

Candidate Forename		Candidate Surname	
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Centre Number						Candidate Number				
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**OXFORD CAMBRIDGE AND RSA EXAMINATIONS
GENERAL CERTIFICATE OF SECONDARY EDUCATION**

A322/01

TWENTY FIRST CENTURY SCIENCE

CHEMISTRY A

**UNIT 2: Modules C4 C5 C6
(Foundation Tier)**

MONDAY 28 JUNE 2010: Morning

DURATION: 40 minutes

SUITABLE FOR VISUALLY IMPAIRED CANDIDATES

Candidates answer on the Question Paper

A calculator may be used for this paper

OCR SUPPLIED MATERIALS:

None

OTHER MATERIALS REQUIRED:

Pencil

Ruler (cm/mm)

READ INSTRUCTIONS OVERLEAF

INSTRUCTIONS TO CANDIDATES

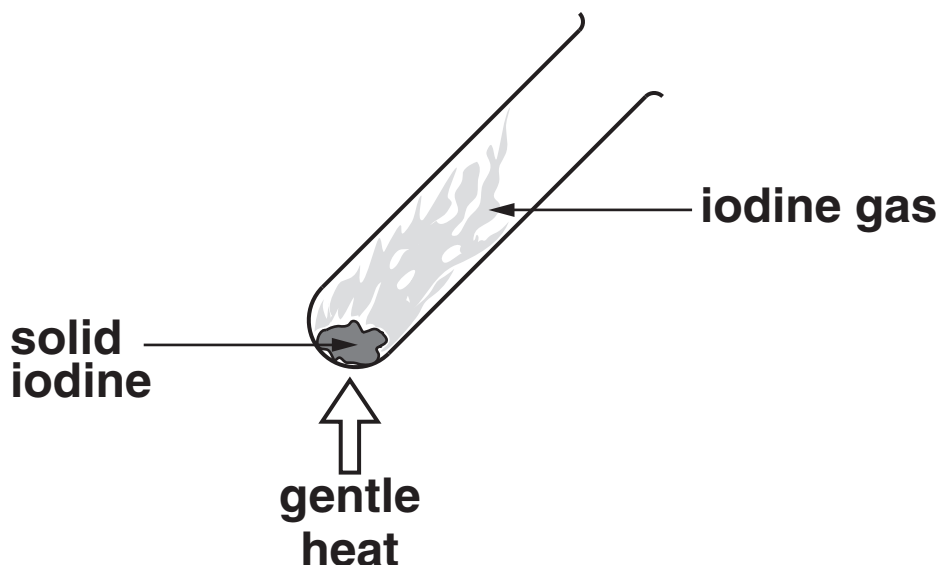
- Write your name clearly in capital letters, your Centre Number and Candidate Number in the boxes above.
- Use black ink. Pencil may be used for graphs and diagrams only.
- Read each question carefully and make sure that you know what you have to do before starting your answer.
- Answer **ALL** the questions.
- Write your answer to each question in the space provided. Additional paper may be used if necessary but you must clearly show your Candidate Number, Centre Number and question number(s).

INFORMATION FOR CANDIDATES

- The number of marks is given in brackets [] at the end of each question or part question.
- The total number of marks for this paper is **42**.
- The Periodic Table is printed on the back page.

Answer ALL the questions.

- 1 Iodine is a halogen in Group 7 of the Periodic Table.
Marty warms a small crystal of iodine in a test tube.



- (a) What colour change will Marty see when solid iodine changes into iodine gas?

Put a tick (✓) in the box next to the correct answer.

dark grey to orange

orange to yellow

dark grey to purple

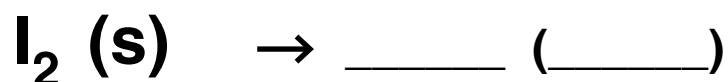
green to brown

[1]

- (b) The equation shows what happens when iodine changes from a solid to a gas.

Complete the symbol equation by filling in the missing formula and state symbols.

iodine solid → iodine gas



[2]

- (c) Marty finds some information about other halogens in Group 7.

NAME OF HALOGEN	MELTING POINT IN °C	BOILING POINT IN °C	REACTIVITY	FORMULA OF POTASSIUM SALT
fluorine	-220	-188	most reactive halogen	_____
chlorine	-101	-35	less reactive than fluorine more reactive than bromine	KCl
bromine	-7	59	less reactive than chlorine more reactive than iodine	KBr

- (i) Fill in the missing formula for potassium fluoride.

[1]

- (ii) Use information from the table to describe the trends in properties of the elements down Group 7.

[3]

- (d) Astatine is another halogen. It is below iodine in Group 7.

Complete the sentences about astatine by putting a **ring** around the correct word in each sentence.

Astatine is a **METAL** / **NON-METAL**.

The atomic mass of astatine is

BIGGER / **SMALLER** than the atomic mass
of iodine.

Astatine is **MORE** / **LESS** reactive than
iodine. [2]

[Total: 9]

2 Lithium is in Group 1.

Alex watches a video about lithium.

The video shows a piece of lithium being cut with a knife.

(a) What does Alex see when the freshly cut surface of lithium comes in contact with air?

Put a tick (✓) in the box next to the correct answer.

It starts to fizz.

It expands.

It catches fire.

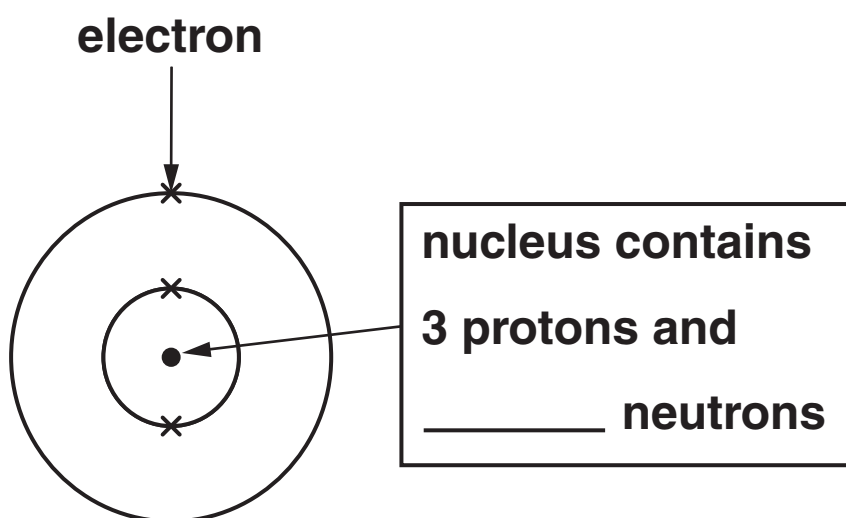
It goes from shiny to dull.

[1]

(b) The Periodic Table shows this information for lithium.

7
Li
lithium
3

The diagram shows an atom of lithium.

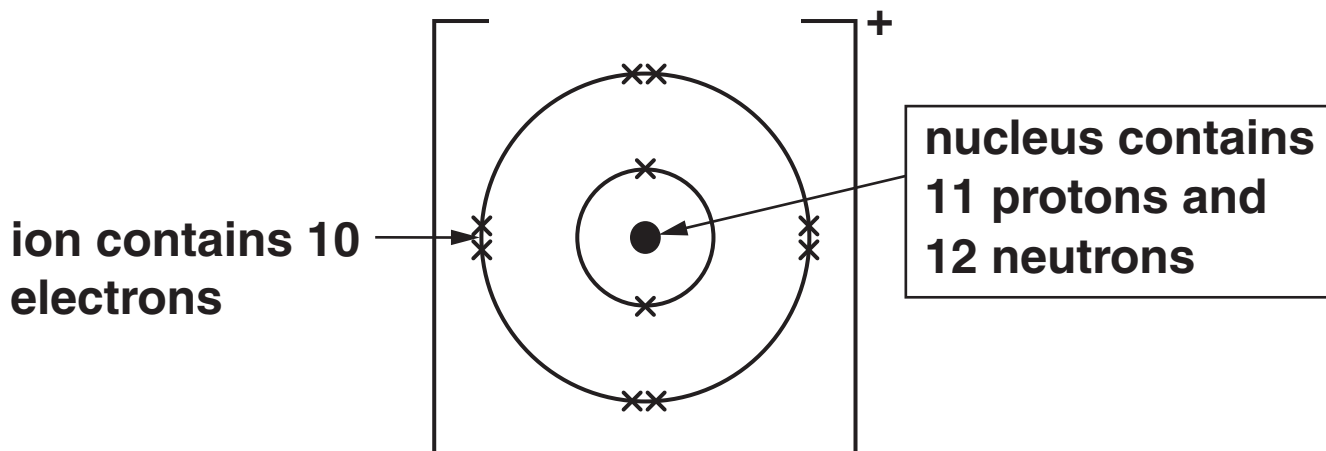


[1]

Complete the label on the diagram to show the number of neutrons in the nucleus of a lithium atom.

(c) Sodium is another element in Group 1.

The diagram shows a sodium ION.



Which one of the following statements about a sodium ION is correct?

Put a tick (✓) in the box next to the correct answer.

A sodium ion has more electrons than protons.

A sodium ion has a lower mass than a lithium atom.

A sodium ion has more protons than neutrons.

A sodium atom forms a sodium ion by losing one electron.

[1]

(d) Alex looks at lines in the spectra of lithium, sodium and potassium.

lithium



sodium



potassium



(i) How can you tell by looking at the spectra that these are three different elements?

_____ [1]

(ii) 'Healthy Salt' is a type of table salt that contains compounds of more than one Group 1 element.

This is the line spectrum for 'Healthy Salt'.



Which two Group 1 elements does 'Healthy Salt' contain?

answer _____ and _____ [1]

[Total: 5]

3 Bauxite is an ore of aluminium.

Bauxite contains the ionic compound, aluminium oxide.

Aluminium is extracted from aluminium oxide by electrolysis.

(a) The first stage of the process involves heating solid aluminium oxide until it melts.

What happens to the ions in aluminium oxide when it melts?

Put ticks (✓) in the boxes next to the TWO correct answers.

The ions become free to move.

The ions spread very far apart.

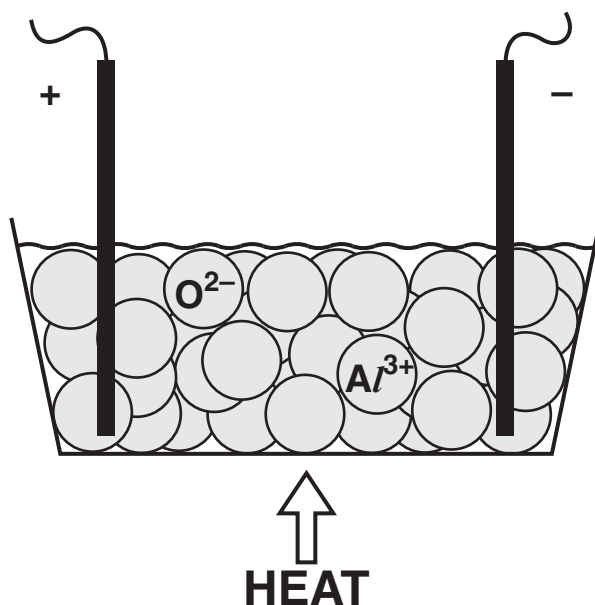
New bonds form between the ions.

The arrangement of ions becomes more random.

The ions move into a regular arrangement.

[2]

(b) The diagram shows how the electrolysis of aluminium oxide can be set up.



(i) Draw an arrow on the diagram to show which way the aluminium ions, Al^{3+} , move when the electrical current is switched on. [1]

(ii) Aluminium is one product of the electrolysis. What is the name of the OTHER product?

_____ [1]

- (c) One use of aluminium is to make overhead power cables.

Why is aluminium a good material to use to make overhead power cables?

Put ticks (✓) in the boxes next to the TWO best reasons.

Aluminium is a good conductor of heat.

Aluminium is less dense than other metals.

Aluminium has a lower melting point than some other metals.

Aluminium is a good electrical conductor.

Aluminium is softer than most other metals.

[2]

- (d) What type of bonding holds aluminium atoms together?

Put a **ring** around the correct answer.

ATOMIC

COVALENT

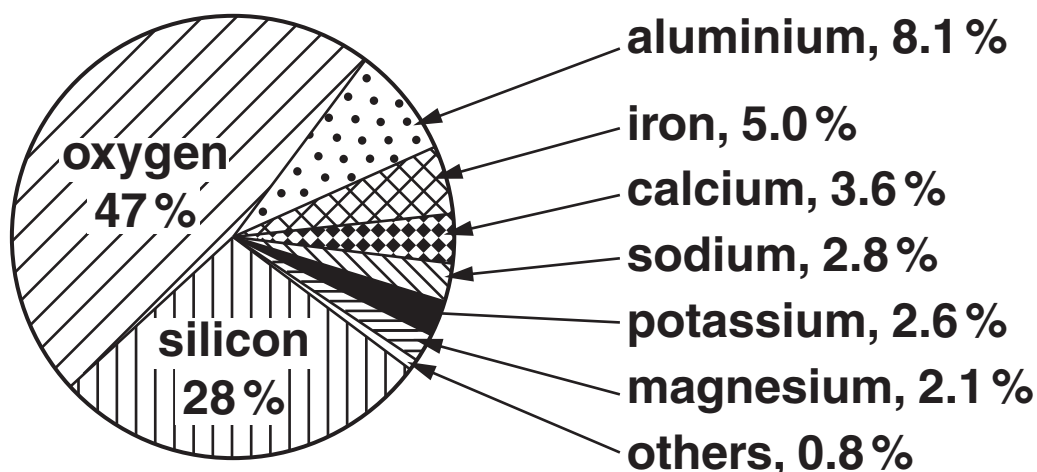
IONIC

METALLIC

[1]

[Total: 7]

- 4 The pie chart shows the percentages of common elements found in the Earth's crust.

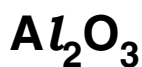


- (a) Three elements make up most of the Earth's crust.

The list below shows compounds found in the Earth's crust.

Which two compounds, when taken together, contain ALL THREE of these most common elements?

Put a **ring** around the TWO correct answers.



[2]

- (b) Sodium chloride, NaCl , is found in some parts of the Earth's crust. It is left behind when sea water evaporates.

Sodium appears in the pie chart of common elements in the Earth's crust but chlorine does not.

Which of the following statements explain why?

Put ticks (✓) in the boxes next to the TWO best answers.

There is much less chlorine than sodium in the Earth's crust.

Chlorine is a gas.

Sodium occurs in other compounds, not only sodium chloride.

The pie chart only shows metals.

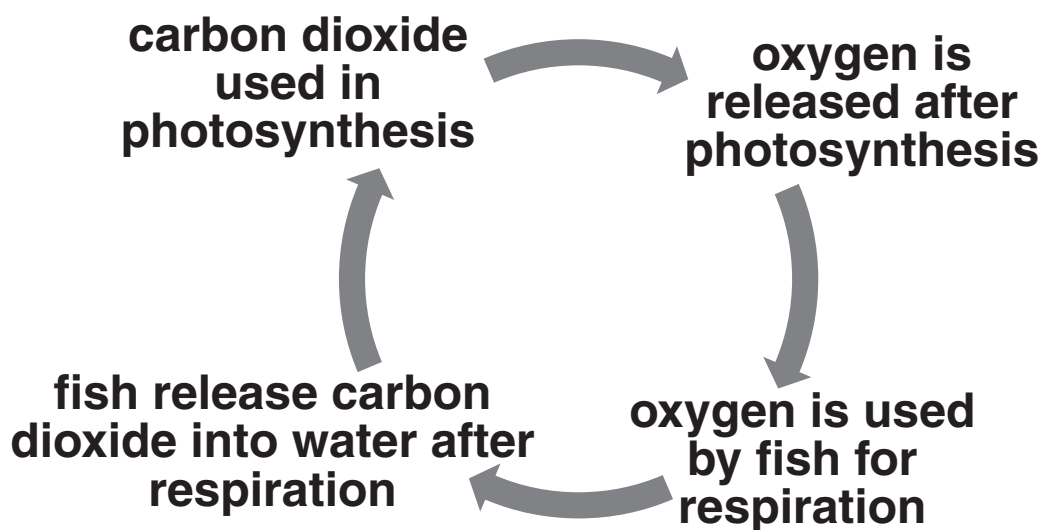
There is only a very small amount of chlorine in the sea.

[2]

[Total: 4]

5 A fish pond contains plants and fish.

The diagram shows how carbon dioxide and oxygen are involved in a cycle of changes in the fish pond.



(a) What effect do the FISH have on the amount of carbon dioxide and oxygen in the fish pond?

Put a tick (✓) in the correct box in each row.

	INCREASES	STAYS THE SAME	DECREASES
amount of carbon dioxide in the water			
amount of oxygen in the water			

[1]

(b) The diagram below shows the structure of a carbon dioxide molecule and an oxygen molecule.



Explain why carbon dioxide is a COMPOUND but oxygen is an ELEMENT.

[2]

[Total: 3]

6 Vehicles produce pollutant gases.

(a) One of the pollutant gases is nitrogen dioxide.

Nitrogen dioxide makes nitric acid when it dissolves in water. Nitric acid is very acidic.

How does the pH of the water change if a large volume of nitrogen dioxide is bubbled through it?

Put a ring around the correct answer in EACH ROW.

Pure water has a pH of

1 5 7 9 14

When a large volume of nitrogen dioxide is bubbled through water the pH changes to

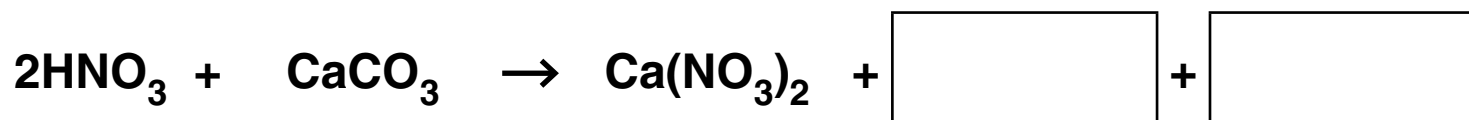
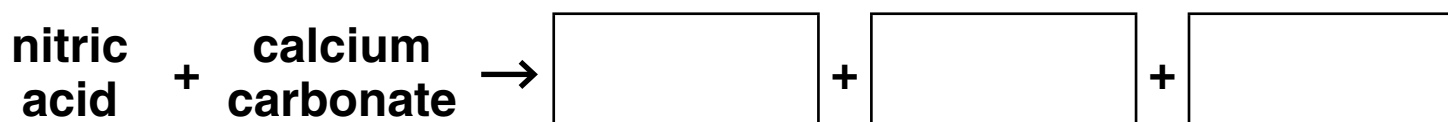
1 5 7 9 14

[1]

(b) Concrete contains calcium carbonate.

Nitric acid in rain water damages concrete because it reacts with calcium carbonate.

Complete the WORD and SYMBOL equations for this reaction by filling in the empty boxes.

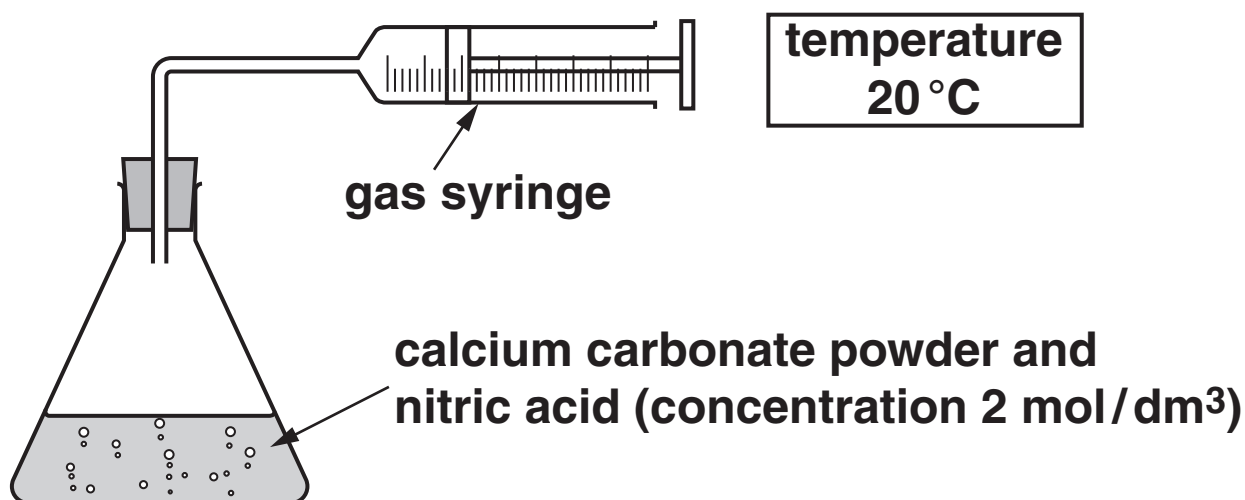


[3]

(c) Joe carries out an experiment to find out how quickly nitric acid reacts with calcium carbonate.

He follows the rate of reaction by measuring the amount of gas given off every 30 s.

This is how he sets his experiment up.



(i) The reaction happens too quickly for Joe to measure the volume of gas accurately.

Suggest **THREE** changes that Joe could make to his experiment so that the reaction happens more slowly.

1 _____

2 _____

3 _____ [3]

(ii) Joe decides to use another method of following the rate of reaction.

He puts the flask on a balance and measures the change in mass during the reaction.

The mass decreases.

Explain why.

_____ [1]

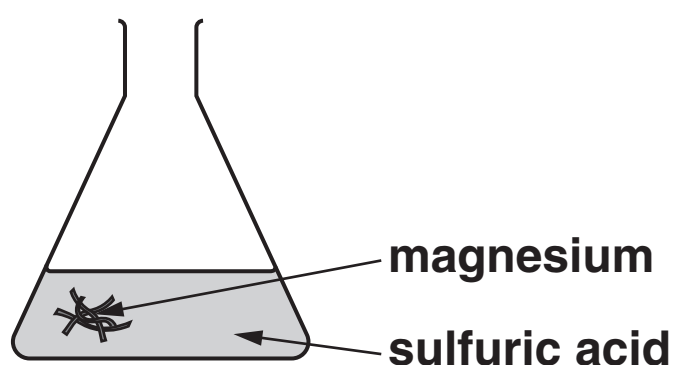
[Total: 8]

7 Jack works for a company that makes chemicals to make soil more fertile.

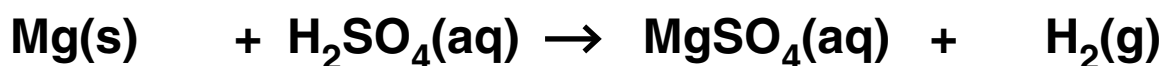
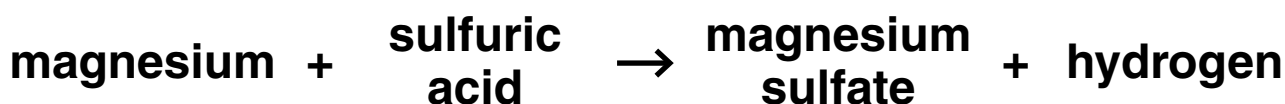
The company sells magnesium sulfate to treat soil that does not have enough magnesium.

Jack makes a small amount of magnesium sulfate to test.

He makes the magnesium sulfate by reacting magnesium metal with dilute sulfuric acid.



These are the word and symbol equations for the reaction.



(a) What will Jack SEE during the reaction?

[2]

(b) Jack keeps adding magnesium until the reaction finishes.

The sentences below show some steps he might carry out to make magnesium sulfate crystals.

They are in the wrong order.

- A pat dry with filter paper
- B leave to crystallise
- C filter off unreacted magnesium
- D evaporate some liquid by heating

Write the letters in the boxes to show the correct order of steps Jack needs to carry out.

The first one has been done for you.

correct order

C			
---	--	--	--

 [1]

(c) Jack works out the theoretical yield of magnesium sulfate for this reaction.

He measures the actual yield he has made.

The box shows his results.

theoretical yield: 2.5g
actual yield: 2.0g

What is Jack's PERCENTAGE YIELD?

Put a ring around the correct answer.

0.5% 5% 20% 25% 80%

(d) Jack wants to make a LARGER amount of magnesium sulfate.

What TWO changes does he need to make to his experiment to make more magnesium sulfate?

Put ticks (✓) in the boxes next to the TWO correct answers.

use more acid

heat the reaction mixture to a higher temperature

use smaller pieces of magnesium

use a catalyst

use more magnesium

leave the reaction for a longer time

[2]

[Total: 6]

END OF QUESTION PAPER

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