

**CAMBRIDGE INTERNATIONAL EXAMINATIONS**  
International General Certificate of Secondary Education

## **MARK SCHEME for the May/June 2013 series**

### **0620 CHEMISTRY**

**0620/52**

Paper 5 (Practical), maximum raw mark 40

This mark scheme is published as an aid to teachers and candidates, to indicate the requirements of the examination. It shows the basis on which Examiners were instructed to award marks. It does not indicate the details of the discussions that took place at an Examiners' meeting before marking began, which would have considered the acceptability of alternative answers.

Mark schemes should be read in conjunction with the question paper and the Principal Examiner Report for Teachers.

Cambridge will not enter into discussions about these mark schemes.

Cambridge is publishing the mark schemes for the May/June 2013 series for most IGCSE, GCE Advanced Level and Advanced Subsidiary Level components and some Ordinary Level components.

Page 2	Mark Scheme	Syllabus	Paper
	IGCSE – May/June 2013	0620	52

- 1 (a) Table of results for Experiment 1
- initial and final volumes and differences completed correctly (1)
- within  $\pm 2$  Supervisor (1)
- all results (both tables) to 1 or 2 decimal places (including 0.0) (1) [3]
- (b) Table of results for Experiment 2
- initial and final volumes and differences completed correctly (1)
- titre lower than experiment (1)
- within  $\pm 2$  Supervisor (1) [3]
- (c) (i) to speed up the reaction / owtte (1) [1]
- (ii) colourless (1) **not:** clear, to brown / pink / purple / lilac / mauve (1) [2]
- (iii) not an acid / alkali reaction or potassium manganate is coloured / self indicating owtte (1) [1]
- (d) (i) experiment 1 **allow:** ecf from results (1) [1]
- (ii) experiment 1 (about) 2x volume experiment 2 – quantitative relationship required.  
**allow:** ecf from results [1]
- (iii) solution **B** / experiment 1 more concentrated / stronger or converse (1)
- (about) 2x as concentrated – quantitative statement (1) [2]
- (e) half value from table result for experiment 2 (1)  $\text{cm}^3$  (1)
- half volume / amount (of **C**) used (1) [3]
- (f) (both) oxidation (1)  
(and) reduction (occur) (1)
- accept:** answers using definitions of oxidation in terms of:  
oxygen / hydrogen / electrons / oxidation numbers
- transfer of electrons scores 2 [2]
- (g) advantage: easy to use / quick / convenient (1)
- disadvantage: not accurate owtte (1) [2]

Page 3	Mark Scheme	Syllabus	Paper
	IGCSE – May/June 2013	0620	52

- 2 bubbles / fizz (ignore references to colour / ppt) (1) [1]
- (a) pH = 7 (accept any in range 5 to 7, must be a number) (1) [1]
- (b) (i) white precipitate (1)  
dissolves / clears (1) [2]
- (ii) white precipitate (1)  
Insoluble / does not dissolve (1) (dependent on a ppt having been formed) [2]
- (c) no change / colourless solution / no ppt / no reaction (1) [1]
- (d) white (1) precipitate (1) [2]
- (e) bubbles / fizz / effervescence (1)
- limewater (1) milky (1) [3]
- white (1) precipitate (1) [2]
- (f) aluminium (1) sulfate (1) [2]
- (g) carbon dioxide (1) [1]
- (h) calcium (1) carbonate (1) [2]