

GENERAL CERTIFICATE OF SECONDARY EDUCATION

GATEWAY SCIENCE

B624/02

ADDITIONAL SCIENCE B

Unit 2 Modules B4 C4 P4
(Higher Tier)

Candidates answer on the question paper
A calculator may be used for this paper

OCR Supplied Materials:
None

Other Materials Required:

- Pencil
- Ruler (cm/mm)

Wednesday 10 June 2009

Afternoon

Duration: 1 hour



Candidate Forename		Candidate Surname	
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Centre Number						Candidate Number				
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MODIFIED LANGUAGE

INSTRUCTIONS TO CANDIDATES

- Write your name clearly in capital letters, your Centre Number and Candidate Number in the boxes above.
- Use black ink. Pencil may be used for graphs and diagrams only.
- Read each question carefully and make sure that you know what you have to do before starting your answer.
- Answer **all** the questions.
- Do **not** write in the bar codes.
- Write your answer to each question in the space provided, however additional paper may be used if necessary.

INFORMATION FOR CANDIDATES

- The number of marks is given in brackets [] at the end of each question or part question.
- A list of physics equations is printed on page two.
- The Periodic Table is printed on the back page.
- The total number of marks for this paper is **60**.
- This document consists of **24** pages. Any blank pages are indicated.

2

EQUATIONS

$$\text{speed} = \frac{\text{distance}}{\text{time taken}}$$

$$\text{acceleration} = \frac{\text{change in speed}}{\text{time taken}}$$

$$\text{force} = \text{mass} \times \text{acceleration}$$

$$\text{work done} = \text{force} \times \text{distance}$$

$$\text{power} = \frac{\text{work done}}{\text{time}}$$

$$\text{kinetic energy} = \frac{1}{2} mv^2$$

$$\text{potential energy} = mgh$$

$$\text{weight} = \text{mass} \times \text{gravitational field strength}$$

$$\text{resistance} = \frac{\text{voltage}}{\text{current}}$$

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Question 1 begins on page 4.

PLEASE DO NOT WRITE ON THIS PAGE

Answer **all** the questions.

Section A – Module B4

1 Read this newspaper article carefully.

	<p>The blue hole</p>
	<p>Scientists have just discovered a deep, blue hole in a forest in the Bahamas.</p>
	<p>The hole is full of water and is about 35 metres deep.</p>
	<p>The water is pure at the surface. It becomes more and more salty and contains less oxygen deeper in the hole.</p>
	<p>At the bottom of the hole scientists have found the bodies of animals and plants that have not decayed. They are thousands of years old.</p>
	<p>“The plants are so well preserved they still have green chloroplasts” said one scientist.</p>

(a) The decomposers can **not** decay the dead animals and plants at the bottom of the hole.

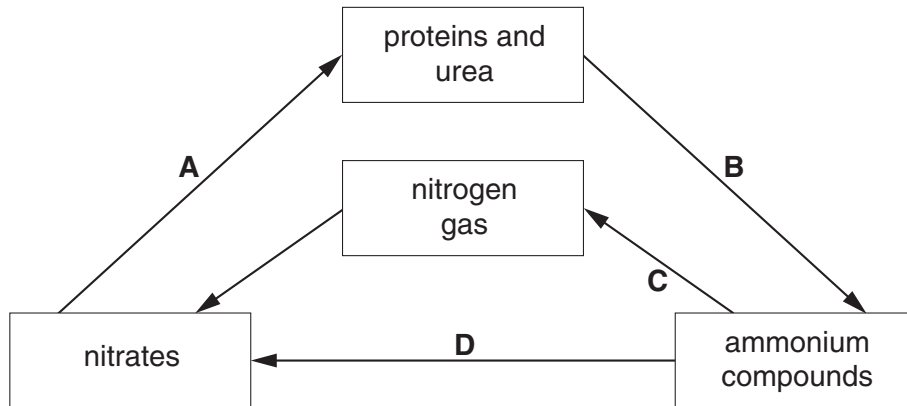
Write down **two** reasons why.

1

2 [2]

(b) The decay of dead plants and animals is an important step in the nitrogen cycle.

The diagram shows part of the cycle.



(i) Which step in the cycle shows decay?

Choose from **A**, **B**, **C** or **D**.

answer

[1]

(ii) Nitrogen gas can be converted into nitrogen compounds, such as nitrates.

Write down **two** ways that this can happen.

1

2 [2]

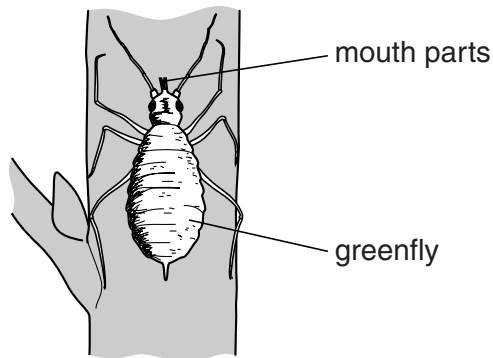
(c) The scientist says that the plants still have green chloroplasts.

Which cells in plant leaves contain most chloroplasts?

..... [1]

[Total: 6]

- 2 The diagram shows a greenfly feeding from the stem of a tomato plant.



The greenfly pushes a hollow tube into one of the tissues in the plant stem.

It can then take sugar from this tissue.

- (a) Suggest which tissue the greenfly is most likely to pierce to get the sugar solution.

Put a ring around the answer in this list.

epidermis palisade phloem xylem

[1]

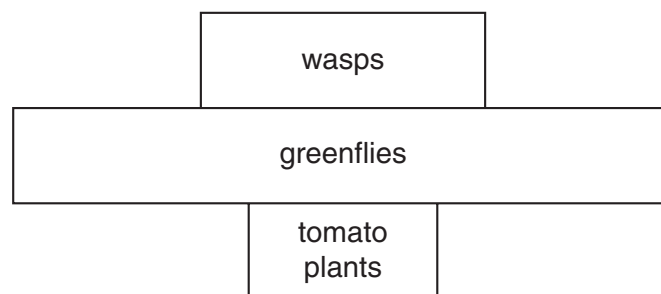
- (b) Tomato plants are often grown in glasshouses.

The plants produce fewer tomatoes when greenflies feed on them.

A gardener releases some wasps into his glasshouse.

The wasps eat the greenflies.

- (i) The following diagram gives information about the food chain in the glasshouse.



Write down the name of this type of diagram.

..... [1]

(ii) Using wasps to eat greenflies is an example of biological control.

Many gardeners prefer to use biological control instead of chemical pesticides.

Write down **two** advantages of using biological control.

1

2 [2]

[Total: 4]

3 Plants need minerals to grow.

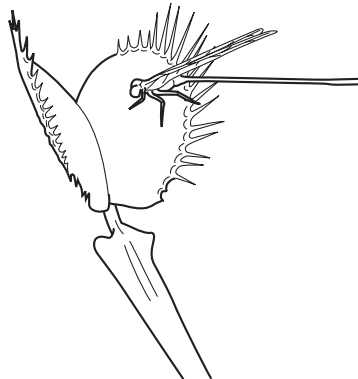
They usually get these minerals from the soil.

(a) Some plants can **not** get enough minerals from the soil.

Their leaves are adapted to trap insects.

They digest the insects to get the minerals they need.

One plant that does this is the venus fly trap.



The venus fly trap does not get enough nitrates from the soil.

Instead it gets nitrogen compounds from the insects.

(i) What do plants look like if they do not get enough nitrates?

..... [1]

(ii) The venus fly trap also needs magnesium compounds.

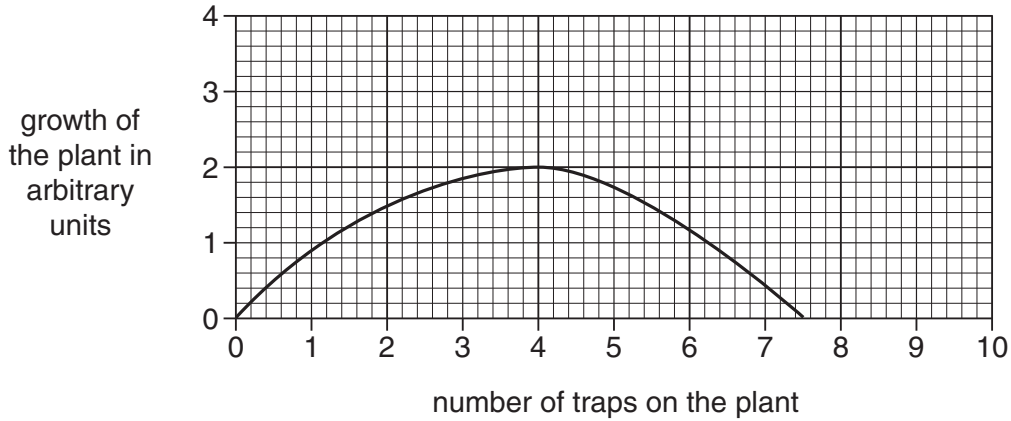
Write down **one** chemical in the plant that contains magnesium.

..... [1]

(b) Plants like the venus fly trap use energy to make traps to catch insects.

This energy comes from sugars made in photosynthesis.

The graph shows how different numbers of traps affect the growth of a plant.



(i) Use the graph to predict how many traps a plant should make for maximum growth.

..... [1]

(ii) Suggest why the plant grows **less** well if it makes more or less traps.

.....
.....
..... [2]

(iii) Most plant leaves are **not** adapted to catch insects.

The leaves are adapted for photosynthesis by being broad and thin.

Explain how these adaptations help with photosynthesis.

Leaves are broad because

Leaves are thin because

[Total: 7]

4 Anil is growing some lettuce plants in his garden.

Normally they look like the plants in the first diagram.



Anil goes outside on a hot day to look at the lettuce plants.

They look different.



(a) Anil's lettuce plants have lost water and wilted.

Why does loss of water from the plants cause wilting?

.....

.....

..... [2]

(b) In what conditions would Anil's plants wilt fastest?

Put a tick (✓) in the box next to the correct answer.

light and humid

dark and warm

dry and windy

windy and humid

[1]

[Total: 3]

Section B – Module C4

5 This question is about fertilisers.

Look at the diagram. It shows the label on a bag of fertiliser.



(a) Ammonium nitrate, NH_4NO_3 , is a fertiliser.

(i) Anna makes some ammonium nitrate crystals.

She uses ammonia solution and an acid.

Write down the **name** of the acid.

..... [1]

(ii) What is the relative formula mass (M_r) of ammonium nitrate, NH_4NO_3 ?

The relative atomic mass (A_r) of H is 1, of N is 14 and of O is 16.

.....
.....

relative formula mass is [1]

(b) The overuse of fertilisers sometimes causes eutrophication.

Eutrophication may cause living things in the water to die.

Write about what happens during eutrophication.

Include in your answer

- how the fertiliser gets into the water
- what the fertiliser does in the water
- how this affects the living things in the water.

.....
.....
.....
..... [3]

[Total: 5]

6 This question is about washing powders.

(a) Suggest a reason, other than cost, why it is good to wash clothes at **40°C** rather than at **50°C**.

.....
 [1]

(b) A detergent molecule has two ends

- a hydrophilic head
- a hydrophobic tail.



(i) What does the **hydrophilic** head do during cleaning?

..... [1]

(ii) What does the **hydrophobic** tail do during cleaning?

..... [1]

(c) Another way of cleaning clothes is to use a dry cleaning solvent.

(i) What is meant by **dry** cleaning?

..... [1]

(ii) Why is dry cleaning sometimes used instead of normal washing?

..... [1]

[Total: 5]

7 This question is about water.

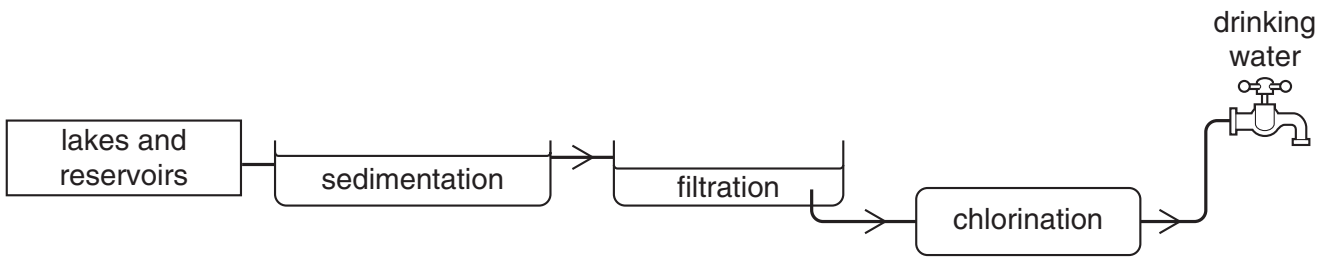
Water is taken from lakes and reservoirs.

The water contains microbes, soluble materials and insoluble materials.

Water is treated to make it safe for drinking.

Look at the diagram.

It shows some of the stages used in treating water.



(a) (i) Explain what happens during sedimentation.

..... [1]

(ii) Explain why the water goes through a chlorination process.

..... [1]

(b) River water may contain many substances before it is purified.

The water may contain **pesticides**.

The pesticides get into the river from the land.

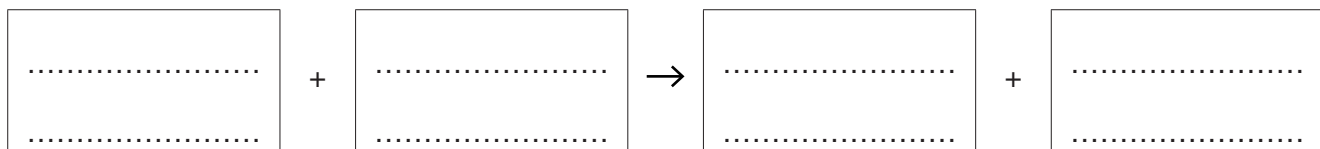
Suggest how pesticides get into the river.

..... [1]

(c) Sodium chloride reacts with silver nitrate.

Sodium nitrate and silver chloride are made.

Write a **word** equation for this reaction.



[1]

(d) Barium chloride, BaCl_2 , reacts with sodium sulfate, Na_2SO_4 .

Barium sulfate, BaSO_4 , and sodium chloride, NaCl , are made.

Write a **balanced symbol** equation for this reaction.

..... [2]

[Total: 6]

8 This question is about acids and bases.

An acid and base react together.

A salt and water are made.

(a) Look at the list.

It shows the names of some salts.

ammonium chloride

ammonium nitrate

ammonium sulfate

potassium sulfate

sodium chloride

sodium nitrate

Sodium hydroxide reacts with hydrochloric acid.

Write down the name of the salt made.

Choose from the list.

..... [1]

(b) Calcium nitrate, $\text{Ca}(\text{NO}_3)_2$, is another salt.

(i) How many different **elements** are there in this formula?

..... [1]

(ii) John makes some calcium nitrate in the laboratory.

He expects to make 80 g of calcium nitrate.

He only makes 64 g.

Calculate his percentage yield.

.....
.....
.....

answer % [2]

[Total: 4]

Section C – Module P4

9 This question is about static electricity.

(a) Static electricity can be dangerous when refuelling an aircraft.

Suggest why.

.....
..... [1]

(b) A plastic ruler is rubbed with a cloth.

The ruler becomes **positively** charged.

Explain how the ruler becomes positively charged.

.....
.....
..... [2]

(c) Static electricity can also be useful.

It is used in hospitals.



A doctor can **restart** a patient's **heart**.

He puts the paddles on the patient's chest.

The paddles are charged.

Describe what happens next.

In your answer write about

- how the heart restarts
- the precautions taken.

.....

.....

.....

..... [2]

[Total: 5]

10 A hair dryer is an electrical appliance.

(a) The hair dryer is **double insulated**.



It is not earthed.

Explain why the hair dryer is not earthed.

.....
..... [1]

(b) The hair dryer is connected to a 230V mains supply.

The current through the hair dryer is 5 A.

Calculate the **resistance** of the hair dryer.

The equations on page 2 may help you.

.....
.....

answer ohms [2]

[Total: 3]

11 **Ultrasound** scans are used in hospitals.



(a) Ultrasound waves are used to build up a picture of an unborn baby.

Explain how the waves build up this picture.

.....
.....
..... [2]

(b) X-rays are **not** used for scanning unborn babies.

One reason is that X-rays can damage living cells.

Write down **one other** reason why X-rays are not used.

.....
..... [1]

[Total: 3]

12 Gamma rays and X-rays are used to treat cancer.

(a) Gamma rays come from the nucleus of some radioactive materials.

How are X-rays made?

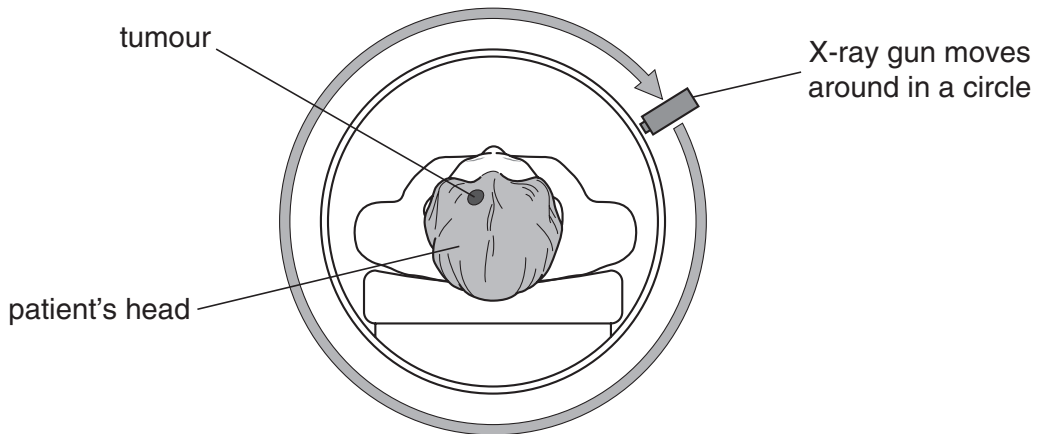
.....
..... [1]

(b) (i) X-rays are often better than gamma rays for cancer treatment.

Suggest why.

.....
..... [1]

(ii) An X-ray gun is rotated around a cancer patient's head.



The X-rays are always aimed at the tumour.

This kills the cancer cells from all directions.

Give **one other** reason why the X-ray gun is moved around the cancer patient's head.

.....
..... [1]

[Total: 3]

13 There are three types of nuclear radiation.

These are alpha, beta and gamma.

Look at the table.

type of radiation	description
alpha	${}^4_2\text{He}$
beta	${}^0_{-1}\beta$
gamma	electromagnetic wave

(a) The nucleus of a radioactive atom decays and emits alpha, beta or gamma radiation.

What is special about a radioactive nucleus?

Finish the sentence.

A radioactive atom decays because its nucleus is [1]

(b) An alpha particle is a helium nucleus.

What is a beta particle?

Finish the sentence.

A beta particle is a high speed [1]

(c) Background radiation is always present in the environment.

Where does background radiation come from?

Write down **two** sources of background radiation.

1

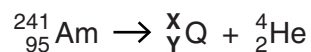
2 [2]

(d) An americium nucleus, Am, decays.

It emits an alpha particle.

A new element, Q, is produced.

Look at the equation.



(i) What is the **mass number (X)** of Q?

..... [1]

(ii) What is the **atomic number (Y)** of Q?

..... [1]

[Total: 6]

END OF QUESTION PAPER

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The Periodic Table of the Elements

1	2	3	4	5	6	7	0	
7 Li lithium 3	9 Be beryllium 4	11 Na sodium 11	12 C carbon 6	13 Al aluminium 13	14 N nitrogen 7	15 O oxygen 8	16 F fluorine 9	17 Ne neon 10
19 K potassium 19	20 Ca calcium 20	23 V vanadium 23	24 Cr chromium 24	25 Mn manganese 25	26 Fe iron 26	27 Co cobalt 27	28 Ni nickel 28	29 Cu copper 29
37 Rb rubidium 37	38 Sr strontium 38	40 Zr zirconium 40	41 Nb niobium 41	42 Mo molybdenum 42	43 Tc technetium [98]	44 Ru ruthenium 44	45 Rh rhodium 45	46 Pd palladium 46
55 Cs caesium 55	56 Ba barium 56	57 La* lanthanum 57	58 Ce cerium 58	59 Pr praseodymium 59	60 Nd neodymium 60	61 Pm promethium [61]	62 Sm samarium 62	63 Eu europium 63
87 Fr francium 87	88 Ra radium 88	89 Ac* actinium 89	90 Th thorium 90	91 Pa protactinium 91	92 U uranium 92	93 Np neptunium [93]	94 Pu plutonium 94	95 Am americium 95
133 Cs caesium 55	137 Ba barium 56	139 La* lanthanum 57	140 Ce cerium 58	141 Pr praseodymium 59	142 Nd neodymium 60	143 Pm promethium [61]	144 Sm samarium 62	145 Eu europium 63
209 Po polonium 84	210 At astatine 85	211 Rn radon 86	212 Fr francium 87	213 Ra radium 88	214 Ac actinium 89	215 Th thorium 90	216 Pa protactinium 91	217 U uranium 92
204 Tl thallium 81	207 Pb lead 82	208 Bi bismuth 83	209 Po polonium 84	210 At astatine 85	211 Rn radon 86	212 Fr francium 87	213 Ra radium 88	214 Ac actinium 89
115 In indium 49	119 Sn tin 50	122 Sb antimony 51	126 Te tellurium 52	128 I iodine 53	131 Xe xenon 54	133 Cs caesium 55	137 Ba barium 56	139 La* lanthanum 57
112 Cd cadmium 48	115 In indium 49	119 Sn tin 50	122 Sb antimony 51	126 Te tellurium 52	128 I iodine 53	131 Xe xenon 54	133 Cs caesium 55	137 Ba barium 56
65 Zn zinc 30	69 Ga gallium 31	73 Ge germanium 32	75 As arsenic 33	79 Se selenium 34	81 Br bromine 35	83 Kr krypton 36	85 Rb rubidium 37	87 Fr francium 87
63.5 Cu copper 29	65 Zn zinc 30	69 Ga gallium 31	73 Ge germanium 32	75 As arsenic 33	79 Se selenium 34	81 Br bromine 35	83 Kr krypton 36	85 Rb rubidium 37
59 Co cobalt 27	59 Ni nickel 28	63.5 Cu copper 29	65 Zn zinc 30	69 Ga gallium 31	73 Ge germanium 32	75 As arsenic 33	79 Se selenium 34	81 Br bromine 35
103 Rh rhodium 45	106 Pd palladium 46	108 Ag silver 47	112 Cd cadmium 48	115 In indium 49	119 Sn tin 50	122 Sb antimony 51	126 Te tellurium 52	128 I iodine 53
192 Ir iridium 77	195 Pt platinum 78	197 Au gold 79	201 Hg mercury 80	204 Tl thallium 81	207 Pb lead 82	209 Bi bismuth 83	210 At astatine 85	211 Rn radon 86
268 Mt meitnerium 109	271 Ds darmstadtium 110	272 Rg roentgenium 111	Elements with atomic numbers 112-116 have been reported but not fully authenticated					

1	H	hydrogen	1
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relative atomic mass
atomic symbol
name
atomic (proton) number

* The lanthanoids (atomic numbers 58-71) and the actinoids (atomic numbers 90-103) have been omitted.

The relative atomic masses of copper and chlorine have not been rounded to the nearest whole number.