

GCSE 2004

June Series



Mark Scheme

Science: Single Award

Specification B (Co-ordinated)

3463/2F

Mark schemes are prepared by the Principal Examiner and considered, together with the relevant questions, by a panel of subject teachers. This mark scheme includes any amendments made at the standardisation meeting attended by all examiners and is the scheme which was used by them in this examination. The standardisation meeting ensures that the mark scheme covers the candidates' responses to questions and that every examiner understands and applies it in the same correct way. As preparation for the standardisation meeting each examiner analyses a number of candidates' scripts: alternative answers not already covered by the mark scheme are discussed at the meeting and legislated for. If, after this meeting, examiners encounter unusual answers which have not been discussed at the meeting they are required to refer these to the Principal Examiner.

It must be stressed that a mark scheme is a working document, in many cases further developed and expanded on the basis of candidates' reactions to a particular paper. Assumptions about future mark schemes on the basis of one year's document should be avoided; whilst the guiding principles of assessment remain constant, details will change, depending on the content of a particular examination paper.

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SCIENCE: SINGLE AWARD CO-ORDINATED

INFORMATION FOR EXAMINERS

1. General

The mark scheme for each question shows:

- the marks available for each part of the question;
- the total marks available for the question;
- the typical answer or answers which are expected;
- extra information to help the Examiner make his or her judgement and help to delineate what is acceptable or not worthy of credit or, in discursive answers, to give an overview of the area in which a mark or marks may be awarded.

The extra information is aligned to the appropriate answer in the left-hand part of the mark scheme and should only be applied to that item in the mark scheme.

At the beginning of a part of a question a reminder may be given, for example:
where consequential marking needs to be considered in a calculation;
or the answer may be on the diagram or at a different place on the script.

In general the right hand side of the mark scheme is there to provide those extra details which confuse the main part of the mark scheme yet may be helpful in ensuring that marking is straightforward and consistent.

2. Emboldening

- 2.1** In a list of acceptable answers where more than one mark is available ‘any **two** from’ is used, with the number of marks emboldened. Each of the following lines is a potential mark.
- 2.2** A bold **and** is used to indicate that both parts of the answer are required to award the mark.
- 2.3** Alternative answers acceptable for a mark are indicated by the use of **or**. (Different terms in the mark scheme are shown by a / ; e.g. allow smooth / free movement.)

3. Marking points

3.1 Marking of Quality of Written Communication

Where *Quality of written communication* appears in the mark scheme, one mark is to be awarded for either of the following points:

- Using correct scientific terms
- Correct sequencing or linking of ideas or points

The mark scheme will specify which of the points is to be awarded in a particular question. A QoWC mark can be awarded for a scientific answer, even if it is not accurate. It cannot be awarded for a nonsensical or non-scientific answer.

On the script, the QoWC tick should be identified by a ‘q’ written next to it.

3.2 Marking of lists

This applies to questions requiring a set number of responses, but for which candidates have provided extra responses. The general principle to be followed in such a situation is that ‘right + wrong = wrong’.

Each error/contradiction negates each correct response. So, if the number of error/contradictions equals or exceeds the number of marks available for the question, no marks can be awarded.

However, responses considered to be neutral (indicated as * in example 1) are not penalised.

Example 1: What is the pH of an acidic solution? (1 mark)

Candidate	Response	Marks awarded
1	4,8	0
2	green, 5	0
3	red*, 5	1
4	red*, 8	0

Example 2: Name two planets in the solar system. (2 marks)

Candidate	Response	Marks awarded
1	Pluto, Mars, Moon	1
2	Pluto, Sun, Mars, Moon	0

3.3 Use of chemical symbols/formulae

If a candidate writes a chemical symbol/formula instead of a required chemical name, full credit can be given if the symbol/formula is correct and if, in the context of the question, such action is appropriate.

3.4 The marking of quantitative relationships

Full credit can be given for a correct quantitative relationship expressed in:

- named units;
- physical quantities;
- standard symbols;
- a combination of physical quantities and units.

No credit can be given for any quantitative relationship expressed in terms of:

- a combination of physical quantities, units and symbols;
- a diagram, e.g. the ohm's law triangle, unless the rest of the answer shows clearly that the candidate understands the relationships involved.

3.5 Marking procedure for calculations

3.5.1 Full marks can be given for a correct numerical answer, as shown in the column 'answers', without any working shown. However:

- if the answer is incorrect, mark(s) can be gained by correct substitution/working and this is shown in the 'extra information' column;
- if the answer is correct, but an incorrect relationship is written in the working, then no marks can be awarded (see 3.5.2).

3.5.2 Where calculations are based on incorrectly recalled relationships, neither the incorrectly recalled relationship, nor the resulting calculation based on the incorrect relationship, will be credited.

3.6 Interpretation of 'it'

Answers using the word 'it' should be given credit only if it is clear that the 'it' refers to the correct subject.

3.7 Errors carried forward

There should be no error carried forward from a previous answer which has been based on wrong science. Any error in the answers to a structured question should be penalised once only.

Examples

- (a) A candidate who calculates average speed using $\text{speed} = \text{time}/\text{distance}$ **and** then proceeds to use this incorrect answer to calculate an acceleration based on the correct quantitative relationship should be given credit for the use of the correct acceleration relationship but none for either numerical answer.
- (b) A candidate who incorrectly calculates average speed using $\text{speed} = \text{distance}/\text{time}$ and then proceeds to use this incorrect value to calculate an acceleration based on the correct quantitative relationship, should be given credit for the use of both correct quantitative relationships **and** for the correct substitution and use of the incorrect value in the calculation of the rate of acceleration.

Papers should be constructed in such a way that the number of times errors can be carried forward are kept to a minimum. Allowances for errors carried forward are most likely to be restricted to calculation questions and should be shown by the abbreviation e.c.f. in the marking scheme.

3.8 Phonetic spelling

The phonetic spelling of correct scientific terminology should be credited **unless** there is a possible confusion with another technical term.

3.9 Brackets

(.....) is used to indicate information which is not essential for the mark to be awarded but is included to help the examiner identify the sense of the answer required.

3.10 Interpretation of marginal points

There will be times when the answer is almost, but not quite, correct. Some examiners would award a mark while others would not. In any one script, an attempt should be made to balance these nearly correct answers by giving the mark on some occasions but not on others. If this is not done, the marking would end up being too lenient or too harsh.

3.11 Unexpected Correct Answers not in the Mark Scheme

The Examiner should use professional judgement to award credit where a candidate has given an unexpected correct answer which is not covered by the mark scheme. The Examiner should consult with the Team Leader to confirm the judgement. The Team Leader should pass this answer on to the Principal Examiner with a view to informing all examiners.



**Single Award
Foundation Tier 3463/2F**

3463/2F Q1

question	answers	extra information	mark
	hydrogen / H ₂ / H	accept correct symbol or formula	1
	nickel / Ni		1
	astatine / At / At ₂		1
	bromine / Br / Br ₂		1
	lithium / Li		1
total			5

3463/2F Q2

question	answers	extra information	mark
(a)	carbon	accept C	1
(b)	protons		1
total			2

3463/2F Q3

question	answers	extra information	mark
(a)	it can provide oxygen which can make other substances burn more fiercely	list principle applies	1
	it can cause reddening or blistering of the skin		1
(b)	gloves		1
	to prevent contact with skin / hands	reason must be linked to clothing	1
	or laboratory coat to prevent contact with skin / clothes etc. or goggles / safety glasses / visor to prevent contact with eyes / splashes etc.	accept any other item of safety clothing e.g. mask for 1 mark only	
(c)(i)	measure volume / mass of gas produced	1 mark is for a sensible way of measuring the amount of product produced and 1 mark is for the idea of timing (measuring the rate at which bubbles are produced e.g. number of bubbles in 30 seconds gains only 1 mark unless an enclosed system is used)	1
	in a certain time period		1
	e.g. measure volume of gas produced at regular time intervals or time taken to fill a test tube with the gas or collect a certain volume of gas or measure decrease in mass of flask and contents at regular time intervals or time taken for the mass to decrease by certain amount		
(c)(ii)	increases rate (owtte)		1
(c)(iii)	change the concentration or add a catalyst or change the surface area or <u>lower</u> the temperature	accept 'expose to sunlight' (owtte) or change the amount of water / powder / solution used ignore 'stirring'	1
total			8

3463/2F Q4

question	answers	extra information	mark
(a)	contains nitrogen and potassium	list principle – must be names	1
(b)	600 g or 0.6 kg	accept 600 with no unit but incorrect unit loses the answer mark if 0.6 is given as the answer then there must be a unit $1000 \times \left(\frac{60}{100} \right)$ for 1 mark	2
(c)	nitric acid		1
	potassium hydroxide		1
	water		1
total			6

3463/2F Q5

question	answers	extra information	mark
(a)	bacteria	list principle applies	1
(b)	lactic	list principle applies	1
total			2

3463/2F Q6

question	answers	extra information	mark
(a)(i)	coke	allow coal ignore carbon	1
(ii)	any plausible description of a difference based on diagram or their knowledge e.g. use of water power or smaller scale or use of bellows or use of cold air or slower speed / less efficient	accept converse argument but must be clear	1
(b)(i)	to remove / lose oxygen (owtte)	accept answer in terms of electrons (gain electrons) but not hydrogen	1
(ii)	iron and carbon dioxide	both for 1 mark accept Fe and CO ₂ do not accept Fe and CO ₂	1
(c)	to remove impurities / acidic compounds	accept to form slag (owtte) accept correct chemical description suggesting impurities allow to stop furnace getting blocked	1
(d)	sodium is <u>higher</u> in the reactivity series than carbon	it = sodium accept sodium is (very) high in the reactivity series accept sodium is <u>too</u> reactive accept sodium is more reactive than iron / calcium sodium is reactive = 0 marks must be a comparison with carbon for both marks	1 1
(e)(i)	chromium or nickel	accept Cr or Ni	1
(ii)	does not rust / corrode or does not react with water or does not oxidise	ignore references to heat conductivity / conduction or appearance (e.g. tarnish) ignore melting point any reference to electrical conductivity = 0 marks	1
total			9

3463/2F Q7

question	answers	extra information	mark
(a)	catalyst or speeds up the reaction (owtte)	accept lowers activation energy not just helps reaction to take place ignore increased surface area	1
(b)	C ₈ H ₁₈	allow H ₁₈ C ₈ must be upper case do not accept powers	1
(c)	B because it is a gas or because it has small molecules or because <u>they</u> are small	position and reason for mark allow it has a lower / <u>very</u> low boiling point than A ignore references to solubility accept does not condense do not accept light molecules or bubbles into B do not accept <u>it</u> is small	1
(d)	breakdown of a substance (owtte) by the action of heat (owtte)	do not accept decompose unqualified	1 1
(e)	Quality of written communication	if the written communication makes sense and it is in context then award Q mark Q ✓ Q X	1
	large to small molecules or scientific word that implies smaller, e.g. alkene / ethane / petrol	any name or formula of alkane / alkene smaller than decane	1
	either advantages of smaller molecules or disadvantages of larger molecules e.g. hydrocarbons with large molecules are limited in their usefulness	or converse for smaller molecules	1
	large hydrocarbon molecules do not ignite easily / do not flow easily / are not very volatile	or converse for smaller molecules	
	more large hydrocarbon molecules are produced than are needed	or converse for smaller molecules	
	smaller molecules are useful as fuels alkenes / products can be used to make polymers		
(f)	(fractional) distillation	accept fractionation accept good description do not accept just diagram	1
total			9

3463/2F Q8

question	answers	extra information	mark
(a)(i)	argon and potassium or tellurium and iodine or cobalt and nickel	accept correct symbols allow argon and calcium	1
(ii)	it would place them in incorrect <u>groups</u> (owtte)	idea of <u>not</u> being placed with elements which have similar properties or converse argument accept would not have same number of electrons in outer shell allow it would put them in wrong period if linked to argon and potassium / calcium do not accept reasons based just on protons do not accept metals and non-metals mixed up	1
(b)	any two from: • each successive atom has one more electron • atoms in the same group have the same number of electrons in their <u>outer shells</u> / energy levels • across a period an energy level / shell is being filled • in the next period the next energy level / shell is being filled	number of electrons in outer shell = group number accept period number = shell number	2
total			4