

**MARK SCHEME for the May/June 2009 question paper**  
**for the guidance of teachers**

<b>0620/06</b>	<b>0620 CHEMISTRY</b> Paper 6 (Alternative to Practical), maximum raw mark 60
----------------	--

This mark scheme is published as an aid to teachers and candidates, to indicate the requirements of the examination. It shows the basis on which Examiners were instructed to award marks. It does not indicate the details of the discussions that took place at an Examiners' meeting before marking began, which would have considered the acceptability of alternative answers.

Mark schemes must be read in conjunction with the question papers and the report on the examination.

- CIE will not enter into discussions or correspondence in connection with these mark schemes.

CIE is publishing the mark schemes for the May/June 2009 question papers for most IGCSE, GCE Advanced Level and Advanced Subsidiary Level syllabuses and some Ordinary Level syllabuses.

Page 2	Mark Scheme: Teachers' version	Syllabus	Paper
	IGCSE – May/June 2009	0620	06

- 1 (a) balance (1) stirring/(glass) rod/stirrer (1) not thermometer beaker (1) [3]
- (b) (i) excess (1) not residue [1]
- (ii) filtration/decant (1)  
not sieve/strain/centrifuge [1]
- (c) heat/evaporate (1) to crystallising point or description e.g. using glass rod (1) [2]
- 2 (a) to reach room temperature/be at same temperature owtte (1) [1]
- (b) insulator/to minimise heat loss (1) [1]
- (c) exothermic (1) [1]
- (d) (i) 40 cm<sup>3</sup> volume of acid (1) [1]
- (ii) two straight lines, missing error point (1) extended to intersect (1) [2]
- (iii) 22.5 +/- 0.5 (1) or read from graph cm<sup>3</sup> (1) [2]
- 3 (a) add dilute acid (1) fizz, no fizz (1) or correct chloride test [2]
- (b) litmus paper/named indicator (1) turns blue (1) bleached (1) [3]
- (c) sodium hydroxide/ammonia (solution) (1) green (precipitate) (1)  
brown (precipitate) (1) [3]
- 4 (a) Table of results
- |   |    |    |    |    |    |
|---|----|----|----|----|----|
| initial temperature boxes completed correctly (2) | 24 | 33 | 40 | 51 | 60 |
| final temperature boxes correctly completed (2)   | 24 | 31 | 38 | 47 | 54 |
| average temperature boxes correctly completed (1) | 24 | 32 | 39 | 49 | 57 |
- [5]
- (b) 5 points correctly plotted (3), -1 for any incorrect smooth line graph (1) [4]
- (c) (i) experiment 5 (1) [1]
- (ii) more energy owtte (1) particles move faster (1) more kinetic energy = 2  
more collisions (1) [3]

Page 3	Mark Scheme: Teachers' version	Syllabus	Paper
	IGCSE – May/June 2009	0620	06

- (d) idea of a fair test/to compare effect of changing the temperature (1) [1]
- (e) (i) value from graph approx 20 (1) unit (1)  
extrapolation shown (1) [3]
- (ii) curve sketched on grid below original curve (1) [1]
- (f) change e.g. use of data logger/colourimeter (1) or use of lagging/insulation  
/repeat experiments or more values/use a burette or pipette
- explanation e.g. timing of reaction more accurate (1) to reduce heat losses  
/average readings for times/volumes more accurate [2]
- 5 tests on solid S**
- (c) (i) blue precipitate (1) [1]
- (ii) blue (1) precipitate (1) [2]
- dissolves/clears (1) deep/royal blue (1) [2]
- (iii) white (1) precipitate (1) [2]
- (f) (i) V is more reactive or converse (1) [1]
- (ii) oxygen (1) [1]
- (g) catalyst/transition metal/manganese oxide any two points (2)  
V is a better catalyst = 2 [2]
- 6** (a) add water (1)  
crush/mix/warm (1)  
filter/decant or pipette off liquid/sieve (1) [3]
- (b) add indicator solution to acid (and note colour) (1)  
add indicator solution to alkali or named alkali (and note colour) (1) not base  
conclusion e.g. colours should be different owtte (1) [3]