

UNIVERSITY OF CAMBRIDGE INTERNATIONAL EXAMINATIONS International General Certificate of Secondary Education

CHEMISTRY

Paper 1 Multiple Choice

0620/12

May/June 2011

45 Minutes

Additional Materials:

Multiple Choice Answer Sheet Soft clean eraser Soft pencil (type B or HB is recommended)

READ THESE INSTRUCTIONS FIRST

Write in soft pencil.

Do not use staples, paper clips, highlighters, glue or correction fluid.

Write your name, Centre number and candidate number on the Answer Sheet in the spaces provided unless this has been done for you.

There are **forty** questions on this paper. Answer **all** questions. For each question there are four possible answers A, B, C and D.

Choose the **one** you consider correct and record your choice in **soft pencil** on the separate Answer Sheet.

Read the instructions on the Answer Sheet very carefully.

Each correct answer will score one mark. A mark will not be deducted for a wrong answer. Any rough working should be done in this booklet. A copy of the Periodic Table is printed on page 16. You may use a calculator.

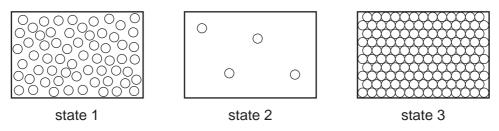
This document consists of 16 printed pages.



UNIVERSITY of CAMBRIDGE International Examinations

[Turn over

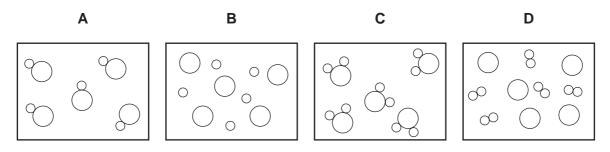
1 The diagrams show the arrangement of particles in three different physical states of substance X.



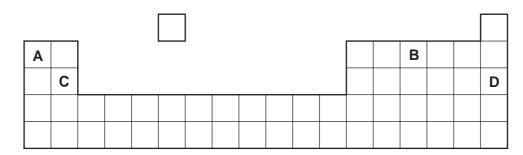
Which statement about the physical states of substance X is correct?

- A Particles in state 1 vibrate about fixed positions.
- **B** State 1 changes to state 2 by diffusion.
- **C** State 2 changes directly to state 3 by condensation.
- **D** The substance in stage 3 has a fixed volume.
- 2 In the diagrams, circles of different sizes represent atoms of different elements.

Which diagram represents hydrogen chloride gas?



3 The diagram shows part of the Periodic Table.



Which element is correctly matched with its electronic structure?

	electronic structure	
Α	A 2,8,1	
в	2,4	
С	2,8,2	
D	2,8	

0620/12/M/J/11

4 An aqueous solution is coloured.

Which method of separation would show that the solution contains ions of different colours?

- **A** chromatography
- B crystallisation
- C distillation
- **D** filtration
- 5 The table gives the solubility of four substances in ethanol and in water.

A mixture containing all four substances is added to ethanol, stirred and filtered.

The solid residue is added to water, stirred and filtered.

The filtrate is evaporated to dryness, leaving a white solid.

Which is the white solid?

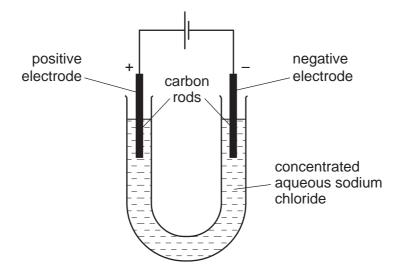
	solub	solubility in		
	ethanol water			
Α	insoluble	insoluble		
в	insoluble	soluble		
С	soluble	insoluble		
D	soluble	soluble		

6 Which two elements react together to form an ionic compound?

		element		electronic stru	ucture	Э	
		W		2,4			
		Х		2,8			
		Y		2,8,1			
		Z		2,8,7			
в	X and	ΙY	С	Y and Z	D	Ζa	and W

A W and X

7 Electricity is passed through concentrated aqueous sodium chloride, as shown.



What is the test for the gas formed at the positive electrode?

- A bleaches damp litmus paper
- B 'pops' with a lighted splint
- **C** relights a glowing splint
- D turns damp red litmus paper blue
- 8 Electricity from a power station passes through overhead cables to a substation and then to a school where it is used to electrolyse concentrated hydrochloric acid using inert electrodes.

Which substances are used for the overhead cables and for the electrodes?

	overhead cables	electrodes
Α	aluminium	copper
В	aluminium	platinum
С	copper	platinum
D	platinum	aluminium

9 The nucleon number and proton number of the lithium atom are shown by the symbol $\frac{7}{3}$ Li.

What is the correct symbol for the lithium ion in lithium chloride?

 $\mathbf{A} \quad {}^{6}_{2} \mathrm{Li}^{-} \qquad \mathbf{B} \quad {}^{6}_{3} \mathrm{Li}^{+} \qquad \mathbf{C} \quad {}^{7}_{3} \mathrm{Li}^{+} \qquad \mathbf{D} \quad {}^{7}_{3} \mathrm{Li}^{-}$

© UCLES 2011

0620/12/M/J/11

10 Three processes are listed.

burning methane in air

radioactive decay of ²³⁵U

reacting hydrogen with oxygen.

Which statements about these processes are correct?

- 1 Hydrogen and methane are being used as fuels.
- 2 All the processes involve oxidation.
- 3 All the processes are used to produce energy.
- **A** 1 and 2 only **B** 1 and 3 only **C** 2 and 3 only **D** 1, 2 and 3
- 11 Which statement about the electrolysis of molten lead(II) bromide is correct?
 - **A** A colourless gas is seen at the cathode.
 - **B** A grey metal is seen at the anode.
 - **C** A red/brown gas is seen at the anode.
 - **D** A red/brown metal is seen at the cathode.
- **12** What is the relative molecular mass (M_r) of HNO₃?
 - **A** 5 **B** 31 **C** 32 **D** 63
- **13** The equation for the effect of heat on hydrated sodium carbonate is as shown.

 $Na_2CO_3.10H_2O(s) \rightleftharpoons Na_2CO_3(s) + 10H_2O(g)$

Statements made by four students about the reaction are given.

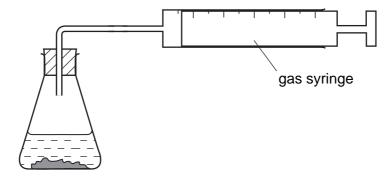
- P Anhydrous sodium carbonate is formed.
- Q Steam is formed.
- **R** There is a colour change from blue to white.
- **S** The reaction is reversible.

Which students' statements are correct?

- A P, Q and R only
- B P, Q and S only
- C Q, R and S only
- D P, Q, R and S

0620/12/M/J/11

14 The apparatus shown can be used to measure the rate of some chemical reactions.

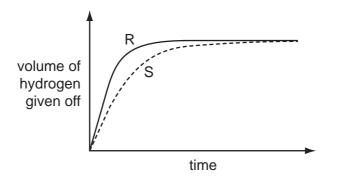


For which two reactions would the apparatus be suitable?

1 and 2 B	1 and 3 C	2 and 4	D	3 and 4
reaction 4	ZnCO ₃ (s) + 2HC <i>l</i> (a	$aq) \rightarrow ZnCl_2(aq)$	+ C	$O_2(g) + H_2O(I)$
reaction 3	$MgO(s) + 2HCl(aq) \rightarrow MgCl_2(aq) + H_2O(I)$			
reaction 2	$2H_2O_2(aq) \rightarrow 2H_2O(I) + O_2(g)$			
reaction 1	AgNO ₃ (aq) + HC1(aq) \rightarrow AgC $l(s)$ +	HN	O₃(aq)

15 A student investigates the rate of reaction between magnesium and excess sulfuric acid.The volume of hydrogen given off in the reaction is measured over time.

The graph shows the results of two experiments, R and S.



Which change in conditions would cause the difference between R and S?

- **A** A catalyst is added in S.
- **B** The acid is more concentrated in R than in S.
- **C** The magnesium is less finely powdered in R than in S.
- **D** The temperature in R is lower than in S.

Α

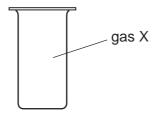
0620/12/M/J/11

16 Butane, ethanol and hydrogen are fuels.

Which substances produce both carbon dioxide and water when used as a fuel?

	butane	ethanol	hydrogen
Α	\checkmark	\checkmark	√
в	\checkmark	\checkmark	x
С	\checkmark	×	1
D	X	\checkmark	x

17 X is a monatomic gas.



Which statement about X is correct?

- A X burns in air.
- **B** X is coloured.
- C X is unreactive.
- **D** X will displace iodine from potassium iodide.
- **18** The equation shows the reaction between a halogen and aqueous bromide ions.

Which words correctly complete gaps 1, 2 and 3?

	1	2	3
Α	chlorine	brown	colourless
в	chlorine	colourless	brown
С	iodine	brown	colourless
D	iodine	colourless	brown

0620/12/M/J/11

19 Carbon dioxide is an acidic oxide that reacts with aqueous calcium hydroxide.

Which type of reaction takes place?

- A decomposition
- **B** fermentation
- **C** neutralisation
- D oxidation
- **20** A solution contains barium ions and silver ions.

What could the anion be?

- A chloride only
- B nitrate only
- C sulfate only
- D chloride or nitrate or sulfate
- **21** A mixture containing two anions was tested and the results are shown below.

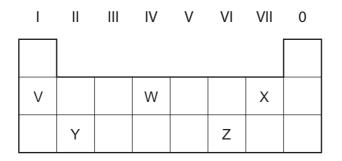
test	result
dilute nitric acid added	effervescence of a gas which turned limewater milky
dilute nitric acid added, followed by aqueous silver nitrate	yellow precipitate formed

Which anions were present?

- A carbonate and chloride
- B carbonate and iodide
- C sulfate and chloride
- D sulfate and iodide
- 22 Which is not a typical property of an acid?
 - A They react with alkalis producing water.
 - **B** They react with all metals producing hydrogen.
 - **C** They react with carbonates producing carbon dioxide.
 - **D** They turn litmus paper red.

0620/12/M/J/11

23 The diagram shows a section of the Periodic Table.



Which elements will conduct electricity at room temperature?

A V, W and X B V, Y and W C W, X and Z D Y and Z

24 Water from a reservoir flows to the water works where purification processes 1 takes place followed by process 2.

What are purification processes 1 and 2?

	purification process 1	purification process 2
Α	chlorination	filtration
В	filtration	chlorination
С	fractional distillation	filtration
D	filtration	fractional distillation

25 The properties of a metal are important in deciding its use.

Which row lists a property that is **not** correct for the use given?

	use of the metal	metal property needed
Α	aluminium in aircraft wings	low density
в	aluminium in food containers	resists corrosion
С	mild steel in car bodies	high density
D	stainless steel in cutlery	does not rust

0620/12/M/J/11

26 Brass is an alloy of copper and zinc.

Which statement is correct?

- A Brass can be represented by a chemical formula.
- **B** Brass is formed by a chemical reaction between copper and zinc.
- **C** The alloy will dissolve completely in dilute hydrochloric acid.
- **D** The zinc in the alloy will dissolve in dilute hydrochloric acid.
- 27 Which statement is correct for the element of proton number 19?
 - A It is a gas that dissolves in water.
 - **B** It is a hard metal that is not very reactive with water.
 - **C** It is a non-metal that burns quickly in air.
 - **D** It is a soft metal that is highly reactive with water.
- 28 Which row describes the conditions used to make steel from the iron produced by a blast furnace?

	calcium oxide (lime)	oxygen	heat
Α	\checkmark	\checkmark	\checkmark
в	\checkmark	\checkmark	x
С	X	\checkmark	\checkmark
D	X	\checkmark	X

0620/12/M/J/11

29 The table shows the results of adding three metals, P, Q and R, to dilute hydrochloric acid and to water.

metal	dilute hydrochloric acid	water
Р	hydrogen produced	hydrogen produced
Q	no reaction	no reaction
R	hydrogen produced	no reaction

What is the order of reactivity of the metals?

	most reactive		least reactive
Α	Р	R	Q
в	Р	Q	R
С	R	Q	Р
D	R	Р	Q

30 Which substance is a metal?

	electrical conductivity (solid)	electrical conductivity (molten)			
Α	high	high			
В	high	low			
С	low	high			
D	low	low			

31 Greenhouse gases may contribute to climate change.

Two of these gases are emitted into the atmosphere as a result of processes within animals.

Gas1..... is produced by process3......

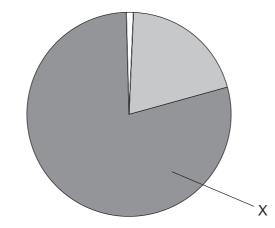
Gas2..... is produced by process4......

Which words correctly complete gaps 1, 2, 3 and 4?

	1	2	3	4
Α	СО	C_2H_6	digestion	respiration
в	CO	C_2H_6	respiration	digestion
С	CO_2	CH_4	digestion	respiration
D	CO ₂	CH_4	respiration	digestion

0620/12/M/J/11

- 12
- 32 The diagram shows the composition by volume of air.



What is X?

- A argon
- B carbon dioxide
- **C** nitrogen
- D oxygen
- **33** The table gives the composition of the atmosphere of four newly discovered planets.

planet	composition of atmosphere
W	argon, carbon dioxide and oxygen
х	argon, nitrogen and oxygen
Y	argon, carbon dioxide and methane
Z	methane, nitrogen and oxygen

On which planets is the greenhouse effect likely to occur?

- A W only
- B W, X and Z
- C W and Y only
- **D** W, Y and Z

0620/12/M/J/11

- **34** Which two substances, when reacted together, would form a salt that contains two of the essential elements provided by fertilisers?
 - A potassium hydroxide and nitric acid
 - B potassium hydroxide and sulfuric acid
 - C sodium hydroxide and nitric acid
 - D sodium hydroxide and sulfuric acid
- **35** Statement 1: Alloying iron with other materials to form stainless steel prevents iron from rusting by excluding oxygen.

Statement 2: Painting, oiling and electroplating are all methods of preventing iron from rusting.

Which is correct?

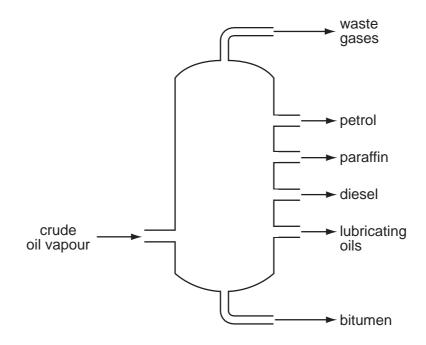
- A Both statements are correct and statement 2 explains statement 1.
- **B** Both statements are correct but statement 2 does not explain statement 1.
- **C** Statement 1 is correct but statement 2 is incorrect.
- **D** Statement 2 is correct but statement 1 is incorrect.
- 36 What is the main constituent of natural gas?
 - A carbon dioxide
 - B ethane
 - C hydrogen
 - D methane
- 37 What is not essential for the formation of ethanol by fermentation?
 - A light
 - B sugar
 - C yeast
 - D water

0620/12/M/J/11

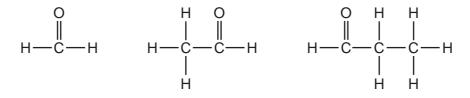
www.XtremePapers.net

[Turn over

38 Which industrial process is shown in the diagram?



- A cracking
- B fermentation
- **C** fractional distillation
- D polymerisation
- **39** The diagram shows the structures of three compounds.



Why do these three compounds belong to the same homologous series?

- **A** They all contain carbon, hydrogen and oxygen.
- **B** They all contain the same functional group.
- **C** They are all carbon based molecules.
- **D** They are all flammable liquids.

0620/12/M/J/11

40 Compounds containing five carbon atoms in a molecule may have names beginning with 'pent...'. What is the name of the compound shown?

key e carbon e oxygen e hydrogen

- A pentane
- B pentanoic acid
- C pentanol
- D pentene

0620/12/M/J/11

	0	4 Helium 2	20 Neon 10 Ar Ar Ar	84 Krypton 36	131 Xenon 54	Radon 86		175 Lu Lutetium 71	Lawrencium 103
	∥>		19 9 Fluorine 35.5 35.5 Chorine	80 Bromine 35	127 I lodine 53	At Astatine 85		173 Yb ^{Ytterbium} 70	Nobelium 102
	⋝		16 8 Oxygen 32 32 16 Sultur	79 Selenium 34	128 Te Tellurium 52	Polonium 84		169 Tm 59	Mendelevium 101
	>		14 Nitrogen 31 Phosphorus 15	75 AS Arsenic 33	122 Sb Antimony 51	209 Bi Bismuth		167 Er Erbium 68	Fermium 100
	2		6 Carbon 6 28 28 14 Silicon	73 Ge Germanium 32	119 Sn 50	207 Pb Lead 82		165 HO Holmium 67	Einsteinium 99
	≡		11 B Boron 5 27 27 Aurminium 13	70 Ga Gallium 31	115 In Indium 49	204 T 1 Thailium 81		162 Dy Dysprosium 66	Cf Californium 98
				65 Zn 30 ^{Zinc}	112 Cd Cadmium 48	201 Hg ^{Mercury} 80		159 Tb ^{Terbium} 65	BK Berkeilum 97
				64 Copper 29	108 Ag Silver 47	197 Au Gold 79		157 Gd Gadolinium 64	Ourrium 96
Group				59 Nickel 28	106 Pd Palladium 46	195 Pt Platinum 78		152 Eu Europium 63	Americium 95
				59 CO 27	103 Rh Rhođium 45	192 I r ^{Iridium}		150 Sm Samarium 62	
		¹ Hydrogen		56 Iron 26	101 Rut Ruthenium 44	190 OS Osmium 76		Promethium 61	Neptunium 93
				55 Mn Manganese 25	Tc Technetium	186 Re Rhenium 75		144 Neodymium 60	238 Uranium 92
				52 Chromium 24	96 Mo Molybdenum 42	184 V Tungsten 74		141 Pr Praseodymium 59	Protactinium 91
				51 Vanadium 23	93 Nabium 41	181 Ta ^{Tantalum} 73		140 Ce ^{Cerium}	232 Th 90
				48 Titanium 22	91 Zr Zirconium 40	178 Hafnium 72		'n	nic mass bol nic) number
				45 Scandium 21	89 Yttrium 39	139 La Lanthanum 57 *	227 Act inium 89	l series eries	a = relative atomic mass X = atomic symbol b = proton (atomic) number
	=		9 Beryllum 4 24 Magnesium 12	40 Calcium 20	88 Sr 38	137 Ba ^{Barium} 56	226 Rad 88	*58-71 Lanthanoid series 190-103 Actinoid series	ت × ۳
	_		7 23 23 23 23 23 23 23 23 23 23 23 23 23	39 K Potassium 19	85 Rb Rubidium 37	133 CS Caesium 55	Fr Francium 87	8-71 L	ه ۲ ۵

Permission to reproduce items where third-party owned material protected by copyright is included has been sought and cleared where possible. Every reasonable effort has been made by the publisher (UCLES) to trace copyright holders, but if any items requiring clearance have unwittingly been included, the publisher will be pleased to make amends at the earliest possible opportunity.

University of Cambridge International Examinations is part of the Cambridge Assessment Group. Cambridge Assessment is the brand name of University of Cambridge Local Examinations Syndicate (UCLES), which is itself a department of the University of Cambridge.

© UCLES 2011

0620/12/M/J/11