

**CHEMISTRY**

**0620/11**

Paper 1 Multiple Choice

**October/November 2014**

**45 Minutes**

Additional Materials: Multiple Choice Answer Sheet  
Soft clean eraser  
Soft pencil (type B or HB is recommended)



**READ THESE INSTRUCTIONS FIRST**

Write in soft pencil.

Do not use staples, paper clips, glue or correction fluid.

Write your name, Centre number and candidate number on the Answer Sheet in the spaces provided unless this has been done for you.

**DO NOT WRITE IN ANY BARCODES.**

There are **forty** questions on this paper. Answer **all** questions. For each question there are four possible answers **A, B, C** and **D**.

Choose the **one** you consider correct and record your choice in **soft pencil** on the separate Answer Sheet.

**Read the instructions on the Answer Sheet very carefully.**

Each correct answer will score one mark. A mark will not be deducted for a wrong answer.

Any rough working should be done in this booklet.

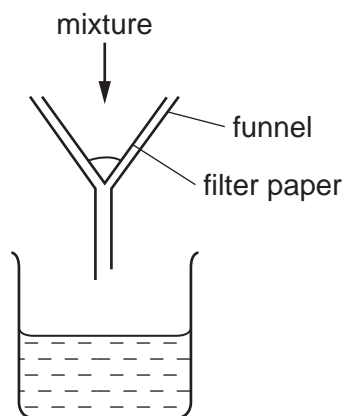
A copy of the Periodic Table is printed on page 16.

Electronic calculators may be used.

The syllabus is approved for use in England, Wales and Northern Ireland as a Cambridge International Level 1/Level 2 Certificate.

This document consists of **13** printed pages and **3** blank pages.

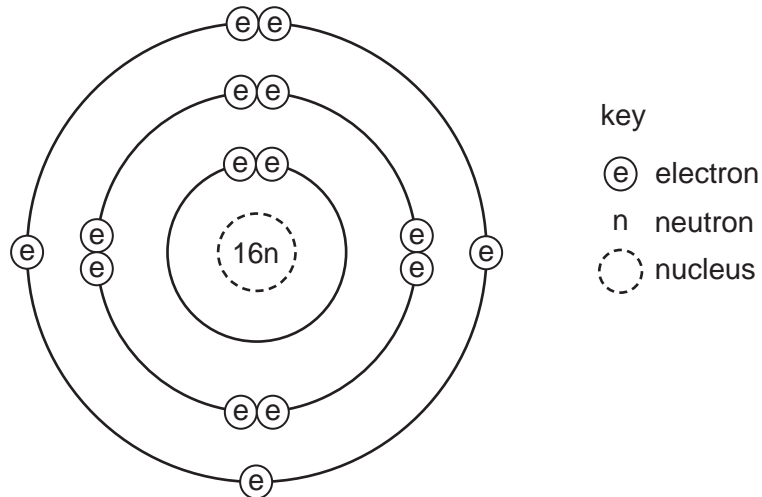
- 1 Which statement is an example of diffusion?
- A A kitchen towel soaks up some spilt milk.
  - B Ice cream melts in a warm room.
  - C Pollen from flowers is blown by the wind.
  - D The smell of cooking spreads through a house.
- 2 A mixture is separated using the apparatus shown.



What is the mixture?

- A aqueous copper chloride and copper
  - B aqueous copper chloride and sodium chloride
  - C ethane and methane
  - D ethanol and water
- 3 Ethanol is made by fermentation.
- How is ethanol obtained from the fermentation mixture?
- A chromatography
  - B crystallisation
  - C electrolysis
  - D fractional distillation
- 4 What is different for isotopes of the same element?
- A nucleon number
  - B number of electron shells
  - C number of electrons in the outer shell
  - D proton number

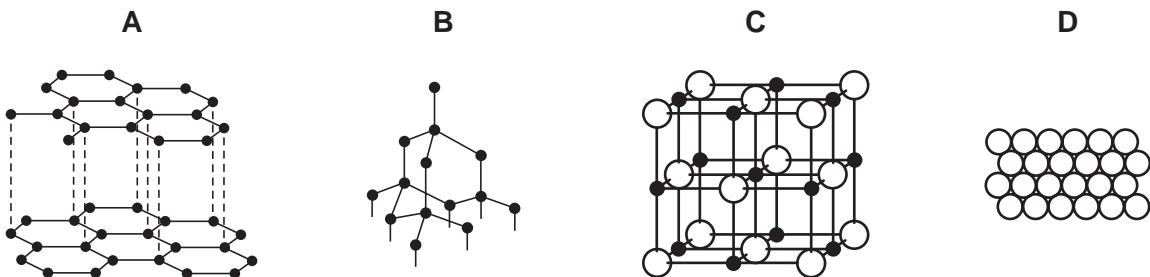
5 Which element has the atomic structure shown?



- A Al                      B P                      C S                      D Si

6 Slate has a layered structure and can easily be split into thin sheets.

Which diagram shows a structure most like that of slate?



7 Sodium chloride is an ionic solid.

Which statement is **not** correct?

- A Ions are formed when atoms lose or gain electrons.
- B Ions in sodium chloride are strongly held together.
- C Ions with the same charge attract each other.
- D Sodium chloride solution can conduct electricity.

- 8 Caesium chloride and rubidium bromide are halide compounds of Group I elements.

Caesium chloride has the formula .....1....., a relative formula mass .....2..... that of rubidium bromide and bonds that are .....3..... .

Which words correctly complete gaps 1, 2 and 3?

	1	2	3
<b>A</b>	$\text{CaCl}$	different from	ionic
<b>B</b>	$\text{CaCl}$	the same as	covalent
<b>C</b>	$\text{CsCl}$	different from	ionic
<b>D</b>	$\text{CsCl}$	the same as	covalent

- 9 How many atoms of hydrogen are there in a molecule of ethanol,  $\text{C}_2\text{H}_5\text{OH}$ ?

**A** 1                      **B** 2                      **C** 5                      **D** 6

- 10 Iron forms an oxide with the formula  $\text{Fe}_2\text{O}_3$ .

What is the relative formula mass of this compound?

**A** 76                      **B** 100                      **C** 136                      **D** 160

- 11 Which metal could **not** be used for electroplating by using an aqueous solution?

**A** chromium  
**B** copper  
**C** silver  
**D** sodium

- 12 Which products are formed at the electrodes when a concentrated solution of sodium chloride is electrolysed?

	cathode (-)	anode (+)
<b>A</b>	hydrogen	chlorine
<b>B</b>	hydrogen	oxygen
<b>C</b>	sodium	chlorine
<b>D</b>	sodium	oxygen

13 Which statements about exothermic and endothermic reactions are correct?

- 1 During an exothermic reaction, heat is given out.
- 2 The temperature of an endothermic reaction goes up because heat is taken in.
- 3 Burning methane in the air is an exothermic reaction.

A 1, 2 and 3    B 1 and 2 only    C 1 and 3 only    D 2 and 3 only

14 A power station was designed to burn gaseous fuels only.

Which two substances could be used?

A carbon dioxide and hydrogen

B carbon dioxide and  $^{235}\text{U}$

C hydrogen and methane

D methane and  $^{235}\text{U}$

15 The rate of a reaction depends on temperature, concentration, particle size and catalysts.

Which statement is **not** correct?

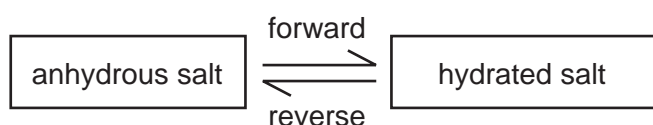
A Catalysts can be used to increase the rate of reaction.

B Higher concentration decreases the rate of reaction.

C Higher temperature increases the rate of reaction.

D Larger particle size decreases the rate of reaction.

16 The diagram shows the change from an anhydrous salt to its hydrated form.



Which statement is correct?

A forward reaction requires heat and water

B forward reaction requires water only

C reverse reaction requires heat and water

D reverse reaction requires water only

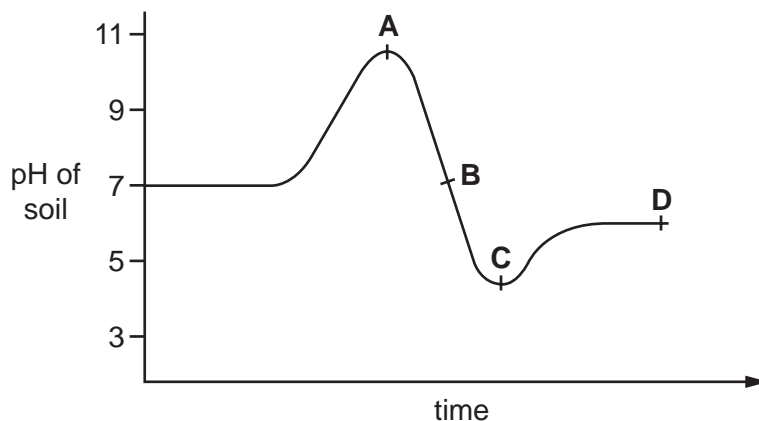


21 How many different salts could be made from a supply of dilute sulfuric acid, dilute hydrochloric acid, copper, magnesium oxide and zinc carbonate?

- A 3                      B 4                      C 5                      D 6

22 The graph shows how the pH of soil in a field changes over time.

At which point was the soil neutral?



23 Elements in Group I of the Periodic Table react with water.

Which row describes the products made in the reaction and the trend in reactivity of the elements?

	products	trend in reactivity
<b>A</b>	metal hydroxide and hydrogen	less reactive down the group
<b>B</b>	metal hydroxide and hydrogen	more reactive down the group
<b>C</b>	metal oxide and hydrogen	less reactive down the group
<b>D</b>	metal oxide and hydrogen	more reactive down the group

24 An element X has the two properties listed.

- 1 It acts as a catalyst.
- 2 It forms colourless ions.

Which of these properties suggest that X is a transition element?

	property 1	property 2
<b>A</b>	✓	✓
<b>B</b>	✓	x
<b>C</b>	x	✓
<b>D</b>	x	x

25 An inert gas X is used to fill weather balloons.

Which descriptions of X are correct?

	number of outer electrons in atoms of X	structure of gas X
<b>A</b>	2	single atoms
<b>B</b>	2	diatomic molecules
<b>C</b>	8	single atoms
<b>D</b>	8	diatomic molecules

26 The table shows the reactions of four different metals with water.

metal	reaction
W	reacts vigorously with cold water
X	no reaction with water
Y	reacts very slowly with water, more vigorously with steam
Z	reacts violently with cold water

What is the correct order of reactivity, from most reactive to least reactive?

- A**  $W \rightarrow X \rightarrow Y \rightarrow Z$
- B**  $W \rightarrow Z \rightarrow Y \rightarrow X$
- C**  $Z \rightarrow W \rightarrow X \rightarrow Y$
- D**  $Z \rightarrow W \rightarrow Y \rightarrow X$

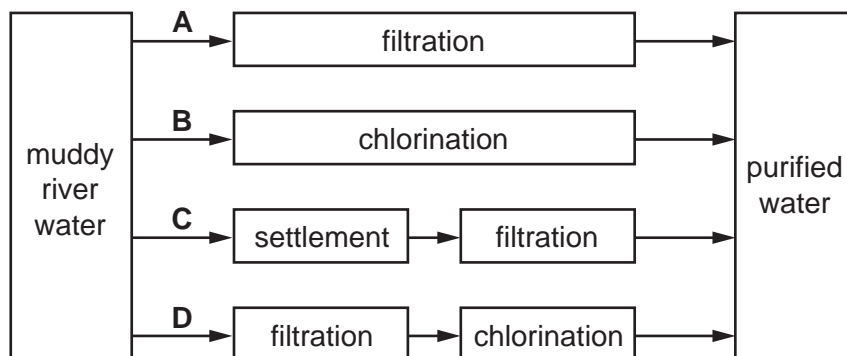


- 27 Which information about an element can be used to predict its chemical properties?
- A boiling point
  - B density
  - C melting point
  - D position in the Periodic Table
- 28 Aluminium is the most common metal in the Earth's crust.
- Which is **not** a property of aluminium?
- A low density
  - B resistance to corrosion
  - C good conductor of electricity
  - D poor conductor of heat
- 29 The oxide of element X is reduced by heating with carbon.
- Element X does not react with cold water, steam or dilute hydrochloric acid.
- What is X?
- A copper
  - B iron
  - C magnesium
  - D zinc
- 30 Which object is **least** likely to contain aluminium?
- A a bicycle frame
  - B a hammer
  - C a saucepan
  - D an aeroplane body
- 31 Which reaction involves oxidation?
- A heating hydrated copper(II) sulfate in the air
  - B polymerisation of ethene
  - C rusting of iron
  - D thermal decomposition of calcium carbonate

32 Which method can be used to obtain ammonia from ammonium sulfate?

- A Heat it with an acid.
- B Heat it with an alkali.
- C Heat it with an oxidising agent.
- D Heat it with a reducing agent.

33 Which method of purification would produce water **most** suitable for drinking?



34 Which statement about methane is **not** correct?

- A It is a liquid produced by distilling petroleum.
- B It is produced as vegetation decomposes.
- C It is produced by animals, such as cows.
- D It is used as a fuel.

35 Which is an air pollutant that affects a part of the body other than the lungs and blood system?

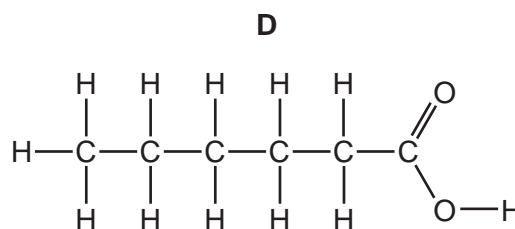
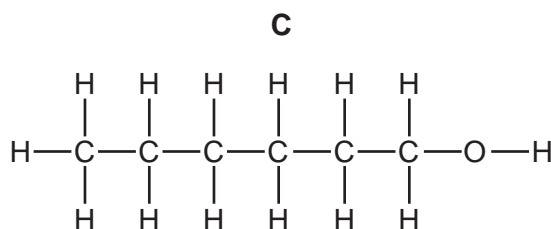
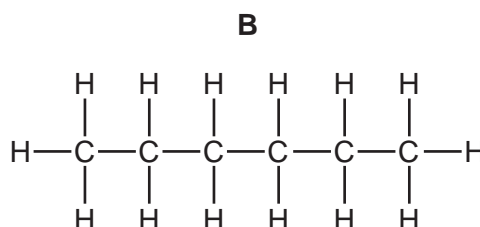
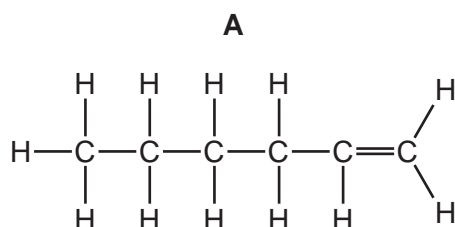
- A lead compounds
- B nitrogen
- C oxides of nitrogen
- D sulfur dioxide

- 36 Increasing the number of atoms in one molecule of a hydrocarbon increases the amount of energy released when it burns.

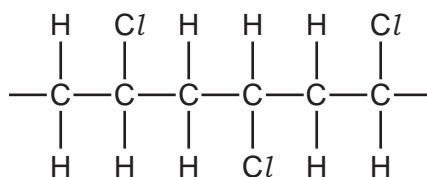
What is the correct order?

	less energy released	→	more energy released
<b>A</b>	ethene		ethane
<b>B</b>	ethene		methane
<b>C</b>	methane		ethane
<b>D</b>	methane		ethene

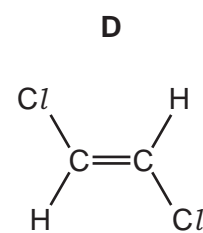
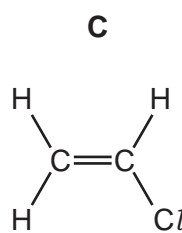
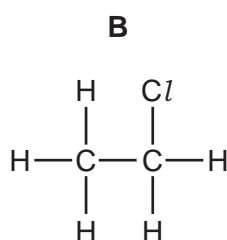
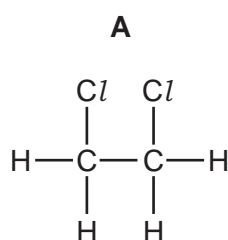
- 37 Which molecular structure shows hexene?



- 38 The diagram shows three repeat units in the structure of an addition polymer.



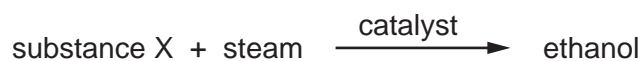
Which alkene monomer is used to make this polymer?



39 Which statement about alkenes is **not** correct?

- A The functional group is C=C.
- B The structural difference between one member and the next is  $-\text{CH}_2-$ .
- C They form a homologous series.
- D They turn aqueous bromine from brown to colourless.

40 Ethanol can be manufactured from substance X.



What is substance X?

- A carbon dioxide
- B ethene
- C hydrogen
- D oxygen







**DATA SHEET**  
**The Periodic Table of the Elements**

		Group																																																																										
		I	II	III	IV	V	VI	VII	0																																																																			
		1 <b>H</b> Hydrogen 1																																																																										
7	9	<b>Li</b> Lithium 3	<b>Be</b> Beryllium 4																																																																									
23	24	<b>Na</b> Sodium 11	<b>Mg</b> Magnesium 12																																																																									
39	40	<b>K</b> Potassium 19	<b>Ca</b> Calcium 20	45 <b>Sc</b> Scandium 21	48 <b>Ti</b> Titanium 22	51 <b>V</b> Vanadium 23	52 <b>Cr</b> Chromium 24	55 <b>Mn</b> Manganese 25	56 <b>Fe</b> Iron 26	59 <b>Co</b> Cobalt 27	59 <b>Ni</b> Nickel 28	64 <b>Cu</b> Copper 29	65 <b>Zn</b> Zinc 30	70 <b>Ga</b> Gallium 31	73 <b>Ge</b> Germanium 32	75 <b>As</b> Arsenic 33	79 <b>Se</b> Selenium 34	80 <b>Br</b> Bromine 35	84 <b>Kr</b> Krypton 36																																																									
85	88	<b>Rb</b> Rubidium 37	<b>Sr</b> Strontium 38	89 <b>Y</b> Yttrium 39	91 <b>Zr</b> Zirconium 40	93 <b>Nb</b> Niobium 41	96 <b>Mo</b> Molybdenum 42	101 <b>Ru</b> Ruthenium 44	101 <b>Rh</b> Rhodium 45	103 <b>Rh</b> Rhodium 45	106 <b>Pd</b> Palladium 46	108 <b>Ag</b> Silver 47	112 <b>Cd</b> Cadmium 48	115 <b>In</b> Indium 49	119 <b>Sn</b> Tin 50	122 <b>Sb</b> Antimony 51	128 <b>Te</b> Tellurium 52	127 <b>I</b> Iodine 53	131 <b>Xe</b> Xenon 54																																																									
133	137	<b>Cs</b> Caesium 55	<b>Ba</b> Barium 56	139 <b>La</b> Lanthanum 57	178 <b>Hf</b> Hafnium 72	181 <b>Ta</b> Tantalum 73	184 <b>W</b> Tungsten 74	190 <b>Os</b> Osmium 76	192 <b>Ir</b> Iridium 77	195 <b>Pt</b> Platinum 78	197 <b>Au</b> Gold 79	201 <b>Hg</b> Mercury 80	204 <b>Tl</b> Thallium 81	207 <b>Pb</b> Lead 82	209 <b>Bi</b> Bismuth 83	210 <b>Po</b> Polonium 84	210 <b>At</b> Astatine 85	210 <b>Rn</b> Radon 86																																																										
	226	<b>Fr</b> Francium 87	<b>Ra</b> Radium 88	227 <b>Ac</b> Actinium 89																																																																								
		*58-71 Lanthanoid series										†90-103 Actinoid series																																																																
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		140 <b>Ce</b> Cerium 58	141 <b>Pr</b> Praseodymium 59	144 <b>Nd</b> Neodymium 60	146 <b>Pm</b> Promethium 61	150 <b>Sm</b> Samarium 62	152 <b>Eu</b> Europium 63	157 <b>Gd</b> Gadolinium 64	159 <b>Tb</b> Terbium 65	162 <b>Dy</b> Dysprosium 66	165 <b>Ho</b> Holmium 67	167 <b>Er</b> Erbium 68	169 <b>Tm</b> Thulium 69	173 <b>Yb</b> Ytterbium 70	175 <b>Lu</b> Lutetium 71	232 <b>Th</b> Thorium 90	238 <b>Pa</b> Protactinium 91	238 <b>U</b> Uranium 92	238 <b>Np</b> Neptunium 93	241 <b>Pu</b> Plutonium 94	244 <b>Am</b> Americium 95	247 <b>Cm</b> Curium 96	251 <b>Bk</b> Berkelium 97	259 <b>Cf</b> Californium 98	265 <b>Es</b> Einsteinium 99	271 <b>Fm</b> Fermium 100	289 <b>Md</b> Mendelevium 101	289 <b>No</b> Nobelium 102	289 <b>Lr</b> Lawrencium 103																																															

The volume of one mole of any gas is 24 dm<sup>3</sup> at room temperature and pressure (r.t.p.).

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