MARK SCHEME for the October/November 2011 question paper

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for the guidance of teachers

0620 CHEMISTRY

0620/53

Paper 5 (Practical), maximum raw mark 40

This mark scheme is published as an aid to teachers and candidates, to indicate the requirements of the examination. It shows the basis on which Examiners were instructed to award marks. It does not indicate the details of the discussions that took place at an Examiners' meeting before marking began, which would have considered the acceptability of alternative answers.

Mark schemes must be read in conjunction with the question papers and the report on the examination.

• Cambridge will not enter into discussions or correspondence in connection with these mark schemes.

Cambridge is publishing the mark schemes for the October/November 2011 question papers for most IGCSE, GCE Advanced Level and Advanced Subsidiary Level syllabuses and some Ordinary Level syllabuses.



Page 2		2	Mark Scheme: Teachers' version	Syllabus	Paper	
			IGCSE – October/November 2011	0620	53	
l (c	initi fina diff	al rea al rea erenc	results adings completed correctly (1) dings completed correctly (1) all readings to 1 d.p. (ses completed correctly (1) able to supervisors (2)	(1)	[6]	
(d	l) pin	k (1)	to colourless (1) not: clear		[2]	
(e	e) neu	utralis	ation / exothermic (1)		[1]	
(f)) (i)	C/3	smallest B/2 largest (1) one correct = 1		[1]	
	(ii)	orde	er is C/3 A/1 B/2 (2) one correct = 1		[2]	
(g	I) exp	erim	ent 2 is twice the volume of experiment 1 or converse	e (1)	[1]	
(h	i) twi	ce va	lue from table result for experiment 3 (1) cm^3 (1)		[2]	
(i)	use	use a pipette / burette [7				
(j)			none / owtte (1) no change in concentration / temperature has no ef affects speed (1)	fect on quantities	or moles / only [2]	
(k	usi	ng sa	ect method that would work – precise details not nee me method with different acids = 0 s (1) method (1) result (1)	eded	[3]	
	e.g	mea	odium hydroxide add named acid (1) asure temperature change (1) est change = strongest / more concentrated solution	(1)		
	e.g	filter	odium hydroxide add named (excess)metal salt solu - precipitate (1) est mass = strongest / more concentrated solution (1			
					[Total: 21]	
2 (a	ı) (i)	yello	ow / brown / orange (1)		[1]	
	(;;)	whit	e / colourless (1)		[1]	

(ii) white / colourless (1) [1]

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(b) (i) no cł	(b) (i) no change / no reaction owtte (1)					
(ii) white	e (1) precipitate (1)		[2]			
(iii) brow	/n (1) precipitate (1)		[2]			
(iv) brow	vn precipitate (1)		[1]			
(c) (i) solic	white (1) condensation at top of tube (1)					
	water / blue litmus (1) milky / red (1) max 3		[3]			
fizz	/ bubbles / effervescence (1)		[1]			
(ii) fizz /	bubbles / effervescence / brown precipitate (1)		[1]			
(d) iron (1) (III) (1) chloride (1)		[3]			
(e) carbon d	ioxide (1)		[1]			
(f) carbonat	carbonate / hydrogen carbonate (1)					
non trans	sition metal / named metal e.g. sodium (1)		[2]			
			[Total: 19]			