

**CAMBRIDGE INTERNATIONAL EXAMINATIONS**  
International General Certificate of Secondary Education

## **MARK SCHEME for the May/June 2014 series**

### **0620 CHEMISTRY**

**0620/31**

Paper 3 (Extended Theory), maximum raw mark 80

This mark scheme is published as an aid to teachers and candidates, to indicate the requirements of the examination. It shows the basis on which Examiners were instructed to award marks. It does not indicate the details of the discussions that took place at an Examiners' meeting before marking began, which would have considered the acceptability of alternative answers.

Mark schemes should be read in conjunction with the question paper and the Principal Examiner Report for Teachers.

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- 1 (a) A, D, E (1)  
 same number of protons and electrons/electrically neutral (1) [2]
- (b) C (1)  
 more electrons than protons /  $36e^-$  and  $34p^+$  / it has gained electrons (1) [2]
- (c) B, F (1) [1]
- (d) they have same number of protons (1)  
 different number of neutrons / neutron number (1) [2]
- [Total: 7]
- 2 (a) (i) filtration (1)  
 chlorination (1) [2]
- (ii) Any **two** from: [2]
- manufacture of ethanol
  - used in the manufacture of sulfuric acid **or** in the Contact process
  - manufacture of hydrogen **or** ammonia **or** for the Haber process
- (iii) Any **two** from: [2]
- cooking
  - washing or laundry
  - drinking
  - toilets
  - watering plants
  - (domestic) heating
- (b) boiling or turning to steam (1)  
**then** condensing / condensation (1) [2]
- [Total: 7]
- 3 (a) (i) (particles) spread to fill total available volume / move from high concentration to low concentration / moves down a concentration gradient (1) [1]
- (ii) mass or  $M_r$  (1) [1]
- (b) (i) helium atoms / molecules are lighter than molecules in air or  $N_2$  **and**  $O_2$   
**or** helium is less dense than air or  $N_2$  **and**  $O_2$ .  
**or** helium diffuses (through the porous barrier) faster than air or  $N_2$  **and**  $O_2$ . (1) [1]

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(ii) faster rate of diffusion / molecules move faster (at high temperatures). (1) [1]

(c) (i)  $\text{CH}_4 + 2\text{O}_2 \rightarrow \text{CO}_2 + 2\text{H}_2\text{O}$  (1) [1]

(ii) would get a mixture of helium and carbon dioxide  
**or** would get a mixture of gases  
**or** waste of methane / natural gas / fossil fuel (1) [1]

(iii) fractional distillation (1) [1]

[Total: 7]

4 (a) (i)

Group number	I	II	III	IV	V	VI	VII
symbol	Na	Mg	Al	Si	P	S	Cl
number of valency electrons	1	2	3	4	5	6	7
valency	1	2	3	4	3	2	1

(1) for each line [2]

(ii) number of valency electrons = the group number (1) [1]

(iii) for Na to Al  
the valency is the same as the number of valency (outer) electrons (1)

(because) this is the number of electrons **lost** (for full energy level) (1)

for P to Cl

the valency is 8 – [number of valency (outer) electrons]

**or** valency + valency electrons = 8 (1)

(because) this is number of electrons **needed** (or to be **gained**) (for full energy level) (1)

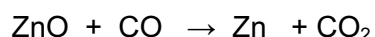
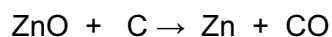
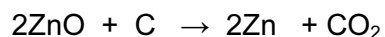
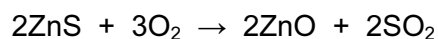
(b) (i) Assume change is from L to R unless clearly stated:  
basic to amphoteric to acidic (2) [2]

(ii) ionic (metal) chlorides on the left (1)  
covalent (non-metal) chlorides on the right (1) [2]

[Total: 11]

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- 5 (a) M1: (zinc sulfide) heated / roasted / burnt in **air** (1)
- M2: zinc oxide formed (1)
- M3: zinc oxide **reduced** (1)
- M4: (by adding) coke or carbon (1)
- M5: Balanced equation (any one of) (1) [5]



- (b) Any **two** from: [2]

- (making) brass **or** alloys (1)
- galvanising (1)
- sacrificial protection (1)
- batteries (1)

[Total: 7]

- 6 (a) (i) rate at  $t_2$  less than at  $t_1$  **or** the rate decreases (1)
- rate at  $t_3$  zero / reaction stopped (1) [2]
- (ii) rate at  $t_2$  less than at  $t_1$  because **concentration** of hydrogen peroxide is less at  $t_2$  **or concentration** of hydrogen peroxide is decreasing. (1)
- (rate at  $t_3$  zero / reaction stopped because) hydrogen peroxide is used up (1) [2]
- (b) (i) steeper and must come from the origin (1)
- final volumes the same (1) [2]
- (ii) Any **two** from: [2]
- steeper curve because of a faster rate
- faster rate because of increased surface area
- same amount / volume / mass / no of mol of hydrogen peroxide
- ecf for M1 for a shallower curve because of slower rate.

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- (c) filter (and rinse/wash) (1)
- dry manganese (IV) oxide (1)
- weigh/measure mass manganese(IV) oxide after reaction (1)
- the mass should be 0.1 g **or** unchanged. (1) [4]

- (d) number of moles of O<sub>2</sub> formed = 0.096/24 = 0.004 (1)
- number of moles of H<sub>2</sub>O<sub>2</sub> in 40 cm<sup>3</sup> of solution = 0.004 × 2 = 0.008 (1)
- concentration of the hydrogen peroxide in mol/dm<sup>3</sup> = 0.008/0.04 = 0.2 (1) [3]

[Total:15]

7 (a) (i)

aqueous solution	lead Pb	magnesium Mg	zinc Zn	silver Ag
lead (II) nitrate		✓	✓	x
magnesium nitrate	x x		x	x
zinc nitrate	x	✓		x
silver(I) nitrate	✓	✓	✓	

each horizontal line correct (1) [3]

(ii) Zn (1)

An arrow **from** Zn **to** Zn<sup>2+</sup> (1) [2]

(iii) Zn + 2Ag<sup>+</sup> → Zn<sup>2+</sup> + 2Ag (1) [1]

(b) (i) correct direction from zinc to lead (1) [1]

(ii) metals react by **losing electrons** (1)

the more reactive metal/zinc will lose electrons more readily (making the electrode negatively charged). (1) [2]

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- (iii) manganese **and** zinc are more reactive than lead (and/or copper) (1)  
 lead is more reactive than copper (1) [2]

- (iv) the **polarity** of a Mn/Zn (cell)  
**or** the **voltages** of Zn/Pb **and** Mn/Pb (cells) (1) [1]

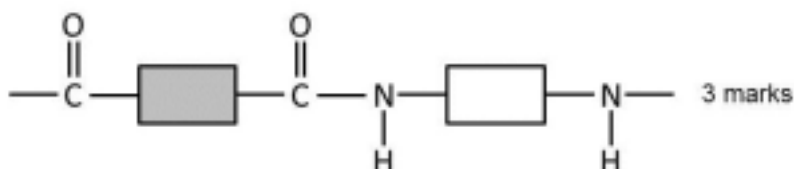
[Total: 12]

- 8 (a) (i)  $\text{CH}_3\text{-CH=CH-CH}_3$  (1) [1]

- (ii) one correct amide linkage between two rectangles (1)

correct sequencing of a second amide link and monomers (1)

two correct amide links **and** rest of structure correct (including additional monomers if seen) **and** correct continuation bonds (1) [3]



- (iii) protein **or** polypeptide **or** named protein (1) [1]

- (iv) addition: **only** the polymer **or** one product is formed (1)

condensation: the polymer **and** a small molecule/water/HCl is formed (1) [2]

- (b) (i) does not break down **or** rot **or** decompose (1)

by microbes **or** fungi **or** bacteria **or** by living organisms (1) [2]

- (ii) Any **three** from: [3]  
 visual pollution (1)

(shortage of) landfill sites (1)

danger to wildlife/animals (including at sea) (1)

toxic gases when burnt **or** greenhouse gases produced when burned (1)

- (c) Any **two** from: [2]  
 resistant to corrosion/unreactive to water/more durable (1)

lighter/less dense (1)

easier to manufacture/can be moulded (1)

good insulator/keeps the water cold (1)

[Total: 14]