



UNIVERSITY OF CAMBRIDGE INTERNATIONAL EXAMINATIONS  
International General Certificate of Secondary Education

CANDIDATE  
NAME

CENTRE  
NUMBER

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NUMBER

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**CO-ORDINATED SCIENCES**

**0654/02**

Paper 2 (Core)

**May/June 2008**

**2 hours**

Candidates answer on the Question Paper.

No Additional Materials are required.

**READ THESE INSTRUCTIONS FIRST**

Write your Centre number, candidate number and name on all the work you hand in.

Write in dark blue or black pen.

You may use a soft pencil for any diagrams, graphs, tables or rough working.

Do not use staples, paper clips, highlighters, glue or correction fluid.

**DO NOT WRITE IN ANY BARCODES.**

Answer **all** questions.

A copy of the Periodic Table is printed on page 24.

At the end of the examination, fasten all your work securely together.

The number of marks is given in brackets [ ] at the end of each question or part question.

For Examiner's Use	
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2	
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<b>Total</b>	

This document consists of **23** printed pages and **1** blank page.



1 Fig. 1.1 shows a section through a human eye.

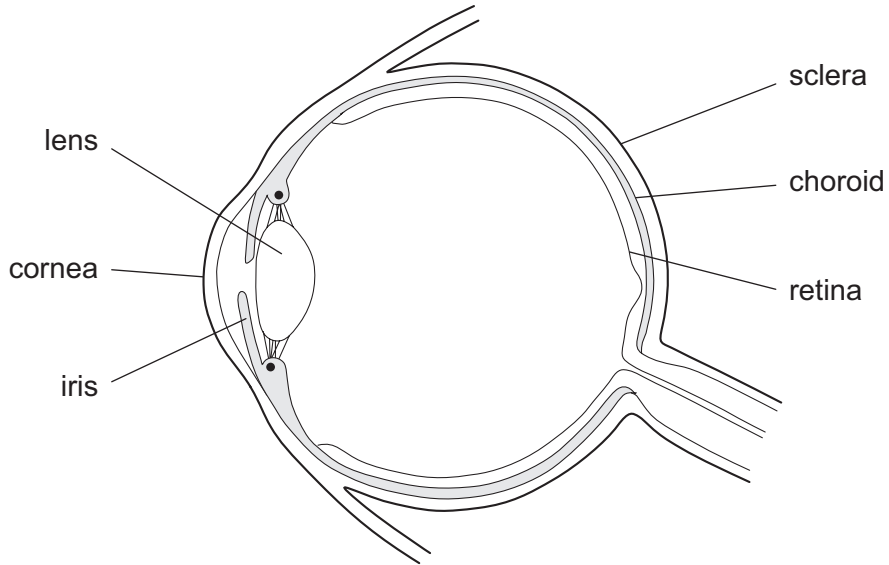


Fig. 1.1

(a) Using the labels on Fig. 1.1, list, in order, the parts of the eye through which light passes to reach the retina.

..... [1]

(b) Describe the function of the following parts of the eye.

(i) the lens

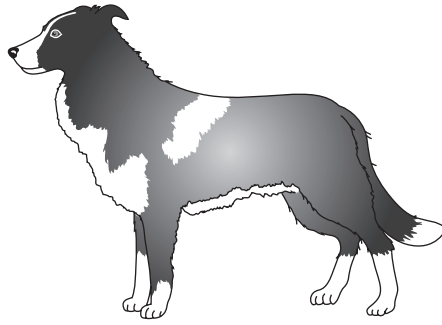
.....  
.....  
..... [2]

(ii) the retina

.....  
.....  
..... [2]

- (c) Collies are a breed of dog that have been bred to herd sheep and cattle. A recessive allele, **a**, in collies causes the choroid to develop abnormally. This can cause blindness.

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- (i) What is the phenotype of a collie with the genotype **aa**?

..... [1]

Breeders of collies try to make sure that none of the puppies that are born inherit this disease.

A collie breeder mates a male dog with the genotype **AA**, and a female dog with the genotype **Aa**.

- (ii) Complete the genetic diagram to explain whether any of their puppies will inherit the choroid disease.

parents	<b>AA</b>		<b>Aa</b>
gametes	all <b>A</b>		..... and .....
offspring genotypes	.....		
offspring phenotypes	..... [3]		

- 2 (a) The mass of a golf ball is 40 g.

Its volume is  $35 \text{ cm}^3$ .

Calculate the density of the golf ball.

State the formula that you use and show your working.

formula

working

.....  $\text{g/cm}^3$  [2]

- (b) A golfer hits the ball.

Calculate the momentum of the golf ball when it has a velocity of  $40 \text{ m/s}$ .

State the formula that you use and show your working.

formula

working

.....  $\text{kg m/s}$  [2]

(c) The golfer's bag of clubs has a mass of 6 kg.

- (i) Calculate the weight of the bag of clubs.  
Assume that the gravitational field strength on Earth is 10 N/kg.

..... N [1]

- (ii) Calculate the work done by the golfer when the bag is lifted 0.5 metres.

State the formula that you use and show your working.

formula

working

..... J [2]

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3 Fig. 3.1 shows some natural processes which occur on and under the Earth's surface.

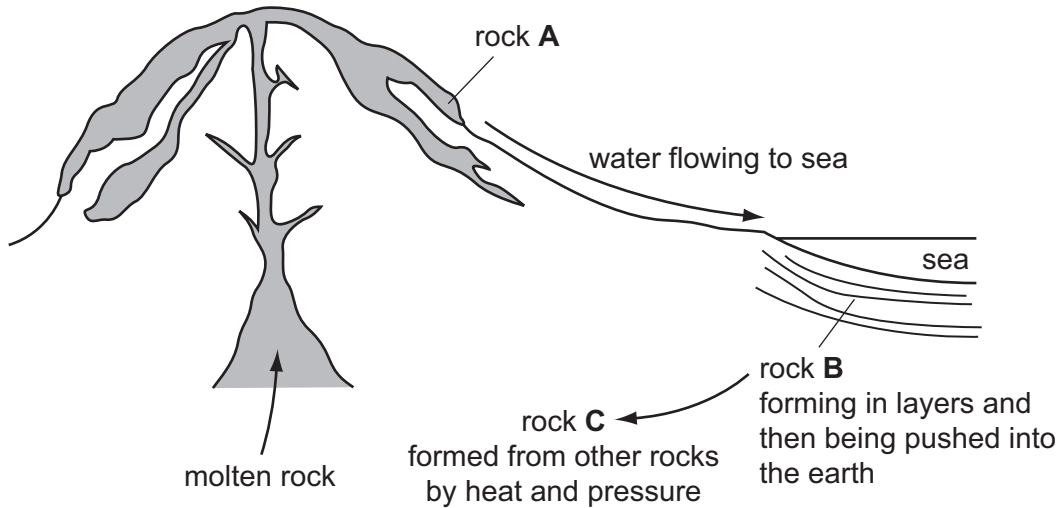


Fig. 3.1

(a) State which rock, **A**, **B** or **C**, was formed when a hot liquid cooled and changed into a solid.

..... [1]

(b) Rock **B** was formed when tiny pieces of solid were washed down into the sea by rivers and compressed. The tiny pieces of solid were produced from rock **A** whose surface had been damaged by weathering.

(i) What general name is given to rocks like rock **B**?

..... [1]

(ii) Describe **one** way in which the surface of rock **A** could have been weathered.

.....  
 .....  
 .....  
 ..... [2]

(iii) Underline the word in the list below which correctly names the type of weathering you have described in part (ii).

- biological**                      **chemical**                      **physical**

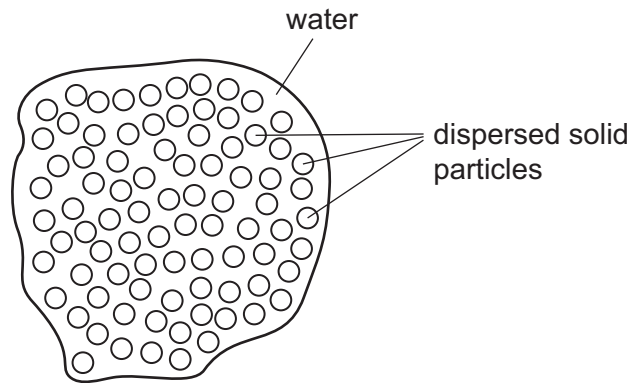
[1]

- (c) A sample of water flowing into the sea, as shown in Fig. 3.1, was taken to a laboratory for testing.

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A student observed a drop of the water under a microscope.

Fig. 3.2 shows a labelled diagram of what he saw.



**Fig. 3.2**

- (i) What **general** name is given to a mixture in which one substance is finely dispersed throughout another?

..... [1]

- (ii) The student stated that the mixture he was observing was an example of an *emulsion*.

Explain whether or not the student's statement was correct.

.....  
 .....  
 ..... [2]

- (iii) The student then added a few drops of acidified barium nitrate solution to some of the water. A white precipitate was formed.

What may be concluded about the water sample from this result?

.....  
 ..... [1]

4 Fig. 4.1 shows a transverse section through a leaf.

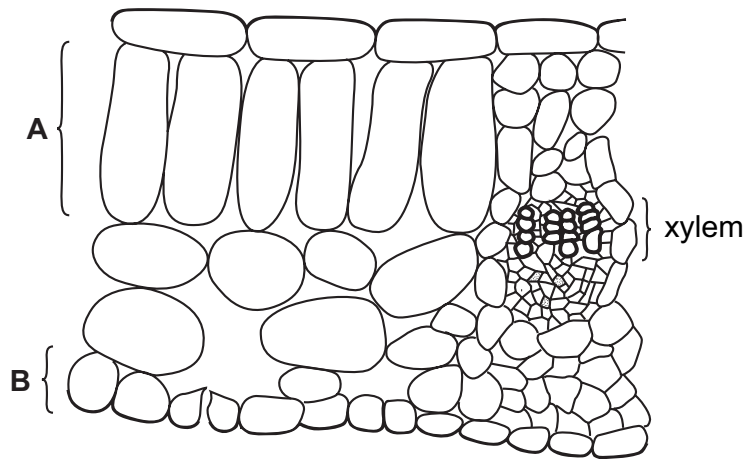


Fig. 4.1

(a) (i) Name the tissues labelled **A** and **B**.

**A** .....

**B** ..... [2]

(ii) State two ways in which a cell in tissue **A** differs from an animal cell.

1. ....

2. .... [2]

(iii) On Fig. 4.1, draw an arrow to show where carbon dioxide enters the leaf. [1]

(b) State two functions of xylem tissue in a leaf.

1. ....

2. .... [2]



- 5 (a) The graph in Fig. 5.1 shows the motion of a dolphin travelling through water.

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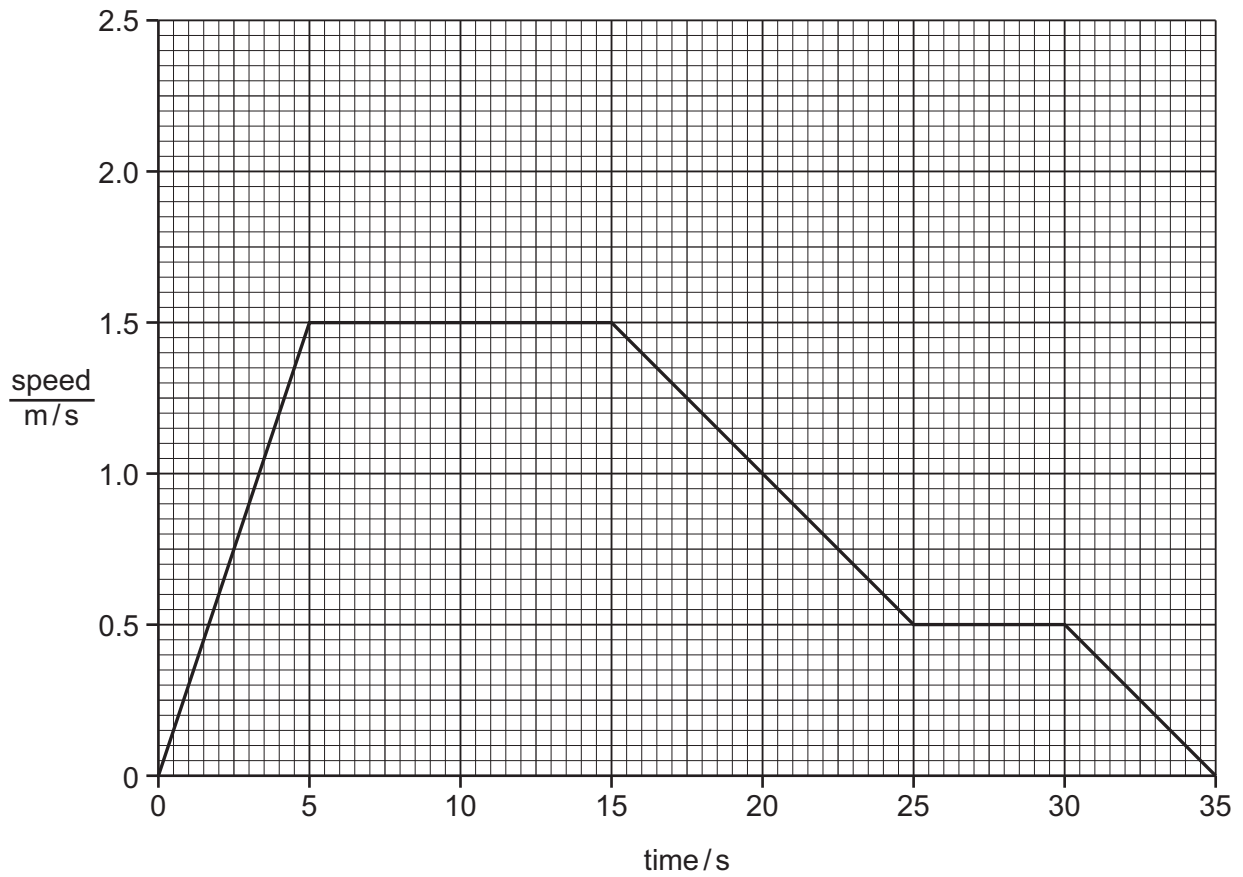


Fig. 5.1

- (i) On the graph, label with an **S** a period when the dolphin was moving at a constant speed. [1]

- (ii) Describe the motion of the dolphin between 0 s and 5 s.

..... [1]

(b) Table 5.1 shows the maximum and minimum frequencies of sounds heard by dolphins, humans and whales.

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**Table 5.1**

animal	maximum frequency / kHz	minimum frequency / Hz
dolphin	110	40
human	20	20
whale	1	2

(i) What is meant by the term *frequency*?

.....  
 ..... [1]

Which animal can hear

(ii) the greatest range of frequencies, ..... [1]

(iii) the sound with the highest pitch? ..... [1]

(c) A dolphin locates an object by emitting a pulse of high frequency sound.

The pulse takes 0.2 s to reach the object and return to the dolphin after reflection. The speed of the sound pulse in water is 1500 m/s.

Calculate the distance between the dolphin and the object.

State the formula that you use and show your working.

formula

working

..... m [3]

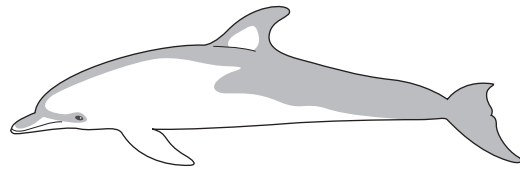
- (d) A man in a boat sees a dolphin under the water. Draw a ray of light on Fig. 5.2 to show how light travels from the dolphin's head to the man's eye.

*For  
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Use*



air

water



**Fig. 5.2**

[3]

6 Fig. 6.1 shows diagrams of some atoms of elements in Group I of the Periodic Table.

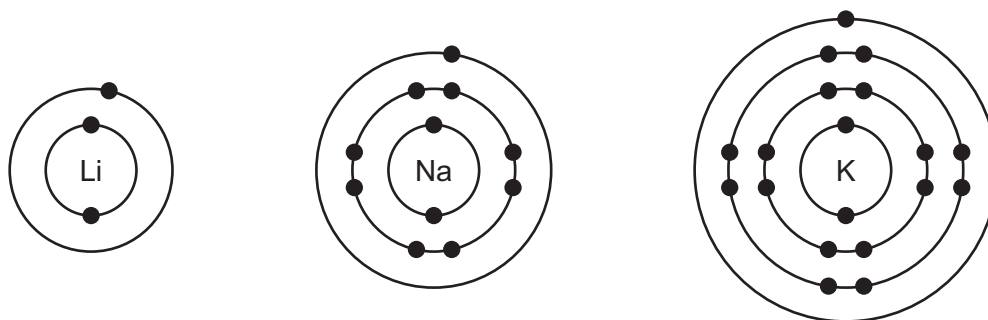


Fig. 6.1

(a) (i) Describe briefly two differences in the properties of lithium and potassium.

1. ....

.....

2. ....

..... [2]

(ii) When sodium reacts with water, sodium atoms change into sodium ions. Draw a diagram of a sodium ion showing how all the electrons are arranged.

[1]

(iii) Rubidium is another metal in Group I. Explain why a rubidium ion has a single positive electrical charge.

.....

..... [1]

For  
Examiner's  
Use

(b) Fig. 6.2 shows apparatus a student used to investigate electrochemical cells.

For  
Examiner's  
Use

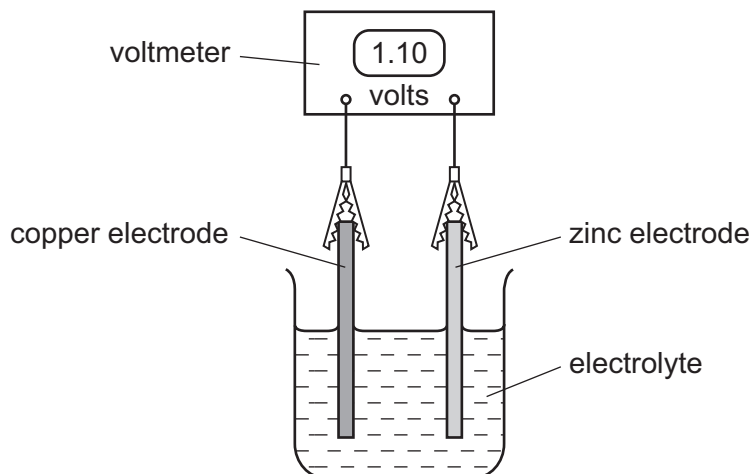


Fig. 6.2

Table 6.1 shows some properties of substances which the student thought might be suitable to produce the electrolyte.

Table 6.1

substance	type of bonding	solubility in water
calcium carbonate	ionic	insoluble
glucose	covalent	soluble
magnesium sulphate	ionic	soluble
silicon dioxide	covalent	insoluble

(i) State and explain which **one** of the substances in Table 6.1 is suitable for making the electrolyte.

.....  
 .....  
 ..... [2]

(ii) Describe briefly what change the student could make to the apparatus in Fig. 6.2 in order to obtain a different value of the cell voltage.

.....  
 ..... [1]

7 A farmer has grown corn (maize) in the same field for several years.

He measured the concentration of nitrate in the soil in 2001 and in 2003.

Fig. 7.1 shows the results.

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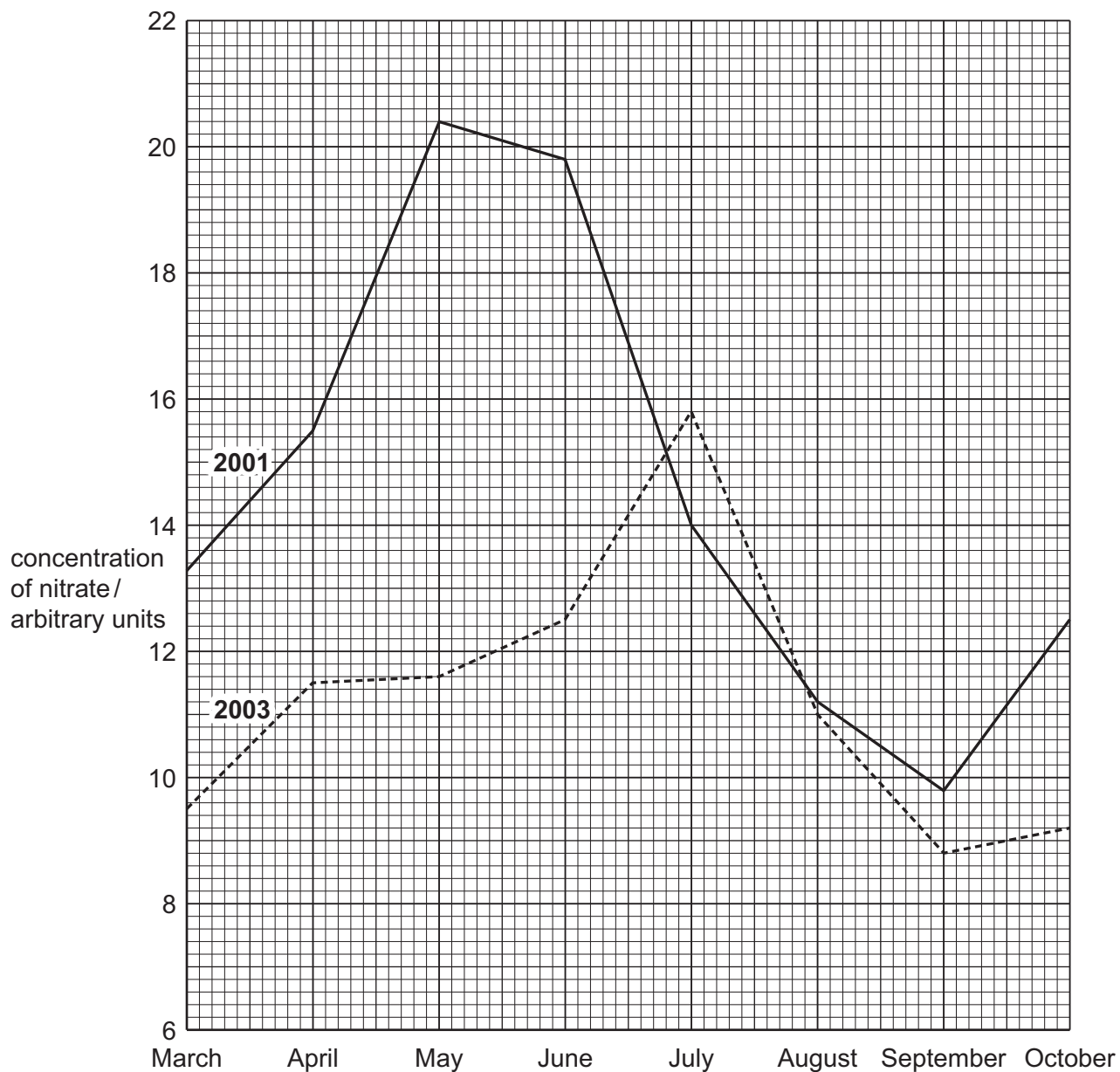


Fig. 7.1

(a) (i) In 2001, in which month was the concentration of nitrate in the soil the highest?

..... [1]

(ii) Describe two ways in which the nitrate concentration in the soil in 2003 was different from the concentration in 2001.

- 1. .... [2]
- 2. .... [2]

(b) The farmer was worried that the nitrate concentration in the field might be too low. He decided to try to increase it.

(i) Explain why increasing the nitrate concentration in the field might help the farmer.

..... [2]  
 .....  
 .....

(ii) Suggest how he could increase the nitrate concentration in the field.

..... [1]

(c) The farmer feeds the maize to cattle. He sells meat from the cattle for people to eat.

(i) Draw a food chain to show this information.

[1]

(ii) What do the arrows in your food chain represent?

..... [1]

(d) When the maize plants are harvested, their roots are left in the soil.

Describe how the carbon compounds in the roots will be turned into carbon dioxide and released into the air.

..... [2]  
 .....  
 .....

8 The bodywork of a car is usually made from steel.

(a) If part of the bodywork goes very rusty it is usually removed and replaced with plastic filler, before being painted.

A car mechanic can use a magnet to find out if parts of the bodywork of a car have been filled with plastic filler.

He tests three areas of a car by placing a magnet near the surface as shown in Fig. 8.1.

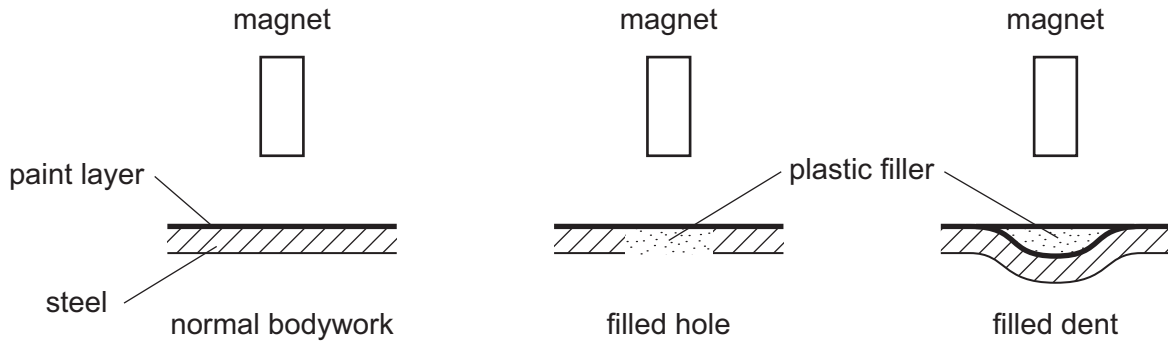


Fig. 8.1

(i) Complete the table.

area	effect on a magnet
normal bodywork	
filled hole	
filled dent	weakly attracted

[2]

(ii) What assumption have you made about the properties of plastic filler?

..... [1]

(iii) Would this method work if the bodywork was made of aluminium?

Explain your answer.  
..... [1]

(iv) Suggest why the bodywork of some cars is made from aluminium rather than steel.

..... [1]



(b) Exhaust gases from a car engine leave the car through a solid steel exhaust pipe.

Complete the sentences below about solids and gases.  
Use **only** the words **solid** or **gas**.

In a ....., the particles are closer together than in a .....

The forces of attraction between particles are stronger in a ..... than in a .....

When a ..... is heated it will eventually turn into a liquid.

In a ....., the particles can only vibrate and not move.

Heat energy will travel through a ..... by conduction.

Heat energy will **not** travel through a ..... by convection.

[4]

For  
Examiner's  
Use

9 Heat energy is obtained when hydrocarbon fuels are burned. Natural gas, methane, is an important hydrocarbon fuel. Natural gas is extracted from the Earth's crust.

(a) State why natural gas is called a *fossil fuel*.

.....  
 ..... [1]

(b) Explain why the burning of hydrocarbon fuels is thought to be causing significant changes to our environment.

.....  
 .....  
 ..... [2]

(c) Biogas is an alternative source of methane made from biodegradable materials. Biogas may be obtained from landfill sites and reaction vessels called digesters.

Some information about two sources of biogas are shown in Table 9.1.

**Table 9.1**

	% of substances in the biogas mixture	
	biogas from a digester	biogas from landfill
methane	60 – 70	45 – 55
carbon dioxide	30 – 40	30 – 40
nitrogen	less than 1	5 – 15

(i) Describe a chemical test which would show that biogas contains carbon dioxide.

.....  
 .....  
 ..... [2]

(ii) Use the information in Table 9.1 to suggest why 1.0 kg of biogas from a digester produces more heat energy when burned than 1.0 kg of biogas from a landfill site.

.....  
 .....  
 ..... [2]

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10 Enzymes are proteins that act as catalysts.

(a) Explain the meaning of the term *catalyst*.

.....  
 .....  
 ..... [2]

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(b) Amylase, protease and lipase are enzymes that digest food in the alimentary canal.

Draw lines to link each enzyme with the food type that it digests, and the molecules that digestion produces.

food digested	enzyme	molecules produced
fats	amylase	amino acids
proteins	protease	fatty acids and glycerol
starch	lipase	maltose (sugar)

[3]

(c) A good diet contains fibre. Fibre cannot be digested.

(i) Describe what happens to fibre that is eaten.

.....  
 .....  
 ..... [2]

(ii) Explain why fibre is an important part of a healthy diet.

.....  
 ..... [1]

(iii) Name **one** food that is a good source of fibre.

..... [1]

11 Starch, cellulose and proteins are compounds found in plants.

- (a) (i) State the **chemical symbols** of the three elements which are combined together in starch.

..... [1]

- (ii) The chemical bonds in starch are formed by atoms sharing pairs of electrons.

Name this type of chemical bonding.

..... [1]

- (b) Plants contain proteins, which are compounds containing nitrogen atoms. These atoms have been obtained from gaseous nitrogen in the air by nitrogen fixation.

- (i) Explain the meaning of the term *nitrogen fixation*.

.....  
 .....  
 ..... [2]

- (ii) When some types of protein are heated in sodium hydroxide solution, a gas is produced which turns damp red litmus paper blue.

Name this gas.

..... [1]

- (iii) A nitrogen atom has a *nucleon number* of 14.

Explain this statement.

.....  
 ..... [2]

- (c) State two important types of compound, other than those used for food, which may be extracted from plants.

1. ....

2. .... [2]

- 12 (a) The circuit in Fig. 12.1 was set up and the current measured by meters  $M_1$ ,  $M_2$ ,  $M_3$ ,  $M_4$  and  $M_5$ .

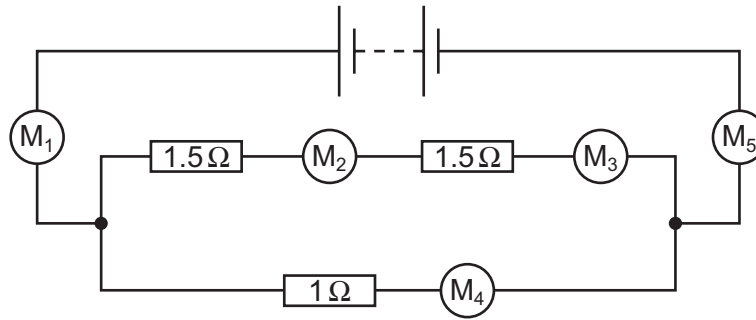


Fig. 12.1

- (i) What type of meter is  $M_1$ ?

..... [1]

- (ii) The readings on  $M_1$ ,  $M_3$ ,  $M_4$ , and  $M_5$  are shown in Table 12.1.

Complete the table for  $M_2$ .

Table 12.1

$M_1 = 4A$
$M_2 =$
$M_3 = 1A$
$M_4 = 3A$
$M_5 = 4A$

- (iii) Calculate the total resistance of the  $1.5 \Omega$  and  $1.5 \Omega$  resistors in series.

[1]

..... [1]

- (iv) The voltage across the  $1\ \Omega$  resistor is 3 V.

Use the formula

$$\text{power} = \text{voltage} \times \text{current}$$

to calculate the power consumed in the  $1\ \Omega$  resistor.

Show your working.

..... W [1]

- (b) The current flows through  $M_1$  for one minute.

Calculate the charge which has passed.

State the formula that you use and show your working.

formula

working

..... C [2]

**DATA SHEET**  
**The Periodic Table of the Elements**

		Group													
		I	II	III	IV	V	VI	VII	VIII	IX	X	0			
		1 <b>H</b> Hydrogen 1											4 <b>He</b> Helium 2		
7	9	3	4											19	20
<b>Li</b> Lithium	<b>Be</b> Beryllium											<b>F</b> Fluorine	<b>Ne</b> Neon		
23	24	11	12											35.5	40
<b>Na</b> Sodium	<b>Mg</b> Magnesium											<b>Cl</b> Chlorine	<b>Ar</b> Argon		
39	40	19	20											79	84
<b>K</b> Potassium	<b>Ca</b> Calcium											<b>Br</b> Bromine	<b>Kr</b> Krypton		
85	88	37	38											127	131
<b>Rb</b> Rubidium	<b>Sr</b> Strontium											<b>I</b> Iodine	<b>Xe</b> Xenon		
133	137	55	56											53	54
<b>Cs</b> Caesium	<b>Ba</b> Barium											<b>At</b> Astatine	<b>Rn</b> Radon		
226	227	87	88											85	86
<b>Fr</b> Francium	<b>Ra</b> Radium											<b>Po</b> Polonium	<b>Rn</b> Radon		
												81	82		
												65	66		
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												1	2		
												0	1		

		140	141	144	150	152	157	159	162	165	167	173	175
		58	59	60	61	62	63	64	65	66	68	70	71
		90	91	92	93	94	95	96	97	98	100	102	103
		232	238									169	175
		<b>Th</b> Thorium	<b>Pa</b> Protactinium	<b>U</b> Uranium	<b>Np</b> Neptunium	<b>Pu</b> Plutonium	<b>Am</b> Americium	<b>Cm</b> Curium	<b>Bk</b> Berkelium	<b>Cf</b> Californium	<b>Fm</b> Fermium	<b>No</b> Nobelium	<b>Lr</b> Lawrencium
		90	91	92	93	94	95	96	97	98	100	102	103

a	<b>X</b>	b	a	X	b
Key			a = relative atomic mass	X = atomic symbol	b = proton (atomic) number

\*58-71 Lanthanoid series  
†90-103 Actinoid series

The volume of one mole of any gas is 24 dm<sup>3</sup> at room temperature and pressure (r.t.p.).

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