidised		
Α	✓	✓		
В	✓	X		
С	×	✓		
D	X	X		

19 Aqueous sodium chloride is electrolysed on a large scale.

Which three substances are manufactured in this way?

- A acid, chlorine and hydrogen
- **B** acid, chlorine and oxygen
- C alkali, chlorine and hydrogen
- D alkali, chlorine and oxygen

20 The diagram shows an experiment to test the hardness of separate samples of distilled water, tap water and boiled tap water.



Soap solution is added, shaking after each drop, until a lather is formed.

Which results could be correct?

	number of drops of soap solution used							
	distilled water tap water boiled tap water							
Α	2	5	10					
В	2	10	5					
С	5	10	2					
D	10	5	2					

21 The waste from a factory is acidic. The factory treats the waste with lime.

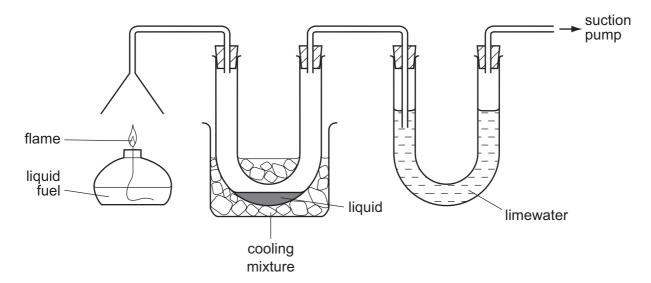
Which pH change takes place?

	pH of waste	pH of treated waste				
Α	6	5				
В	6	7				
С	8	7				
D	8	9				

- 22 From which carbonate is lime manufactured?
 - A calcium carbonate
 - **B** lead(II) carbonate
 - C magnesium carbonate
 - **D** zinc carbonate

23	Wh	hich metal is used with aqueous sodium hydroxide to test for nitrate ions in solution?						
	Α	aluminium						
	В	copper						
	С	magnesium						
	D	tin						
24	Wh	at is an analgesic?						
	A	an alloy						
	В	an antacid						
	С	a monomer						
	D	a painkiller						
25	A s	ample of clay is stirred in a beaker of water.						
	Wh	/hen light is shone through the beaker, the light is scattered.						
	Wh	/hat does the experiment show?						
	Α	An emulsion has been formed.						
	В	Clay in water forms a colloid.						
	С	Clay in water forms a gel.						
	D	Water dissolves clay particles.						

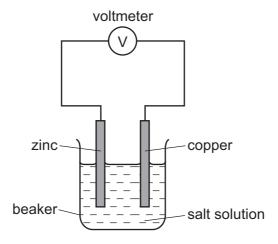
26 A liquid fuel is burned using the following apparatus.



What is being tested for in the gases produced by the burning fuel?

- A carbon monoxide and carbon dioxide
- B carbon monoxide and water
- C carbon dioxide and water
- **D** carbon dioxide and sulfur dioxide

27 The diagram shows a simple cell.



Which change would increase the reading on the voltmeter?

- A adding more solution
- **B** replacing the zinc with magnesium
- C using a larger beaker
- D using a larger piece of zinc

28 A car travels 100 km. The highest speed of the car is 90 km/h, and the lowest speed is 30 km/h. The journey takes two hours.

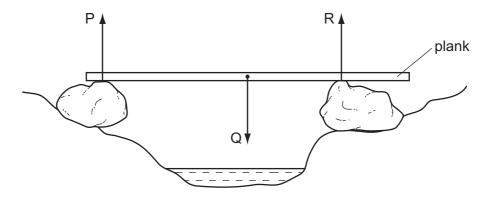
What is the average speed for the journey?

- **A** 30 km/h
- **B** 50 km/h
- **C** 60 km/h
- **D** 90 km/h

29 Which items of apparatus are required to determine the density of a liquid?

- A balance and measuring cylinder
- **B** balance and thermometer
- C metre rule and measuring cylinder
- **D** metre rule and thermometer

30 A wooden plank rests in equilibrium on two boulders on opposite sides of a narrow stream. Three forces of size P, Q and R act on the plank.



How are the sizes of the forces related?

- A P+Q=R
- $\mathbf{B} + \mathbf{R} = \mathbf{Q}$
- **C** P = Q = R
- $\mathbf{D} = \mathbf{Q} + \mathbf{R}$

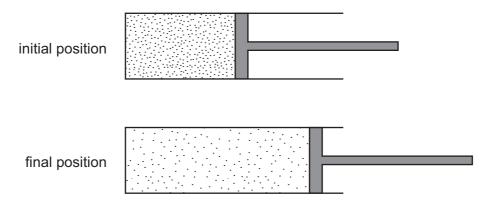
31 Electricity can be obtained from different energy resources.

Which energy resource is used to obtain electricity without producing heat to boil water?

- A coal
- **B** gas
- C hydroelectric
- **D** nuclear

32 A piston traps a certain mass of gas inside a cylinder. Initially the piston is halfway along the length of the cylinder.

The piston is now moved towards the open end of the cylinder. The temperature of the gas remains constant.

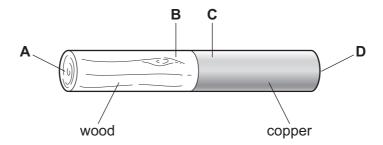


How are the density and the pressure of the gas affected by moving the piston?

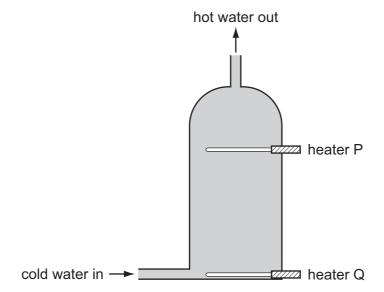
	density	pressure		
Α	decreases	decreases		
B decreases		unchanged		
С	increases	decreases		
D	increases	unchanged		

33 A rod is made up of copper and wood joined together.

After the rod is heated at the join in the centre for about a minute, where would the lowest temperature be?



34 A hot water tank is fitted with two identical heaters P and Q. Heater P is two thirds of the way up the tank and heater Q is at the very bottom. The tank is full of cold water.

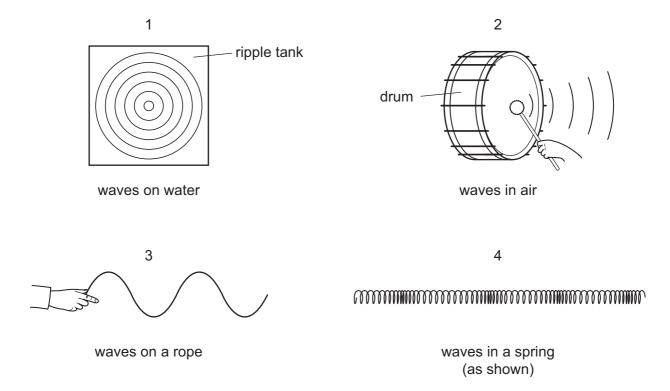


When only heater Q is switched on, it takes a very long time to heat the tank of water to the required temperature of 60 °C.

What happens to the tank of cold water if only heater P is switched on?

- A All the water reaches 60 °C in less time than before.
- **B** All the water reaches 60 °C in the same time as before.
- **C** The bottom two thirds of the water reaches 60 °C in two thirds of the original time
- **D** The top one third of the water reaches 60 °C in one third of the original time.

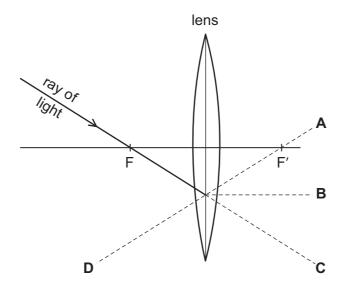
35 The diagrams show examples of wave motion.



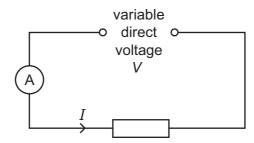
Which are longitudinal waves?

- A 1 only
- **B** 1, 2 and 4
- C 2 and 3 only
- **D** 2 and 4 only
- **36** The diagram shows the path of a ray of light passing through the principal focus F of a lens.

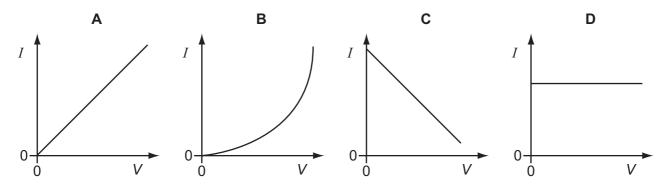
Which broken line shows the direction of the ray after it leaves the lens?



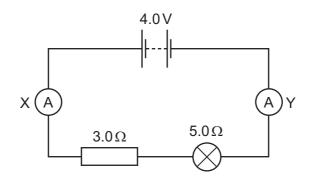
37 Using the circuit shown, the current I is found for various voltages V. The temperature of the resistor does not change.



Which graph shows the results obtained?



38 In the circuit shown, ammeter X reads 0.5 A.



What does ammeter Y read?

A 0

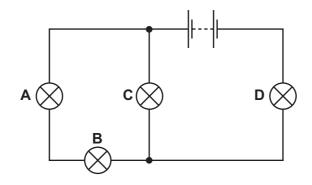
B 0.5 A

C 3.5 A

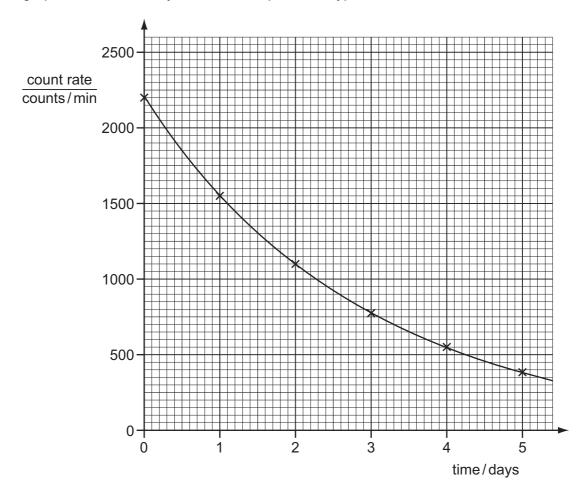
D 4.0 A

39 In the circuit below, one of the lamps breaks, causing all the other lamps to go out.

Which lamp breaks?



40 The graph shows the decay curve for one particular type of radioactive nuclide.



What is the half-life of this nuclide?

A 1.0 day

B 1.5 days

C 2.0 days

D 2.5 days

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DATA SHEET
The Periodic Table of the Elements

	0	Heium 2	Neon 10 Argon 18	84 Kr Krypton 36	131 Xe Xenon 54	Rn Radon		175 Lu Lutetium 71	Lr Lawrencium 103
	II/		19 Fluorine 9 35.5 C 1	80 Br Bromine 35	127 I lodine 53	At Astatine 85		173 Yb Ytterbium 70	Nobelium 102
	N		16 Oxygen 8 32 Sulfur 16	79 Se Selenium 34	128 Te Tellurium 52	Po Polonium 84		169 Tm Thullum 69	Md Mendelevium 101
	>		14 Nitrogen 7 31 Phosphorus 15	75 As Arsenic 33	Sb Antimony 51	209 Bi Bismuth		167 Er Erbium 68	Fm Fermium
	≥		12 Carbon 6 Silicon 14	73 Ge Germanium	Sn Tin	207 Pb Lead		165 Ho Holmium 67	ES Einsteinium 99
	=		11 B Boron 5 27 A1 Aluminium	70 Ga Gallium 31	115 In Indium 49	204 T 1 Thallium		162 Dy Dysprosium 66	Cf Californium 98
				65 Zn Zinc 30	Cadmium 48	201 Hg Mercury 80		159 Tb Terbium 65	BK Berkelium 97
				64 Copper 29	108 Ag Silver 47	197 Au Gold		157 Gd Gadolinium 64	Carium Ourium
Group				Nickel	106 Pd Palladium 46	195 Pt Platinum 78		152 Eu Europium 63	Am Americium 95
G				59 Co Cobalt 27	Rh Rhodium	192 I r Iridium 77		Sm Samarium 62	Pu Plutonium 94
		T Hydrogen		56 Fe Iron	Ruthenium 44	190 Os Osmium 76		Pm Promethium 61	Neptunium
				Manganese	Tc Technetium 43	186 Re Rhenium 75		Neodymium 60	238 U Uranium 92
				52 Cr Chromium 24	96 Mo Molybdenum 42	184 W Tungsten 74		Pr Praseodymium 59	Pa Protactinium 91
				51 V Vanadium 23	Niobium 41	181 Ta Tantalum		140 Cer ium 58	232 Th Thorium 90
				48 Ti Titanium 22	91 Zr Zirconium 40	178 Ha fnium * 72			mic mass nbol nic) number
				Scandium 21	89 X	139 La Lanthanum 57 ,	Actinium 89	d series series	a = relative atomic mass X = atomic symbol b = proton (atomic) number
	=		Be Beryllium 4 24 Mg Magnesium 12	40 Ca Calcium	Strontium 38	137 Ba Barium 56	226 Ra Radium 88	*58-71 Lanthanoid series	<i>a</i> × <i>a</i> ∠
	_		7 Lithium 3 23 Na Sodium 11	39 K Potassium	RB Rubidium 37	133 Csesium 55	Fr Francium 87	*58-71 L 190-103	Key

The volume of one mole of any gas is $24\,\mathrm{dm}^3$ at room temperature and pressure (r.t.p.).

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