



Key

## CHEMISTRY

## Higher Level

Tuesday 16 November 1999 (afternoon)

Paper 1	Grade	1	2	3	4	5	6	7	
	Boundaries	6-10	11-16	17-21	22-25	26-29	29-31	32 →	1 hour

This examination paper consists of 40 questions.

Each question offers 4 suggested answers.

The maximum mark for this paper is 40.

**INSTRUCTIONS TO CANDIDATES**

Do NOT open this examination paper until instructed to do so.

Answer ALL questions.

For each question, choose the answer you consider to be the best and indicate your choice on the answer sheet provided.

Calculators are NOT permitted for this examination paper.

**EXAMINATION MATERIALS**

Required:  
Optically Mark Read (OMR) answer sheet

Allowed:  
A simple translating dictionary for candidates not working in their own language

Periodic Table

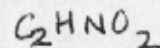
Atomic Number		Atomic Mass	
1	H	1.01	
2	He	4.00	
3	Li	6.94	
4	Be	9.01	
11	Na	22.99	
12	Mg	24.31	
19	K	39.10	
20	Ca	40.08	
21	Sc	44.96	
22	Ti	47.90	
23	V	50.94	
24	Cr	52.00	
25	Mn	54.94	
26	Fe	55.85	
27	Co	58.93	
28	Ni	58.71	
29	Cu	63.55	
30	Zn	65.37	
37	Rb	85.47	
38	Sr	87.62	
39	Y	88.91	
40	Zr	91.22	
41	Nb	92.91	
42	Mo	95.94	
43	Tc	98.91	
44	Ru	101.07	
45	Rh	102.91	
46	Pd	106.42	
47	Ag	107.87	
48	Cd	112.40	
55	Cs	132.91	
56	Ba	137.34	
57 <sup>†</sup>	La	138.91	
87	Fr	(223)	
88	Ra	(226)	
89 <sup>‡</sup>	Ac	(227)	
104	Rf	(261)	
105	Db	(262)	
106	Sg	(263)	
107	Bh	(262)	
108	Hs		
109	Mt		
109	Mt		
72	Hf	178.49	
73	Ta	180.95	
74	W	183.85	
75	Re	186.21	
76	Os	190.21	
77	Ir	192.22	
78	Pt	195.09	
79	Au	196.97	
80	Hg	200.59	
81	Tl	204.37	
82	Pb	207.19	
83	Bi	208.98	
84	Po	(210)	
85	At	(210)	
86	Rn	(222)	
51	Sb	121.75	
52	Te	127.60	
53	I	126.90	
54	Xe	131.30	
50	Sn	118.69	
51	Sb	121.75	
52	Te	127.60	
53	I	126.90	
54	Xe	131.30	
31	Ga	69.72	
32	Ge	72.59	
33	As	74.92	
34	Se	78.96	
35	Br	79.90	
36	Kr	83.80	
13	Al	26.98	
14	Si	28.09	
15	P	30.97	
16	S	32.06	
17	Cl	35.45	
18	Ar	39.95	
5	B	10.81	
6	C	12.01	
7	N	14.01	
8	O	16.00	
9	F	19.00	
10	Ne	20.18	
67	Ho	164.93	
68	Er	167.26	
69	Tm	168.93	
70	Yb	173.04	
71	Lu	174.97	
65	Tb	158.92	
66	Dy	162.50	
64	Gd	157.25	
63	Eu	151.96	
62	Sm	150.35	
61	Pm	146.92	
60	Nd	144.24	
59	Pr	140.91	
58	Ce	140.12	
96	Cm	(247)	
97	Bk	(247)	
98	Cf	(251)	
99	Es	(254)	
100	Fm	(257)	
101	Md	(258)	
102	No	(259)	
103	Lr	(260)	
95	Am	(243)	
94	Pu	(242)	
93	Np	(237)	
92	U	238.03	
91	Pa	231.04	
90	Th	232.04	

†

‡

1. What is the empirical formula for a compound with the molecular formula  $C_6H_3(NO_2)_3$ ?

A. CHNO

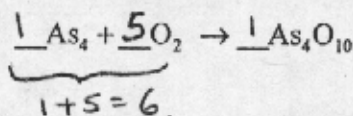


B.  $C_2HNO_2$

C.  $(C_2HNO_2)_3$

D.  $C_6H_3N_3O_6$

2. Arsenic,  $As_4$ , reacts with oxygen to produce the oxide  $As_4O_{10}$ . What is the sum of the coefficients for the reactants in the balanced equation?



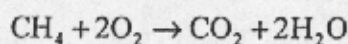
A. 4

B. 5

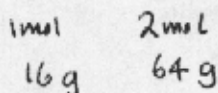
C. 6

D. 7

3. What is the minimum number of grams of  $O_2$  ( $M_R = 32$ ) required to burn 1.6 grams of  $CH_4$  ( $M_R = 16$ ) according to the equation below?



A. 1.6



B. 3.2



C. 6.4

D. 32

4. How many moles of HCl are in  $25 \text{ cm}^3$  of  $0.2 \text{ mol dm}^{-3}$  hydrochloric acid?

A. 0.005

B. 0.008

C. 5

D. 8

$$n = cV = \frac{25}{1000} \text{ dm}^3 \times 0.2 \frac{\text{mol}}{\text{dm}^3} = \frac{5}{1000} \text{ mol} = 0.005 \text{ mol}$$

5. Which species contains 16 protons, 17 neutrons and 18 electrons?

- A.  $^{32}\text{S}^-$       16p 17e<sup>-</sup> x
- (B)**  $^{33}\text{S}^{2-}$       16p, 17n, 18e<sup>-</sup>
- C.  $^{34}\text{S}^-$       16p 17e<sup>-</sup> x
- D.  $^{35}\text{S}^{2-}$       16p 19n 18e<sup>-</sup> x

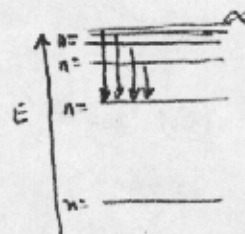
6. Values for the first ionisation energies (IE) for five successive elements in the Periodic Table are given below. Based on these values, which statement is correct?

Element	K	L	M	N	O	P
IE (kJ mol <sup>-1</sup> )	1060	1000	1260	1520	418	?

*← Noble gas      ← Group 1*  
*f      s      d      Ar      He      → Ca*  
*IE increases*

- A. The outermost electron in element K is in a higher energy level than that in element L
  - B. Element M is probably a metal x
  - (C)** Element N is probably a non-metal
  - D. Element P has a lower first ionisation than element O x
7. Spectra have been used to study the arrangements of electrons in atoms. An emission spectrum consists of a series of bright lines that converge at high frequencies. Such emission spectra provide evidence that electrons are moving from

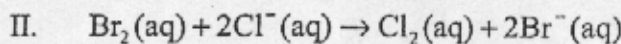
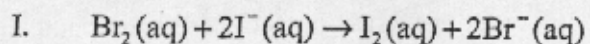
- A. lower to higher energy levels with the higher energy levels being closer together. x
- B. lower to higher energy levels with the lower energy levels being closer together. x
- C. higher to lower energy levels with the lower energy levels being closer together.
- (D)** higher to lower energy levels with the higher energy levels being closer together.



8. Which property of alkali metals changes as stated with increasing atomic number?

- A. Atomic radius decreases x
- B. Reactivity with water decreases x
- C. Electronegativity increases x
- (D)** Melting point decreases      *metals: valence e<sup>s</sup> further from the nucleus*

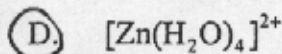
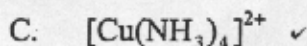
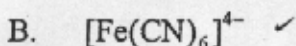
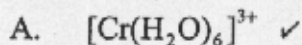
9. Which reaction occurs readily?



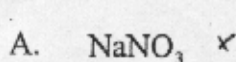
ox. strength  
 $\text{Cl}_2 > \text{Br}_2 > \text{I}_2$   
 $\text{Br}_2$  can displace  $\text{I}_2$  from  $\text{I}^-$   
 $\text{Br}_2$  can NOT displace  $\text{Cl}_2$  from  $\text{Cl}^-$

- (A) I only
- B. II only
- C. Both I and II
- D. Neither I nor II

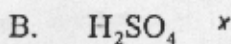
10. Which ion is colourless?



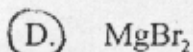
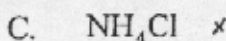
11. Which substance exhibits only ionic bonding?



$\text{Na}^+ \quad \text{NO}_3^-$   
 2 x non-metal  $\Rightarrow$  covalent bonding

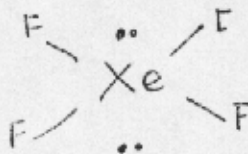
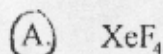


$\text{SO}_4^{2-}$      "  
 $\text{NH}_4^+$         "

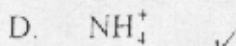
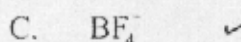
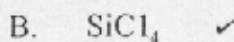


$\text{Mg}^{2+} \quad 2\text{Br}^-$   
 only ionic bonding

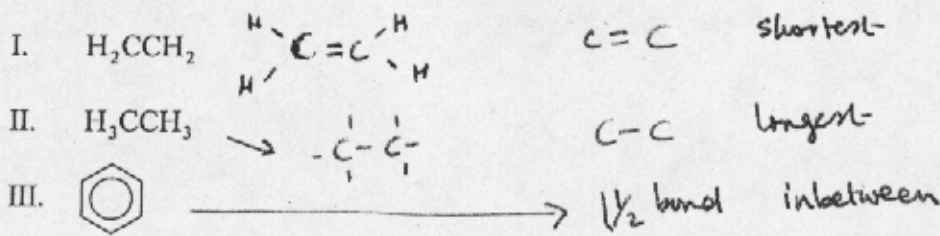
12. Which molecule or ion does not have a tetrahedral shape?



square planar



13. When the substances below are arranged in order of increasing carbon-carbon bond length (shortest bond first), what is the correct order?



Shortest first  
I, III, II

A. I < II < III

B. I < III < II

C. II < I < III

D. III < II < I

14. What type(s) of intermolecular forces is/are present in  $\text{CH}_3\text{OCH}_3$ ?

A. dipole-dipole, hydrogen bonds and van der Waals'

$\rightarrow$  polar molecule, but no H-bonding

B. dipole-dipole and van der Waals' only

C. hydrogen bonds and van der Waals' only

D. van der Waals' only

15. Dry air contains 1 % argon by volume. What is the partial pressure of argon in dry air at one atmosphere pressure (101 kPa)?

A. 1.01 kPa

B. 10.1 kPa

C. 101 kPa

D. 10100 kPa

$$P_x = P_{\text{tot}} \times X_A$$

$$= 101 \text{ kPa} \times \frac{1}{100}$$

$$= 1.01 \text{ kPa}$$

16. The boiling points of four hydrocarbons are given. Which pair will mix most easily at the temperature specified?

	$T_b / K$
cyclohexane	354
cycloheptane	392
cyclooctane	421
cyclononane	444

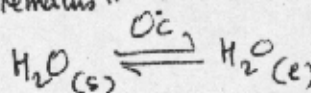
- A. cyclohexane and cycloheptane at 380 K
- B. cycloheptane and cyclooctane at 390 K
- C. cyclooctane and cyclononane at 460 K
- D. cyclononane and cyclohexane at 420 K

temp. at which solid & liquid are at equilibrium with each other.

17. What changes occur when ice at its melting point is converted to liquid water?

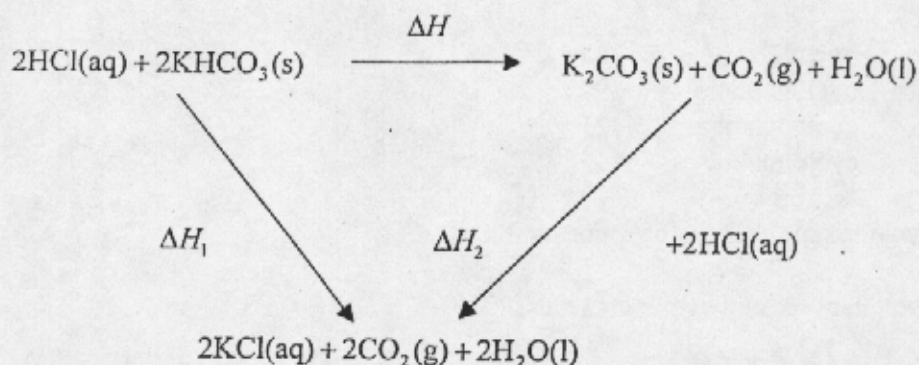
- I. movement of the molecules increases X
- II. distance between molecules increases ✓

temp. remains the same.



- A. I only
- B. II only
- C. Both I and II
- D. Neither I nor II

18.

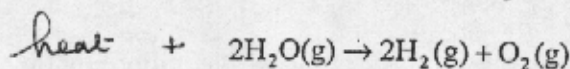


This cycle may be used to determine  $\Delta H$  for the decomposition of potassium hydrogen carbonate. Which expression can be used to calculate  $\Delta H$ ?

- A.  $\Delta H = \Delta H_1 + \Delta H_2$
- (B)**  $\Delta H = \Delta H_1 - \Delta H_2$
- C.  $\Delta H = \frac{1}{2}\Delta H_1 - \Delta H_2$
- D.  $\Delta H = \Delta H_2 - \Delta H_1$
- $\Delta H = \Delta H_1 - \Delta H_2$
19. A sodium hydroxide solution is reacted with excess hydrochloric acid. What information is **not** needed to calculate the molar heat of neutralisation of sodium hydroxide?
- A. Initial temperatures of both solutions ✓
- B. Volumes of both solutions ✓
- (C)** Concentration of the hydrochloric acid solution
- D. Maximum temperature of the mixture ✓
20. Which factor(s) will cause the lattice enthalpy of ionic compounds to increase in magnitude?
- I. an increase in the charge on the ions ✓
- II. an increase in the size of ions ✗
- (A)** I only
- B. II only
- C. Both I and II
- D. Neither I nor II
- Charge ratio  
size  
increase ⇒ higher charge  
or smaller size*



21. How would this reaction at 298 K be described in thermodynamic terms?



(decomposition rxn)

- A. Endothermic ✓ with a significant increase ✓ in entropy
  - B. Endothermic with a significant decrease ✗ in entropy
  - C. Exothermic ✗ with a significant increase in entropy
  - D. Exothermic with a significant decrease ✗ in entropy
22. The rate of reaction of a strip of magnesium and 50 cm<sup>3</sup> of 1 mol dm<sup>-3</sup> HCl is measured at 25° C. In which case would **both** new conditions contribute to an increase in the rate of reaction?

- A. Mg ✓ powder and 100 cm<sup>3</sup> of 1 ✗ mol dm<sup>-3</sup> HCl
- B. Mg ✓ powder and 50 cm<sup>3</sup> of 0.8 ✗ mol dm<sup>-3</sup> HCl
- C. 100 cm<sup>3</sup> of 1 ✗ mol dm<sup>-3</sup> HCl at 30° C ✓
- D. 50 cm<sup>3</sup> of 1.2 ✓ mol dm<sup>-3</sup> HCl at 30° C ✓

23. The rate constant for a certain reaction has the units concentration<sup>-1</sup> time<sup>-1</sup>. What is the order of reaction?

- A. 0
- B. 1
- C. 2
- D. 3

$$\text{rate} = k_2 [A]^2$$

$$\frac{\text{conc}}{\text{time}} = k_2 \text{conc}^2 \quad \therefore k_2 = \frac{1}{\text{conc} \cdot \text{time}} = \text{conc}^{-1} \text{time}^{-1}$$

24. The addition of a catalyst to a chemical reaction alters the rate primarily by

- A. changing the enthalpy of the reaction. ✗
- B. increasing the number of collisions between the reactant molecules in a given time. ✗
- C. increasing the fraction of reactant molecules with a given kinetic energy. ✗
- D. providing a different reaction pathway.

✗ } temp effects

25. Chemical equilibrium is referred to as **dynamic** because, at equilibrium, the

A. equilibrium constant changes.  $\times$

B. reactants and products keep reacting.

*rate of forward rxn = rate of reverse rxn*

C. rates of the forward and backward reactions change.  $\times$

D. concentrations of the reactants and products continue to change.  $\times$

26. An equimolar mixture of propan-1-ol (bp = 97.4° C) and propan-2-ol (bp = 82.4° C) is boiled in a flask to which a distillation column is attached. What is true about the first sample of vapour that enters the distillation column?

*→ lower v.p.r.*

*→ higher v.p.r.*

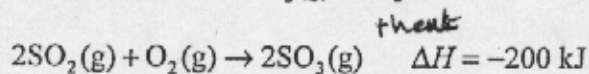
A. It is pure propan-1-ol.  $\times$

B. It is pure propan-2-ol.  $\times$

C. It contains a higher fraction of propan-1-ol than propan-2-ol.  $\times$

D. It contains a higher fraction of propan-2-ol than propan-1-ol.

27. Which change(s) will increase the amount of SO<sub>3</sub>(g) at equilibrium?



*+ heat*

I. Increasing the temperature  $\times$

II. Decreasing the volume  $\checkmark$

III. Adding a catalyst  $\times$

*3 mol → 2 mol*



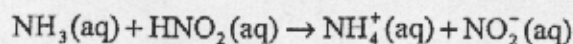
A. I only

B. II only

C. I and III only

D. I, II and III

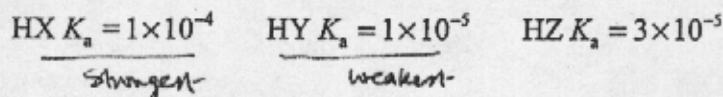
28.



For this reaction, a Brønsted-Lowry acid is proton donor

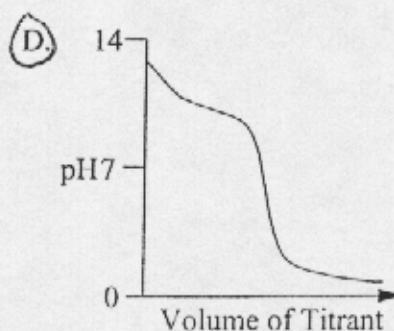
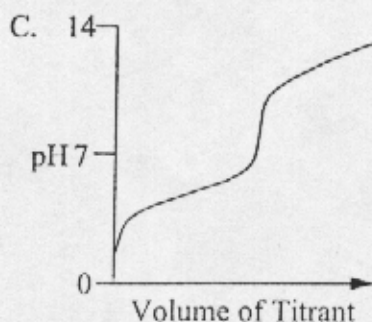
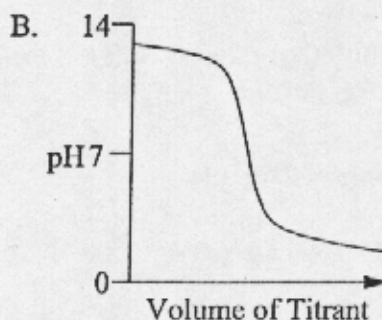
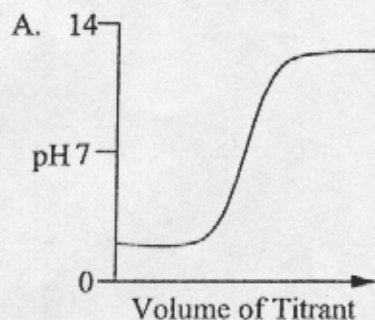
- A.  $\text{NH}_3(\text{aq})$  because it contains the largest number of hydrogen atoms. X
- B.  $\text{NH}_3(\text{aq})$  because it accepts a proton from  $\text{HNO}_2(\text{aq})$ . X
- C.  $\text{HNO}_2(\text{aq})$  because it has lone pairs of electrons on the oxygen atoms. → Lewis Acid
- D.  $\text{HNO}_2(\text{aq})$  because it donates a proton to  $\text{NH}_3(\text{aq})$ . ✓

29. The  $K_a$  values of acids HX, HY and HZ are given. What is the correct order when these acids are arranged in order of increasing strength (weakest first)?



- A.  $\text{HX} < \text{HY} < \text{HZ}$
- B.  $\text{HX} < \text{HZ} < \text{HY}$
- C.  $\text{HZ} < \text{HY} < \text{HX}$
- D.  $\text{HY} < \text{HZ} < \text{HX}$

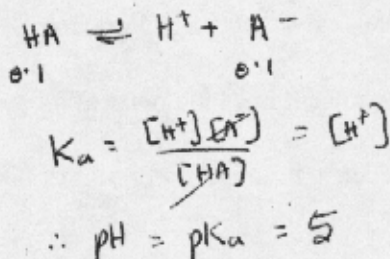
30. Which titration curve represents the titration of a weak base with a strong acid?



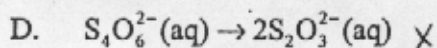
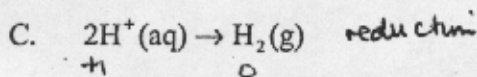
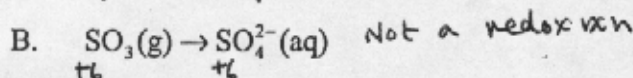
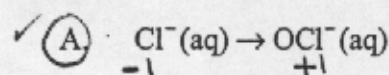
31. What is the pH of a buffer solution that contains  $0.1 \text{ mol dm}^{-3}$  HA and  $0.1 \text{ mol dm}^{-3}$  NaA?

$(HA K_a = 1 \times 10^{-5})$

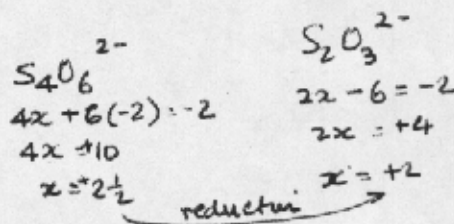
- A. pH = 3
- B. pH = 4
- C. pH = 5
- D. pH = 6



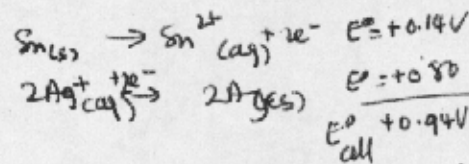
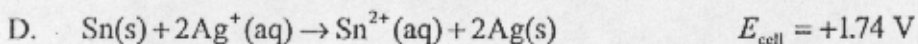
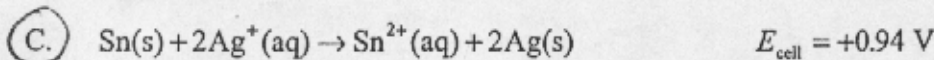
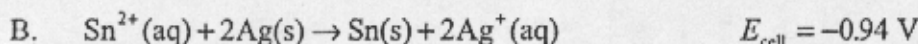
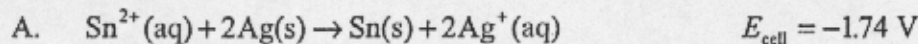
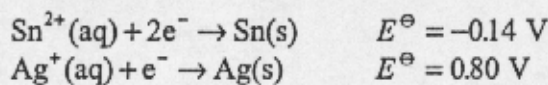
32. For which conversion is an oxidising agent required?



→ (one which is reduced) oxidises another substance



33. The standard electrode potentials for tin and silver are given. What is the equation for the spontaneous reaction together with its cell potential?

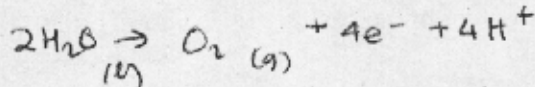


34. What mass of oxygen will be produced when a current of 0.2 A is passed through an aqueous solution of sulphuric acid for 1 hour?

- A.  $\frac{(96500 \times 32)}{(0.2 \times 60 \times 60)}$
- B.  $\frac{(0.2 \times 60 \times 60 \times 32)}{(96500)}$
- C.  $\frac{(0.2 \times 60 \times 60 \times 32)}{(2 \times 96500)}$
- D.  $\frac{(0.2 \times 60 \times 60 \times 32)}{(4 \times 96500)}$

(too many steps involved)

$$C = 0.2 \times 60 \times 60$$

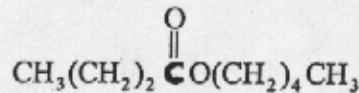


4 x 96500 C per mol of O<sub>2</sub>

$$\therefore \text{amount } n = \frac{0.2 \times 60 \times 60}{4 \times 96500} \text{ mol}$$

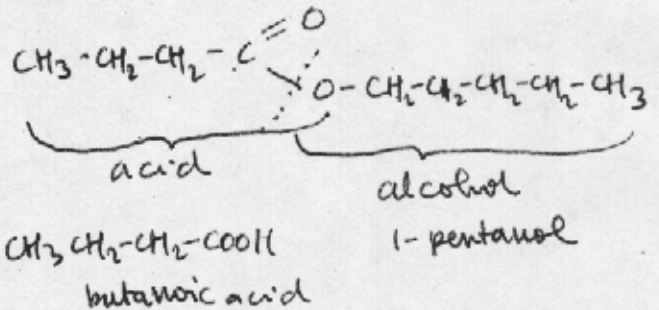
$$\therefore \text{mass} = \frac{0.2 \times 60 \times 60}{4 \times 96500} \text{ mol} \times 32 \text{ g/mol}$$

35.



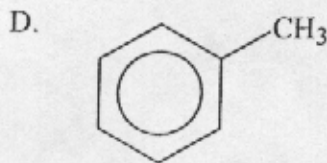
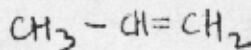
From which two chemicals could this compound be synthesised?

- A. butanoic acid and pentan-1-ol
- B. butanoic acid and butan-1-ol
- C. butanal and pentan-1-ol
- D. pentanoic acid and butan-1-ol



36. Which chemical is most likely to be a starting material for a common polymer?

- A. CH<sub>3</sub>CH<sub>2</sub>CH<sub>3</sub>
- B. CH<sub>3</sub>CH<sub>2</sub>OH
- C. CH<sub>3</sub>CH=CH<sub>2</sub>

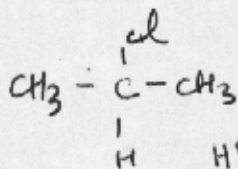


37. 0.1 mol dm<sup>-3</sup> aqueous solutions of these organic compounds were prepared. When these solutions are arranged in order of increasing pH (lowest pH first), what is the correct order?

- I. CH<sub>3</sub>CH<sub>2</sub>CH<sub>2</sub>OH    neutral  
 II. CH<sub>3</sub>CH<sub>2</sub>CH<sub>2</sub>NH<sub>2</sub>    w. base  
 III. CH<sub>3</sub>CH<sub>2</sub>COOH    w. Acid

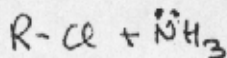
↳ must acidic

- A. I < II < III  
 B. III < I < II  
 C. II < III < I  
 D. III < II < I
38. How many different hydrogen signals would be present in the <sup>1</sup>H NMR spectrum of 2-chloropropane, CH<sub>3</sub>CHClCH<sub>3</sub>?



H's present in two different environments

- A. One  
 B. Two  
 C. Three  
 D. Seven
39. What is the major product when an halogenoalkane (alkyl halide) is reacted with a large excess of ammonia?



- A. An amine  
 B. An amide  
 C. A tetraalkyl ammonium halide  
 D. An alkene
40. What type of reaction does benzene, C<sub>6</sub>H<sub>6</sub>, typically undergo?
- A. Addition  
 B. Elimination  
 C. Reduction  
 D. Substitution