

**CAMBRIDGE INTERNATIONAL EXAMINATIONS**  
International General Certificate of Secondary Education

## **MARK SCHEME for the May/June 2014 series**

### **0620 CHEMISTRY**

**0620/62**

Paper 6 (Alternative to Practical), maximum raw mark 60

This mark scheme is published as an aid to teachers and candidates, to indicate the requirements of the examination. It shows the basis on which Examiners were instructed to award marks. It does not indicate the details of the discussions that took place at an Examiners' meeting before marking began, which would have considered the acceptability of alternative answers.

Mark schemes should be read in conjunction with the question paper and the Principal Examiner Report for Teachers.

Cambridge will not enter into discussions about these mark schemes.

Cambridge is publishing the mark schemes for the May/June 2014 series for most IGCSE, GCE Advanced Level and Advanced Subsidiary Level components and some Ordinary Level components.

Page 2	Mark Scheme	Syllabus	Paper
	IGCSE – May/June 2014	0620	62

- 1 (a) beaker (1) [1]
- (b) (i) electrolysis (1) [1]
- (ii) electrodes (1) [1]  
**allow:** conduct electricity / to transfer electrons  
**ignore:** attract ions
- (c) hydrogen:
- lighted splint (1)
- pops (1)
- OR**
- chlorine:
- litmus (1)
- bleached (1) [2]
- (d) diagram to show test-tubes above electrodes (1)
- containing liquid (1) [2]
- 2 (a) pipette/burette (1) [1]  
**ignore:** measuring cylinder
- (b) Universal / pH indicator / pH paper / full range (1) [1]  
**ignore:** indicator  
**not:** other named indicator

Page 3	Mark Scheme	Syllabus	Paper
	IGCSE – May/June 2014	0620	62

- (c) pH value rises/increases/becomes more alkaline (1)
- steep change in middle (1) [2]  
**allow:** suddenly/drastically/quoted figures
- (d) (i) end/neutralisation/equivalence point/becomes neutral (1) [1]  
**allow:** reaction finished/changes from acid to alkali/basic
- (ii) 12.5 (1)
- cm<sup>3</sup> (1) [2]
- (iii) potassium hydroxide solution is 2 × (1)
- more concentrated/stronger (1) ORA
- half volume of potassium hydroxide used/twice volume of nitric acid used (1) [3]
- (e) evaporation/steam (1)
- solid/crystals formed (1) [2]  
**allow:** decomposes or named products
- 3 (a) Bunsen/burner (1) [1]  
**ignore:** heat/heater
- (b) not pure/not just ethene (1)  
**not:** a different alkane or alkene is formed first
- contains air (from the tube when heated) (1) [2]  
**ignore:** oxygen
- (c) catalyst/to provide a large surface area (1) [1]
- (d) bromine (water) (1)  
**not:** bromide
- colourless/decolourised in alkene **or** stays orange in alkane (1) [2]  
**allow:** colour change ecf

Page 4	Mark Scheme	Syllabus	Paper
	IGCSE – May/June 2014	0620	62

- 4 (c) Experiment 1: Table of results
- initial temperature boxes completed correctly (2) 27, 28, 31, 30, 31
- highest temperature boxes completed correctly (2) 33, 36, 42, 45, 49
- temperature changes correct (1) 6, 8, 11, 15, 18 [5]
- (d) All points correctly plotted (3)  
**guidance:** 5 correct (3); 4 correct (2); 3 correct (1); 2 or fewer correct (0)
- best fit straight line graph drawn with a ruler (1) [4]  
**note:** does not need to go through origin
- (e) value from graph (1), e.g. 21
- °C (1)
- extrapolation to 8 cm / indication shown (1) [3]
- (f) magnesium smaller / disappears / fizzing / bubbles / effervescence (1) [1]  
**ignore:** gas
- (g) (i) Experiment 5 (1) [1]  
**allow:** 7 cm
- (ii) more / most / longest / 7 cm magnesium used (1) [1]  
**ignore:** reactant / sulfuric acid / surface area
- (h) temperature change / reaction faster (1)  
**ignore:** temperature rise
- more surface area (1) [2]
- (i) 3 (°C)  
**allow:** 2–5

Page 5	Mark Scheme	Syllabus	Paper
	IGCSE – May/June 2014	0620	62

- (j) shows gas collected over water (1)  
in labelled measuring cylinder/graduations shown on collection vessel (1)

**OR**

- shows gas collected in a gas syringe (1)  
in labelled (gas) syringe/graduations shown (1) [2]

- (k) error...heat losses/using measuring cylinder/oxide layer (1)  
**ignore:** initial temperature  
improvement...insulation/use burette or pipette/clean/repeat (1) [2]

- 5 (b) pH paper turns blue/pH > 7/reference to smell of the gas (1) [1]

- (c) (i) paper turns blue / pH > 7 (1)  
reference to smell of gas (1)  
**ignore:** fizzing

- (ii) white (1)  
precipitate (1)

- (f) zinc (1)  
**allow:** Zn<sup>2+</sup>  
**ignore:** incorrect formulae  
carbonate (1)  
**allow:** CO<sub>3</sub><sup>2-</sup>  
**ignore:** incorrect formulae

<b>Page 6</b>	<b>Mark Scheme</b>	<b>Syllabus</b>	<b>Paper</b>
	<b>IGCSE – May/June 2014</b>	<b>0620</b>	<b>62</b>

6 crush (1)

with...pestle and mortar/hammer

**OR**

reason...to increase the surface area/to make smaller pieces/to increase the rate of reaction (1)  
[2]

**Followed by:**

heat (1)

with carbon (1)

any **two** from: carbon is more reactive/displaces Pb/takes away oxygen/forms carbon dioxide/reduction (2)  
[4]

**OR**

heat (1)

with a named metal between Mg and Pb in reactivity series, e.g. Fe (1)

more reactive/displaces Pb/takes away oxygen/reduction (1)

separation of Pb and metal oxide (1)  
[4]

**allow:** heat to melt lead and run off/decant

**OR**

heat (1)

with carbon/CO (1)

PbO (1)

heat with carbon/CO (1)  
[4]

**OR**

heat (1)

with iron (1)

PbO (1)

separation (1)  
[4]

**OR**

dilute acid (1)

**allow:** any dilute acid

**ignore:** heating

<b>Page 7</b>	<b>Mark Scheme</b>	<b>Syllabus</b>	<b>Paper</b>
	<b>IGCSE – May/June 2014</b>	<b>0620</b>	<b>62</b>

$\text{Pb}^{2+}_{(\text{aq})}$ / salt/ solution (1)

iron (1)

displaces lead (1)

[4]

**OR**

dilute acid (1)

**allow:** any dilute acid

**ignore:** heating

$\text{Pb}^{2+}_{(\text{aq})}$ / salt/ solution (1)

electrolysis (1)

**ignore:** heating

lead deposited (at cathode) (1)

[4]