

Cambridge International Examinations

Cambridge International General Certificate of Secondary Education

COMBINED SCIENCE 0653/23

Paper 2 Multiple Choice (Extended) October/November 2018

45 minutes

Additional Materials: Multiple Choice Answer Sheet

Soft clean eraser

Soft pencil (type B or HB is recommended)

READ THESE INSTRUCTIONS FIRST

Write in soft pencil.

Do not use staples, paper clips, glue or correction fluid.

Write your name, Centre number and candidate number on the Answer Sheet in the spaces provided unless this has been done for you.

DO NOT WRITE IN ANY BARCODES.

There are **forty** questions on this paper. Answer **all** questions. For each question there are four possible answers **A**, **B**, **C** and **D**.

Choose the **one** you consider correct and record your choice in **soft pencil** on the separate Answer Sheet.

Read the instructions on the Answer Sheet very carefully.

Each correct answer will score one mark. A mark will not be deducted for a wrong answer.

Any rough working should be done in this booklet.

A copy of the Periodic Table is printed on page 16.

Electronic calculators may be used.

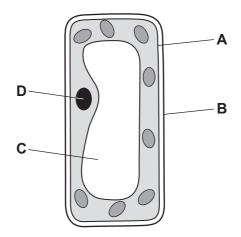


International Examinations

CAMBRIDGE

1 The diagram shows a typical plant cell.

Which part of the cell contains the genetic information of the cell?



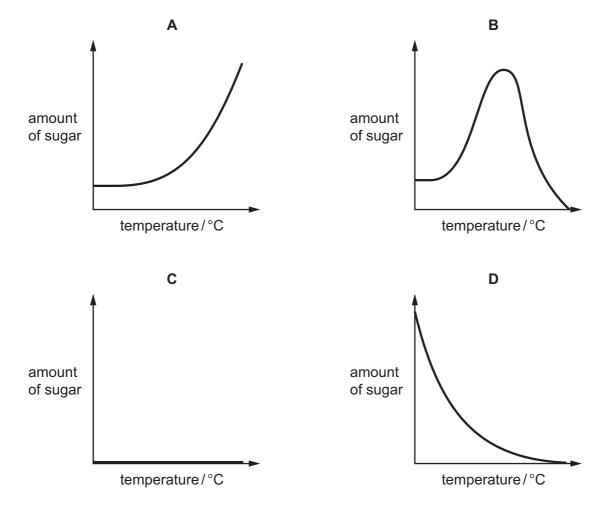
- 2 Which process depends on diffusion?
 - A circulation
 - **B** digestion
 - C gaseous exchange
 - **D** phagocytosis

3 A human enzyme breaks down starch into simple sugars.

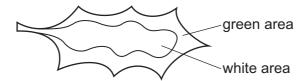
A solution of this human enzyme was heated to 90 °C for 30 minutes.

 $2\,\mathrm{cm}^3$ of this human enzyme solution was added to starch solution in several different test-tubes. The test-tubes were kept at different temperatures for 15 minutes.

Which graph shows the amount of sugar produced in the test-tubes?



4 The diagram shows a leaf that was tested for starch using iodine solution.



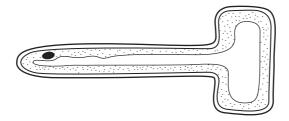
Which row shows the results for this leaf and explains the results?

	green area of leaf after test	white area of leaf after test	explanation
Α	blue-black	blue-black	chlorophyll is found in all parts of the leaf
В	blue-black	brown	chlorophyll is found in only part of the leaf
С	brown	brown	chlorophyll is found in all parts of the leaf
D	brown	blue-black	chlorophyll is found in only part of the leaf

5 A doctor advises a man to change his diet and reduce his intake of saturated fat and salt.

Which condition is the man most likely to be suffering from?

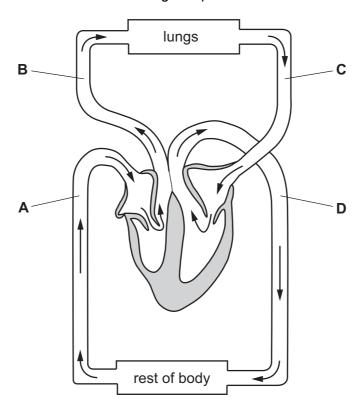
- A constipation
- B coronary heart disease
- **C** dental decay
- **D** starvation
- 6 The diagram shows a cross-section of a root hair cell.



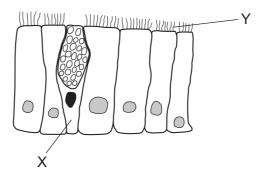
Which row describes the root hair cell and its function?

	animal cell or plant cell	function
Α	animal cell	water and glucose uptake
В	animal cell	water and ion uptake
С	plant cell	water and glucose uptake
D	plant cell	water and ion uptake

7 Which blood vessel contains blood at the highest pressure?



8 The diagram shows two different types of cell which line the trachea in the gas exchange system.

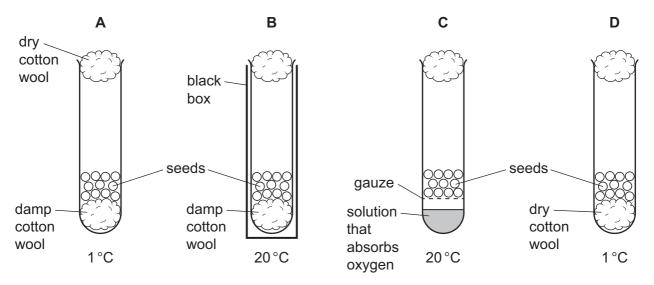


What is the role of X and Y?

	X	Υ
Α	produces mucus	traps pathogens
В	produces mucus	moves pathogens towards the mouth
С	moves pathogens towards the mouth	traps pathogens
D	moves pathogens towards the mouth	moves pathogens towards the mouth

- **9** Which statement about adrenaline is **not** correct?
 - A It decreases blood glucose concentration.
 - **B** It is carried by the blood.
 - **C** It is produced by a gland.
 - **D** The heart is one of its target organs.
- 10 In an investigation, four test-tubes containing seeds were set up as shown in the diagram.

After several days, which test-tube will contain the most germinated seeds?



11 During pregnancy in humans, gas exchange occurs between a mother and her fetus.

Where does this gas exchange occur?

- A amniotic fluid
- B amniotic sac
- C placenta
- **D** umbilical cord
- **12** What is the definition of a trophic level?
 - A It shows how an organism loses energy.
 - **B** It shows the position of an organism in a food chain.
 - **C** It shows the consumers of an organism.
 - **D** It shows the food eaten by an organism.

13 Which are possible harmful effects of deforestation?

	global warming	species extinction
Α	✓	✓
В	✓	x
С	X	✓
D	X	x

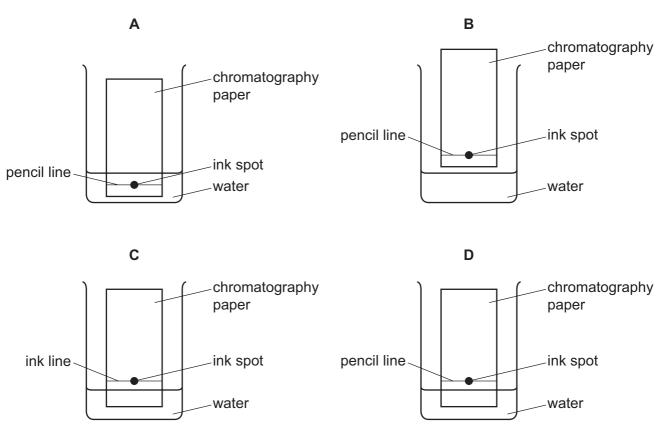
14 Sucrose is a covalent compound.

It is a solid at room temperature.

Which statement about sucrose is correct?

- **A** It is made of atoms that are close together and in continuous random motion.
- **B** It is made of atoms that are far apart and vibrating about a fixed point.
- **C** It is made of molecules that are close together and vibrating about a fixed point.
- **D** It is made of molecules that are far apart and in continuous random motion.

15 Which diagram shows how apparatus is used to separate the different colours in an ink?



16	Wh	ich is the electro	nic s	structure of a no	ble (gas?		
	Α	2,6	В	2,8	С	2,8,1	D	2,8,7
17	Wh	at is the formula	of n	itric acid?				
	Α	HC1	В	HNO ₃	С	NaOH	D	NH_3
18		ring the electrol vards the cathod		of molten pota	ssiu	ım chloride, whi	ich	particles in the electrolyte move
	Α	electrons						
	В	chloride ions						
	С	chlorine molecu	ules					
	D	potassium ions	;					
19	The	e temperature of	som	ne water is recor	ded			
	Soc	dium chloride is	disso	olved in the wate	er ar	d the temperatu	re of	f the solution is recorded.
		temperatur	re of	water = 20 °C				
		temperatur	re of	solution = 18°C	;			
	Wh	ich statement at	oout	the process is c	orre	ct?		
	A	It is endotherm	ic be	ecause chemical	ene	ergy is changed i	nto I	heat energy.
	В	It is endotherm	ic be	ecause heat ene	rgy i	s changed into o	chem	nical energy.
	С	It is exothermic	bec	ause chemical e	ener	gy is changed in	to he	eat energy.
	D	It is exothermic	bec	ause heat energ	gy is	changed into ch	emi	cal energy.
20		ostance X increar reaction.	ases	the rate of a cl	nem	ical reaction, bu	t it r	remains unchanged at the end of
	Wh	ich word describ	es s	ubstance X?				
	Α	catalyst						
	В	electrolyte						
	С	product						
	D	unreactive						

21 Iron oxide reacts with carbon monoxide.

The word equation for the reaction is:

iron oxide + carbon monoxide → iron + carbon dioxide

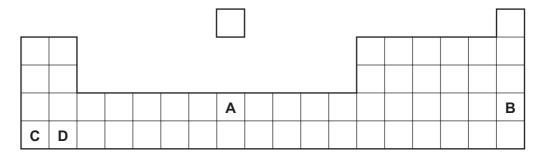
Which statement is **not** correct?

- A Carbon is neither oxidised nor reduced.
- **B** Carbon is oxidised.
- C Iron is reduced.
- **D** This is a redox reaction.
- 22 Rubidium is an element below potassium in Group I of the Periodic Table.

Which property of rubidium is **not** correct?

- A Rubidium is a soft metal.
- **B** Rubidium is less reactive than potassium.
- **C** Rubidium melts at a lower temperature than potassium.
- **D** Rubidium reacts with water forming hydrogen.
- 23 The positions of four elements are shown in the outline of the Periodic Table.

Which element has a high melting point and forms coloured compounds?



- 24 Which gas is used to fill weather balloons?
 - A argon
 - B carbon dioxide
 - C helium
 - **D** nitrogen

25 What is observed when magnesium ribbon is placed into aqueous copper sulfate?

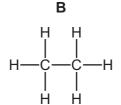
- **A** The blue solution gets darker and a brown solid appears.
- **B** The blue solution gets darker and no solid is visible.
- **C** The blue solution gets paler and a brown solid appears.
- **D** The blue solution gets paler and no solid is visible.

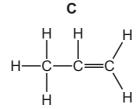
26 Gasoline is a hydrocarbon fuel obtained from petroleum.

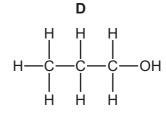
Which statement is correct?

- A Gasoline burns to form carbon dioxide and water.
- **B** Gasoline contains the elements carbon, hydrogen and oxygen.
- **C** Gasoline is used as a fuel in diesel engines.
- **D** The combustion of gasoline is an endothermic reaction.

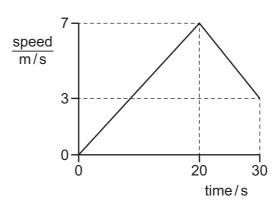
27 Which substance rapidly decolourises bromine?







28 The graph shows the motion of a cyclist over a period of 30 s.



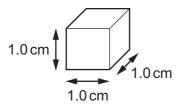
Which distance does she travel?

- **A** 90 m
- **B** 105 m
- **C** 115 m
- **D** 120 m

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29 A cube of aluminium has sides of length 1.0 cm.

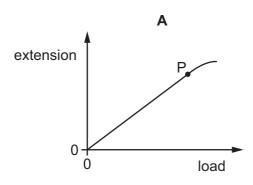


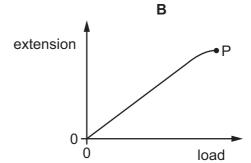
Compared with this cube, which statement about a cube of aluminium with sides of 2.0 cm is correct?

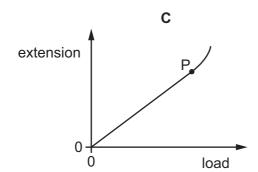
- A It has the same density.
- B It has the same mass.
- **C** It has twice the density.
- **D** It has twice the mass.

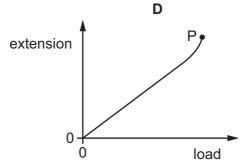
30 A stretching force is applied to a copper wire. The wire obeys Hooke's Law until it reaches the limit of proportionality.

Which is the extension-load graph for the wire and shows the limit of proportionality labelled P?









31 A brick of mass 4.0 kg rests on a window ledge. It falls off the window ledge and drops through a height of $5.0 \,\mathrm{m}$ to the ground. The acceleration of free fall g is $10 \,\mathrm{m/s^2}$.

Air resistance can be ignored.

Which row states the kinetic energy and the speed of the brick just before it hits the ground?

	kinetic energy of brick/J	speed of brick m/s
Α	20	2.2
В	20	3.2
С	200	7.1
D	200	10

32 A liquid evaporates when molecules leave its surface.

Which molecules leave the surface, and what happens to the temperature of the remaining liquid?

- A The more energetic molecules leave and the temperature falls.
- **B** The more energetic molecules leave and the temperature rises.
- **C** The less energetic molecules leave and the temperature falls.
- **D** The less energetic molecules leave and the temperature rises.
- 33 Convection is a process by which thermal energy is transferred from one place to another.

Where can convection take place?

- A in a gas and in a vacuum
- **B** in a liquid and in a gas
- C in a liquid and in a solid
- **D** in a solid and in a vacuum

34 A musical instrument produces a note of frequency 170 Hz. The sound wave produced travels through the air at a speed of 340 m/s.

Which row describes the sound wave?

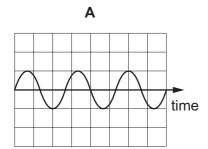
	nature of wave	wavelength / m
Α	longitudinal	0.50
В	longitudinal	2.0
С	transverse	0.50
D	transverse	2.0

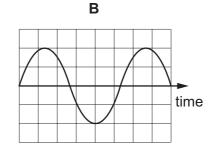
35 A girl stands in front of a plane mirror and observes her image. She walks 2.0 m towards the mirror.

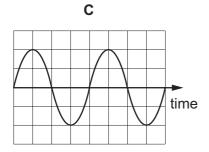
What is the change in the distance between the girl and her image?

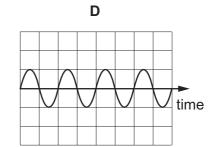
- **A** 0 m
- **B** 1.0 m
- **C** 2.0 m
- **D** 4.0 m
- 36 The diagrams represent four different sound waves. The scales are the same in all the diagrams.

Which sound has the lowest pitch?



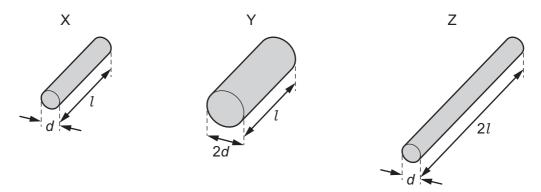






37 Three pieces of resistance wire X, Y and Z are made of the same metal.

The diagram shows the lengths and the diameters of the wires.



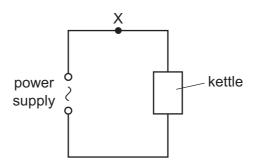
What is the order of the wires when they are placed in order of increasing resistance, least resistance first?

- $A Y \to X \to Z$
- **B** $Y \rightarrow Z \rightarrow X$ **C** $Z \rightarrow X \rightarrow Y$
- $D Z \rightarrow Y \rightarrow X$
- **38** A lamp has a potential difference *V* across it that causes a current *I* in it.

Which equation gives the power *P* produced by the lamp?

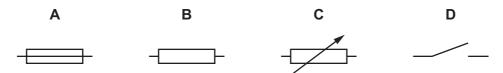
- **A** $P = \frac{I}{V}$ **B** $P = \frac{V}{I}$ **C** P = IV **D** P = I + V

- **39** A kettle is connected to a power supply as shown.



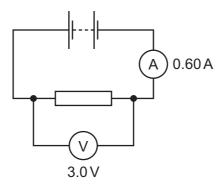
If too much current flows, a component connected at X automatically disconnects the power supply.

Which symbol represents the component at X?

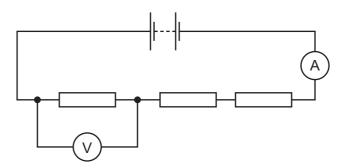


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40 A student sets up the circuit shown in the diagram. The ammeter reads 0.60 A and the voltmeter reads 3.0 V.



The student now takes two resistors that are identical to the original resistor. She connects them in series with the original resistor.



What are the new readings on the ammeter and the voltmeter?

	ammeter/A	voltmeter/V
Α	0.20	1.0
В	0.20	3.0
С	0.60	1.0
D	0.60	3.0

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The Periodic Table of Elements

	II /	5	Ηœ	helium 4	10	Ne	neon 20	18	Ā	argon 40	36	궃	krypton 84	54	Xe	xenon 131	98	R	radon			
	II/				6	ш	fluorine 19	17	Cl	chlorine 35.5	35	ğ	bromine 80	53	П	iodine 127	85	¥	astatine -			
					∞	0	oxygen 16	16	S	sulfur 32	34	Se	selenium 79	52	Те	tellurium 128	84	Ъ	polonium —	116		livermorium -
	>				7	z	nitrogen 14	15	۵	phosphorus 31	33	As	arsenic 75	51	Sp	antimony 122	83	Ξ	bismuth 209			
	2				9	ပ	carbon 12	14	Si	silicon 28	32	Ge	germanium 73	20	Sn	tin 119	82	Pb	lead 207	114	Fl	flerovium
	≡				2	В	boron 11	13	Ν	aluminium 27	31	Ga	gallium 70	49	In	indium 115	81	<i>1</i> L	thallium 204			
								1			30	Zu	zinc 65	48	р О	cadmium 112	80	Нg	mercury 201	112	S	copemicium –
											29	Cn	copper 64	47	Ag	silver 108	62	Au	gold 197	111	Rg	roentgenium -
dn											28	Z	nickel 59	46	Pq	palladium 106	78	귙	platinum 195	110	Ds	darmstadtium -
Group											27	ဝိ	cobalt 59	45	格	rhodium 103	77	'n	iridium 192	109	¥	meitnerium -
		- 1	I	hydrogen 1							26	Fe	iron 56	44	R	ruthenium 101	92	SO	osmium 190	108	Hs	hassium
					J						25	Mn	manganese 55	43	ည	technetium -	75	Re	rhenium 186	107	Bh	bohrium
						loc	SS				24		chromium 52		Mo	molybdenum 96	74	≯	tungsten 184	106	Sg	seaborgium
				Key	atomic number	atomic symbo	name relative atomic mass				23	>	vanadium 51	41	q	niobium 93	73	<u>ra</u>	tantalum 181	105	Q D	dubnium -
					to	ato	rela				22	j	titanium 48	40	Zr	zirconium 91	72	Έ	hafnium 178	104	弘	rutherfordium -
								_			21	လွ	scandium 45	39	>	yttrium 89	57-71	lanthanoids		89–103	actinoids	
	=				4	Be	beryllium 9	12	Mg	magnesium 24	20	Ca	calcium 40	38	ഗ്	strontium 88	56	Ba	barium 137	88	Ra	radium -
	-				3	:=	lithium 7	11	Na	sodium 23	19	¥	potassium 39	37	Rb	rubidium 85	55	Cs	caesium 133	87	뇬	francium -

La lanthanum 139 89	Ce cerium 140	Pr praseodymium 141	Nd neodymium 144	Pm promethium	Sm samarium 150	Eu europium 152 95	Gd gadolinium 157	65 Tb terbium 159	Dy dysprosium 163	Ho holmium 165	68 Erbium 167	thulium 169	70 Yb ytterbium 173	Lu lutetium 175	
Ac	Th	Pa protactinium 231	uranium 238	Np neptunium	Pu plutonium	Am	Cm	Bk berkelium	Cf californium	Estation Est	Fm	Md	No	Lr lawrencium	

The volume of one mole of any gas is $24\,\mathrm{dm}^3$ at room temperature and pressure (r.t.p.).