

Tuesday 31 January 2012 – Morning

**GCSE TWENTY FIRST CENTURY SCIENCE
ADDITIONAL SCIENCE A**

A216/02 Unit 2: Modules B5 C5 P5 (Higher Tier)



Candidates answer on the Question Paper.
A calculator may be used for this paper.

OCR supplied materials:

None

Other materials required:

- Pencil
- Ruler (cm/mm)

Duration: 40 minutes



Candidate forename					Candidate surname				
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Centre number						Candidate number			
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INSTRUCTIONS TO CANDIDATES

- Write your name, centre number and candidate number in the boxes above. Please write clearly and in capital letters.
- Use black ink. HB pencil may be used for graphs and diagrams only.
- Answer **all** the questions.
- Read each question carefully. Make sure you know what you have to do before starting your answer.
- Write your answer to each question in the space provided. Additional paper may be used if necessary but you must clearly show your candidate number, centre number and question number(s).
- Do **not** write in the bar codes.

INFORMATION FOR CANDIDATES

- The number of marks is given in brackets [] at the end of each question or part question.
- The total number of marks for this paper is **42**.
- A list of physics equations is printed on page **2**.
- The Periodic Table is printed on the back page.
- This document consists of **20** pages. Any blank pages are indicated.

TWENTY FIRST CENTURY SCIENCE EQUATIONS

Useful Relationships

Explaining Motion

$$\text{speed} = \frac{\text{distance travelled}}{\text{time taken}}$$

$$\text{momentum} = \text{mass} \times \text{velocity}$$

$$\text{change of momentum} = \text{resultant force} \times \text{time for which it acts}$$

$$\text{work done by a force} = \text{force} \times \text{distance moved in the direction of the force}$$

$$\text{change in energy} = \text{work done}$$

$$\text{change in GPE} = \text{weight} \times \text{vertical height difference}$$

$$\text{kinetic energy} = \frac{1}{2} \times \text{mass} \times [\text{velocity}]^2$$

Electric Circuits

$$\text{resistance} = \frac{\text{voltage}}{\text{current}}$$

$$\frac{\text{voltage across primary coil}}{\text{voltage across secondary coil}} = \frac{\text{number of turns in primary coil}}{\text{number of turns in secondary coil}}$$

$$\text{energy transferred} = \text{power} \times \text{time}$$

$$\text{power} = \text{potential difference} \times \text{current}$$

$$\text{efficiency} = \frac{\text{energy usefully transferred}}{\text{total energy supplied}} \times 100\%$$

The Wave Model of Radiation

$$\text{wave speed} = \text{frequency} \times \text{wavelength}$$

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Question 1 begins on page 4

PLEASE DO NOT WRITE ON THIS PAGE

Answer **all** the questions.

- 1 In 2010 a volcano in Iceland erupted.

It sent clouds of sulfur dioxide gas and volcanic ash into the atmosphere.

Airports closed because planes could not fly through the ash from the volcano.

- (a) Sulfur dioxide has covalent bonds.

Put a tick () in each set of boxes to complete the sentence to explain how a covalent bond is made.

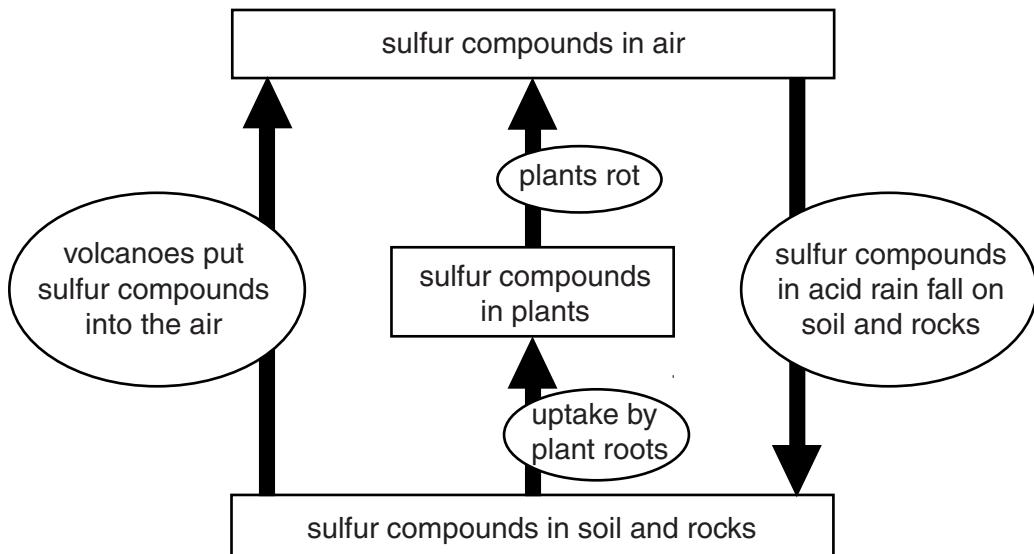
A covalent bond is made when

electrons	<input type="checkbox"/>	are	all in one atom.	<input type="checkbox"/>
ions	<input type="checkbox"/>		shared between two atoms.	<input type="checkbox"/>
neutrons	<input type="checkbox"/>		spread out between many atoms.	<input type="checkbox"/>
protons	<input type="checkbox"/>		transferred from one atom to another.	<input type="checkbox"/>

[2]

- (b) Volcanic eruptions are not the only way that sulfur compounds get into the air.

The diagram shows how some sulfur compounds moved into and out of the air over a period of 4 weeks.



Over the 4 week period, the mass of sulfur compounds in the air did not change overall.

However, the mass of sulfur compounds moving into the air, from volcanoes and rotting plants, was **not** the same as the mass of sulfur compounds moving out of the air as acid rain.

Put a tick (✓) in the box next to the **best** reason for this.

Rainfall varied over the 4-week period.

Volcanoes convert sulfur in rocks into other elements.

Not all processes are shown on the diagram.

Sulfur dioxide molecules are not as heavy as the sulfur-containing molecules taken up by plant roots.

[1]

(c) The volcanic ash also contained silicon dioxide.

Silicon dioxide has a high melting point.

Explain why silicon dioxide has a high melting point.

Use your understanding of the structure and bonding of silicon dioxide in your answer.

.....
.....
.....
.....
.....
.....
.....

[3]

[Total: 6]

- 2 Iron is extracted by heating iron oxide with carbon.

During this reaction the carbon takes the oxygen away from the iron oxide.

- (a) What do we call a reaction in which a metal oxide loses oxygen?

Put a (ring) around the correct answer.

electrolysis

extraction

oxidation

reduction

[1]

- (b) The formula of one oxide of iron is Fe_2O_3 .

What mass of iron is contained in 160g of Fe_2O_3 ?

[relative atomic masses: O = 16, Fe = 56]

Put a (ring) around the correct answer.

48 g

56 g

64 g

112 g

[1]

- (c) Put numbers in the boxes to balance the equation for the formation of iron from iron oxide.



[2]

- (d) Other metals are also extracted by heating their oxides with carbon.

Put ticks (✓) in the boxes next to **two** metals that are extracted in this way.

copper

magnesium

potassium

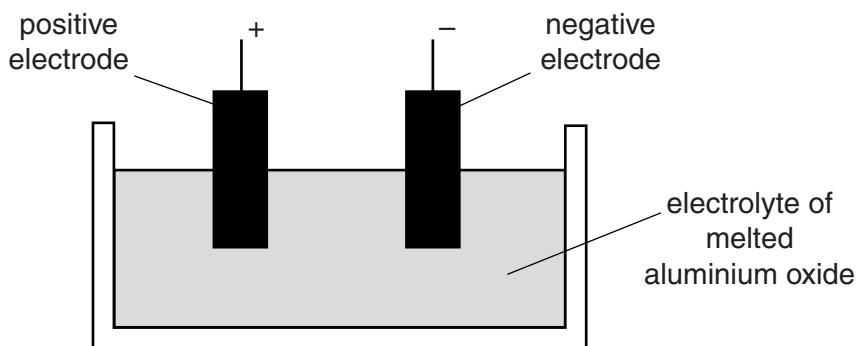
sodium

zinc

[2]

[Total: 6]

- 3 Aluminium is made when melted aluminium oxide is electrolysed.



The aluminium oxide, Al_2O_3 , breaks down into aluminium and oxygen.

Here is part of a description of what happens during electrolysis.

Put a (ring) around the correct **word** in each pair to complete each sentence.

Oxygen ions have a **negative / positive** charge.

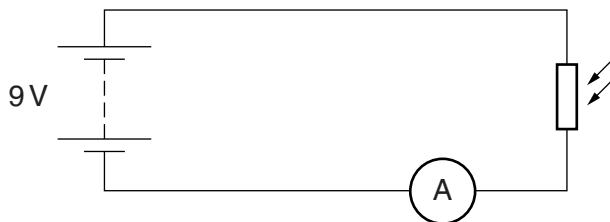
They move towards the **negative / positive** electrode where they **lose / gain** some **electrons / protons** and turn into neutral **atoms / molecules**.

These then combine to form **atoms / molecules** of oxygen gas.

[2]

[Total: 2]

- 4 Alan builds this circuit.



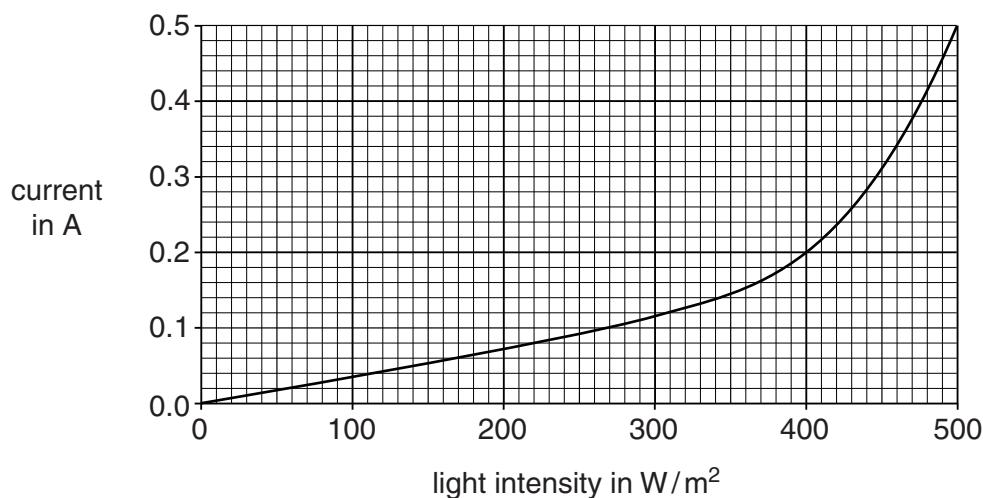
- (a) Alan uses a voltmeter to check that the potential difference across the battery is 9V.

Draw on the circuit diagram to show how the voltmeter should be connected.

[1]

- (b) Alan measures the current in the LDR at different light intensities.

Here is a graph of his results.



- (i) Use data from the graph to calculate the resistance of the LDR at a light intensity of 400W/m^2 . The potential difference across the LDR is 9V.

$$\text{resistance} = \dots \Omega \quad [1]$$

- (ii) The pattern of the graph can be explained by the sentences below.

Complete the sentences. Choose words from this list.

You may use each word once, more than once or not at all.

decreases

increases

stays the same

As the light intensity increases, the current

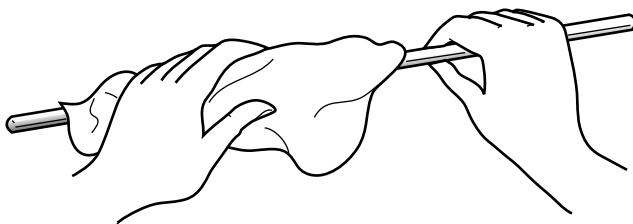
This is because the number of charges that are free to move

Therefore the resistance of the LDR

[2]

[Total: 4]

- 5 Sal rubs a glass rod with a silk cloth.



This gives the rod a **negative** charge.

- (a) Here are some statements about the rod and the cloth.

Put ticks (\checkmark) in the boxes next to the **two** correct statements.

Electrons transfer from the cloth to the rod.

Negative atoms transfer from the cloth to the rod.

Positive electrons transfer from the rod to the cloth.

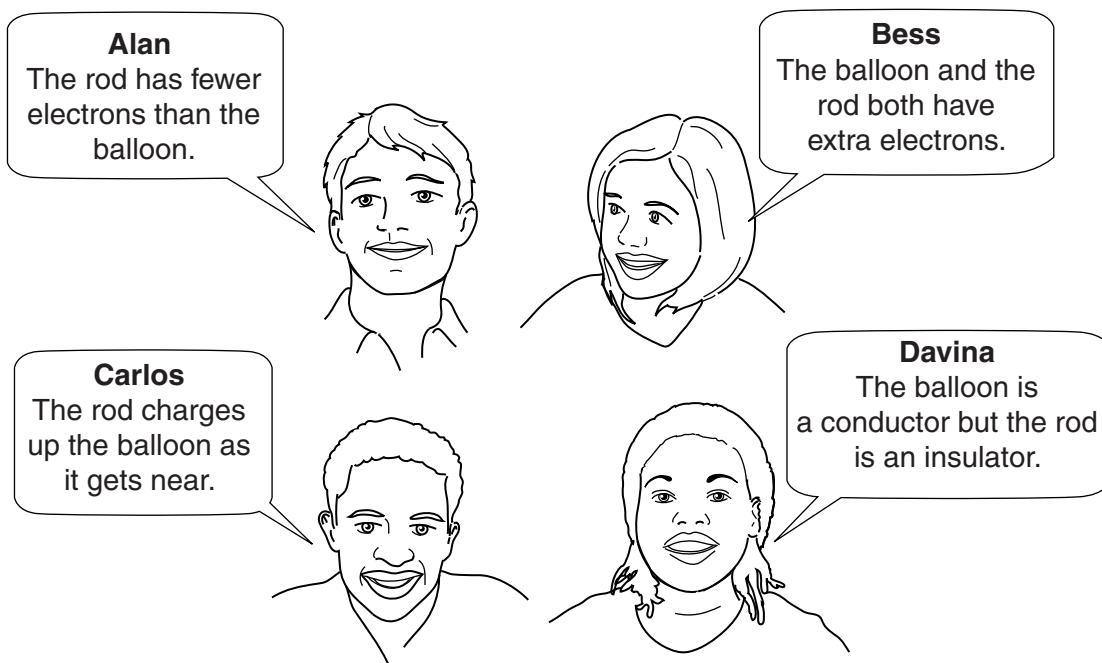
The cloth gains a positive charge as it loses electrons.

The rod and cloth both end up with the same negative charge.

[2]

- (b) Sal holds the charged rod near to a balloon. The rod **repels** the balloon.

Her friends discuss why this happens.

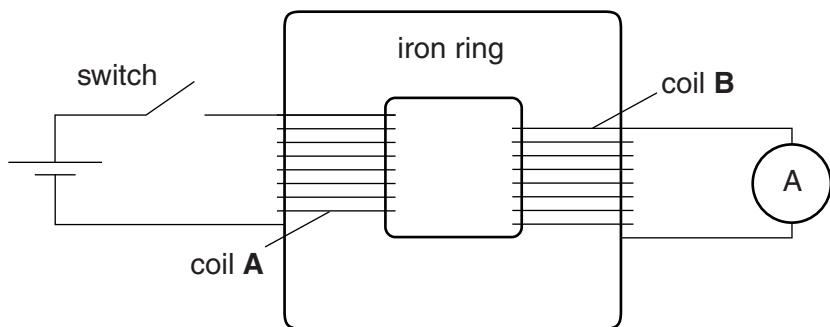


Who correctly explains why the rod repels the balloon?

answer [1]

[Total: 3]

- 6 Zabu builds this circuit with two coils of insulated copper wire and an iron ring.



- (a) Here are some properties of the materials used in the circuit.

Choose the correct word from this list to complete the statement.

a conductor

an insulator

magnetic

a solid

The ring is made of iron because iron is

[1]

- (b) She closes the switch.

The ammeter shows that there is a **brief** pulse of current in coil B.

Explain why there is a brief pulse of current.

Your answer should include what happens

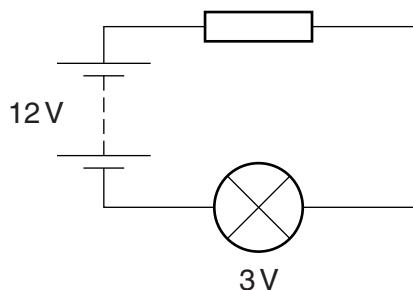
- in coil A
- in the iron ring
- in coil B.

[3]

[Total: 4]

- 7 Celina builds this circuit.

She uses a battery, wires, a resistor and a lamp.



- (a) The current in the lamp is 0.5A when the potential difference across it is 3V.

Calculate the power of the lamp.

$$\text{power} = \dots \text{W} [1]$$

- (b) The potential difference across the battery is 12V.

The resistor ensures that the potential difference across the lamp is 3V.

Suggest how the resistor does this.

.....
..... [1]

- (c) The potential difference across the lamp is 3V.

The potential difference across the battery is 12V.

What is the resistance of the resistor if the current in the lamp is 0.5A?

Put a ring around the correct answer.

6Ω

18Ω

24Ω

30Ω

[1]

[Total: 3]

- 8 Roy is a farmer.

He breeds pigs.

The sperm from his male pig is used to fertilise a female pig.

- (a) Each sperm cell from the male pig contains 19 chromosomes.

How many chromosomes are in each unfertilised egg cell of the female pig?

Put a (ring) around the correct answer.

9

19

36

38

54

[1]

- (b) Once the sperm fertilises the egg, a zygote is formed.

How many chromosomes are in the zygote nucleus?

Put a (ring) around the correct answer.

9

19

36

38

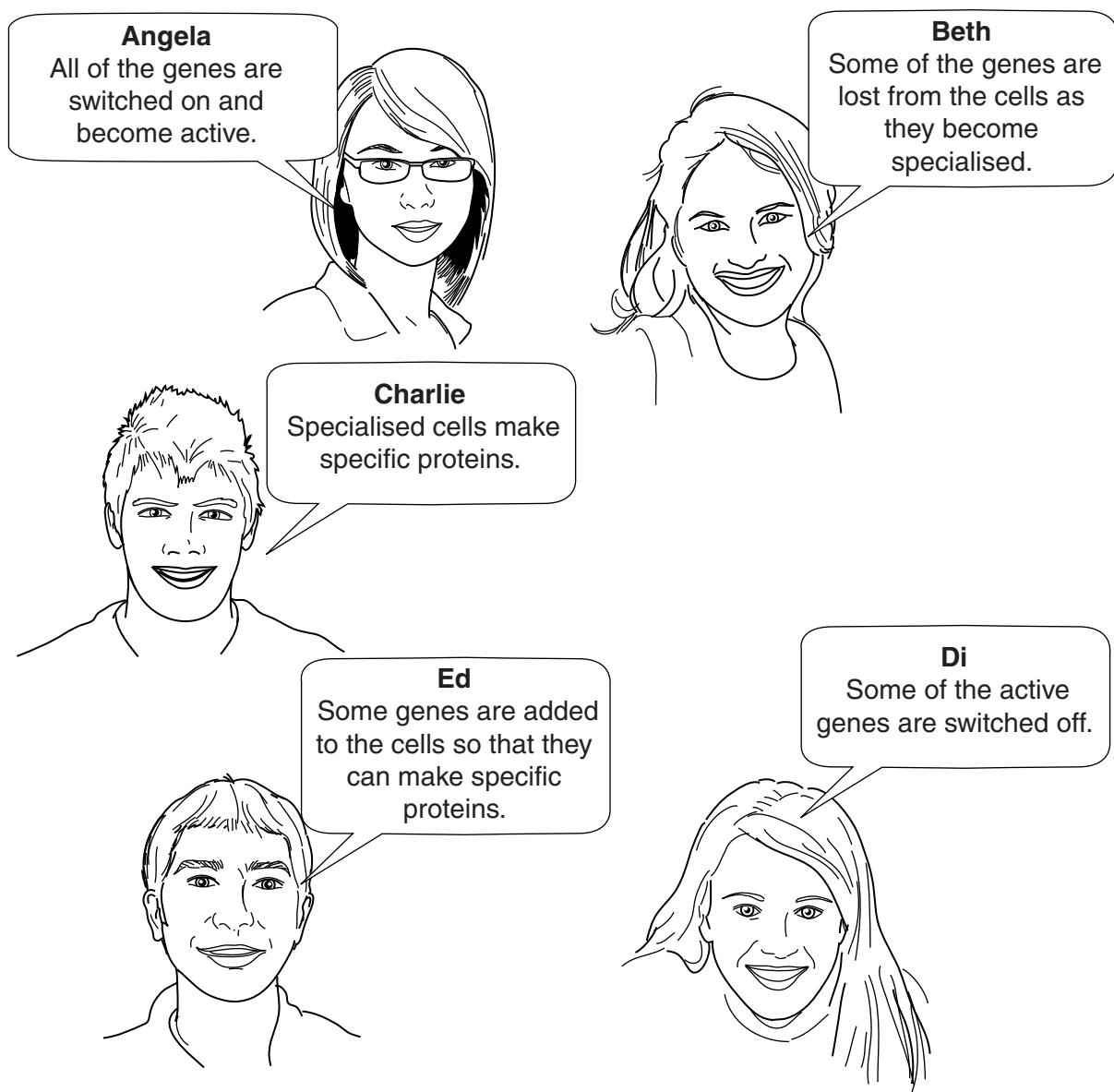
54

[1]

(c) The zygote develops into an embryo.

As the embryo develops, the cells become specialised and form different tissues.

Five people suggest how this can happen.



Which **two** people give correct explanations?

answer and [2]

[Total: 4]

- 9 Red blood cells produce a protein called haemoglobin.

Mature red blood cells do not have a nucleus.

Put ticks (**✓**) in the boxes next to the **two** correct statements about red blood cells.

Red blood cells make haemoglobin all of the time.

Red blood cells have a nucleus at an early stage.

Red blood cells collect haemoglobin in the lungs.

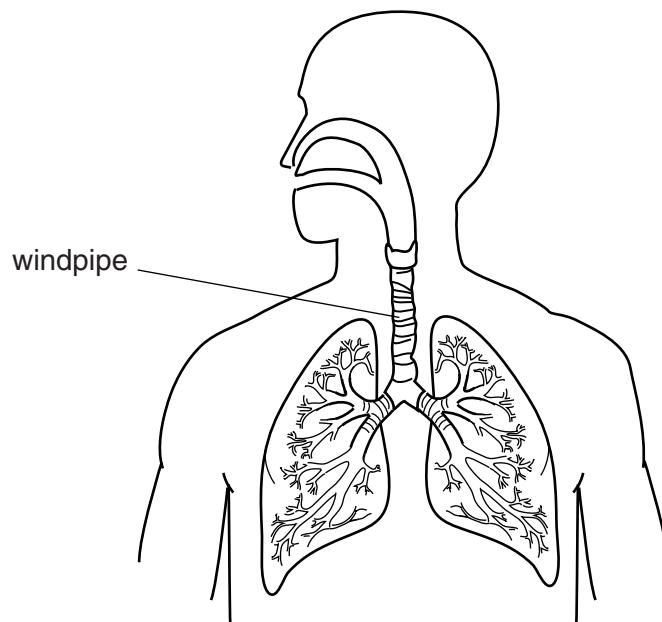
Red blood cells only have a nucleus at the end of their life.

Red blood cells make haemoglobin only when they have a nucleus.

[2]

[Total: 2]

- 10 Doctors can replace a patient's damaged windpipe.



A windpipe is collected from a dead donor.

All of the donor's cells are removed to leave a protein framework.

Stem cells from the patient's bone marrow are placed in the framework.

Once the stem cells have developed, the windpipe is transplanted into the patient.

Explain why stem cells and not muscle cells are used to grow the new cells for the windpipe.

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[3]

[Total: 3]

11 A gene is a length of DNA.

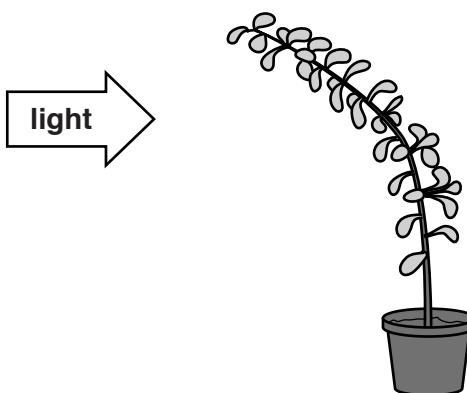
How do different genes make different proteins?

.....
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.....
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.....
.....
..... [3]

[Total: 3]

12 A plant is grown in a pot.

The shoot grows towards the light.



Put a tick (✓) in the box next to the correct word to complete each sentence.

This curve is caused by a change in

auxin	
meristem	
xylem	

concentration.

There is an increase in concentration on the

lit	
shaded	
leafy	

side of the shoot tip.

Its effect is to make cells on this side become

longer.	
greener.	
smaller.	

[2]

[Total: 2]

END OF QUESTION PAPER

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The Periodic Table of the Elements

1 2

1	H	hydrogen	1
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relative atomic mass atomic symbol name atomic (proton) number

7	Li	lithium	3
9	Be	beryllium	4
23	Na	sodium	11
39	Ca	calcium	20
85	Rb	rubidium	37
133	Cs	caesium	55
[223]	Fr	francium	87

3 4

7	Li	lithium	3
9	Be	beryllium	4
23	Na	sodium	11
39	Ca	calcium	20
85	Sr	strontium	38
133	Ba	barium	56
[226]	Ra	radium	88
17	Sc	scandium	21
89	Y	yttrium	39
137	La*	lanthanum	57
[227]	Ac*	actinium	89
45	Ti	titanium	22
89	Zr	zirconium	40
139	Hf	hafnium	72
178	Ta	tantalum	73
181	W	tungsten	74
184	Re	rhenium	75
186	Os	osmium	76
190	Ir	iridium	77
192	Pt	platinum	78
195	Au	gold	79
197	Hg	mercury	80
201	Tl	thallium	81
204	Pb	lead	82
207	Bi	bismuth	83
209	Po	polonium	84
[271]	Ds	darmstadtium	110
[268]	Hs	hassium	108
[277]	Bh	bohrium	107
[266]	Sg	seaborgium	106
[262]	Db	dubnium	105
104			
109			
110			
111			

20

5 6 7 0

1	H	hydrogen	1
4	He	helium	2
11	B	boron	5
12	C	carbon	6
14	N	nitrogen	7
16	O	oxygen	8
19	F	fluorine	9
20	Ne	neon	10
27	Al	aluminium	13
28	Si	silicon	14
31	P	phosphorus	15
32	S	sulfur	16
35.5	Cl	chlorine	17
40	Ar	argon	18

Elements with atomic numbers 112-116 have been reported but not fully authenticated