


## Chemistry Data Sheet

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### 1. Reactivity Series of Metals

Potassium Sodium Calcium Magnesium Aluminium <i>Carbon</i> Zinc Iron Tin Lead <i>Hydrogen</i> Copper Silver Gold Platinum	most reactive  least reactive
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(elements in italics, though non-metals, have been included for comparison)

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### 2. Formulae of Some Common Ions

#### Positive ions

Name	Formula
Hydrogen	H <sup>+</sup>
Sodium	Na <sup>+</sup>
Silver	Ag <sup>+</sup>
Potassium	K <sup>+</sup>
Lithium	Li <sup>+</sup>
Ammonium	NH <sub>4</sub> <sup>+</sup>
Barium	Ba <sup>2+</sup>
Calcium	Ca <sup>2+</sup>
Copper(II)	Cu <sup>2+</sup>
Magnesium	Mg <sup>2+</sup>
Zinc	Zn <sup>2+</sup>
Lead	Pb <sup>2+</sup>
Iron(II)	Fe <sup>2+</sup>
Iron(III)	Fe <sup>3+</sup>
Aluminium	Al <sup>3+</sup>

#### Negative ions

Name	Formula
Chloride	Cl <sup>-</sup>
Bromide	Br <sup>-</sup>
Fluoride	F <sup>-</sup>
Iodide	I <sup>-</sup>
Hydroxide	OH <sup>-</sup>
Nitrate	NO <sub>3</sub> <sup>-</sup>
Oxide	O <sup>2-</sup>
Sulfide	S <sup>2-</sup>
Sulfate	SO <sub>4</sub> <sup>2-</sup>
Carbonate	CO <sub>3</sub> <sup>2-</sup>

**Turn over ►**

### 3. The Periodic Table of Elements

1	2	3	4	5	6	7	0
7 <b>Li</b> lithium 3	9 <b>Be</b> beryllium 4	11 <b>B</b> boron 5	12 <b>C</b> carbon 6	14 <b>N</b> nitrogen 7	16 <b>O</b> oxygen 8	19 <b>F</b> fluorine 9	20 <b>Ne</b> neon 10
23 <b>Na</b> sodium 11	24 <b>Mg</b> magnesium 12	27 <b>Al</b> aluminium 13	28 <b>Si</b> silicon 14	31 <b>P</b> phosphorus 15	32 <b>S</b> sulfur 16	35.5 <b>Cl</b> chlorine 17	40 <b>Ar</b> argon 18
39 <b>K</b> potassium 19	40 <b>Ca</b> calcium 20	45 <b>Sc</b> scandium 21	48 <b>Ti</b> titanium 22	51 <b>V</b> vanadium 23	52 <b>Cr</b> chromium 24	55 <b>Mn</b> manganese 25	56 <b>Fe</b> iron 26
85 <b>Rb</b> rubidium 37	88 <b>Sr</b> strontium 38	89 <b>Y</b> yttrium 39	91 <b>Zr</b> zirconium 40	93 <b>Nb</b> niobium 41	96 <b>Mo</b> molybdenum 42	[98] <b>Tc</b> technetium 43	101 <b>Ru</b> ruthenium 44
133 <b>Cs</b> caesium 55	137 <b>Ba</b> barium 56	139 <b>La*</b> lanthanum 57	178 <b>Hf</b> hafnium 72	181 <b>Ta</b> tantalum 73	184 <b>W</b> tungsten 74	186 <b>Re</b> rhenium 75	190 <b>Os</b> osmium 76
[223] <b>Fr</b> francium 87	[226] <b>Ra</b> radium 88	[227] <b>Ac*</b> actinium 89	[261] <b>Rf</b> rutherfordium 104	[262] <b>Db</b> dubnium 105	[266] <b>Sg</b> seaborgium 106	[264] <b>Bh</b> bohrium 107	[277] <b>Hs</b> hassium 108
112 <b>Cd</b> cadmium 48	112 <b>Cd</b> cadmium 48	112 <b>Cd</b> cadmium 48	112 <b>Cd</b> cadmium 48	112 <b>Cd</b> cadmium 48	112 <b>Cd</b> cadmium 48	112 <b>Cd</b> cadmium 48	112 <b>Cd</b> cadmium 48
65 <b>Zn</b> zinc 30	63.5 <b>Cu</b> copper 29	59 <b>Ni</b> nickel 28	59 <b>Co</b> cobalt 27	59 <b>Co</b> cobalt 27	59 <b>Ni</b> nickel 28	59 <b>Ni</b> nickel 28	65 <b>Zn</b> zinc 30
115 <b>In</b> indium 49	115 <b>In</b> indium 49	115 <b>In</b> indium 49	115 <b>In</b> indium 49	115 <b>In</b> indium 49	115 <b>In</b> indium 49	115 <b>In</b> indium 49	115 <b>In</b> indium 49
119 <b>Sn</b> tin 50	119 <b>Sn</b> tin 50	119 <b>Sn</b> tin 50	119 <b>Sn</b> tin 50	119 <b>Sn</b> tin 50	119 <b>Sn</b> tin 50	119 <b>Sn</b> tin 50	119 <b>Sn</b> tin 50
122 <b>Sb</b> antimony 51	122 <b>Sb</b> antimony 51	122 <b>Sb</b> antimony 51	122 <b>Sb</b> antimony 51	122 <b>Sb</b> antimony 51	122 <b>Sb</b> antimony 51	122 <b>Sb</b> antimony 51	122 <b>Sb</b> antimony 51
127 <b>I</b> iodine 53	127 <b>I</b> iodine 53	127 <b>I</b> iodine 53	127 <b>I</b> iodine 53	127 <b>I</b> iodine 53	127 <b>I</b> iodine 53	127 <b>I</b> iodine 53	127 <b>I</b> iodine 53
131 <b>Xe</b> xenon 54	131 <b>Xe</b> xenon 54	131 <b>Xe</b> xenon 54	131 <b>Xe</b> xenon 54	131 <b>Xe</b> xenon 54	131 <b>Xe</b> xenon 54	131 <b>Xe</b> xenon 54	131 <b>Xe</b> xenon 54
204 <b>Tl</b> thallium 81	204 <b>Tl</b> thallium 81	204 <b>Tl</b> thallium 81	204 <b>Tl</b> thallium 81	204 <b>Tl</b> thallium 81	204 <b>Tl</b> thallium 81	204 <b>Tl</b> thallium 81	204 <b>Tl</b> thallium 81
207 <b>Pb</b> lead 82	207 <b>Pb</b> lead 82	207 <b>Pb</b> lead 82	207 <b>Pb</b> lead 82	207 <b>Pb</b> lead 82	207 <b>Pb</b> lead 82	207 <b>Pb</b> lead 82	207 <b>Pb</b> lead 82
209 <b>Bi</b> bismuth 83	209 <b>Bi</b> bismuth 83	209 <b>Bi</b> bismuth 83	209 <b>Bi</b> bismuth 83	209 <b>Bi</b> bismuth 83	209 <b>Bi</b> bismuth 83	209 <b>Bi</b> bismuth 83	209 <b>Bi</b> bismuth 83
[210] <b>At</b> astatine 85	[210] <b>At</b> astatine 85	[210] <b>At</b> astatine 85	[210] <b>At</b> astatine 85	[210] <b>At</b> astatine 85	[210] <b>At</b> astatine 85	[210] <b>At</b> astatine 85	[210] <b>At</b> astatine 85
[222] <b>Rn</b> radon 86	[222] <b>Rn</b> radon 86	[222] <b>Rn</b> radon 86	[222] <b>Rn</b> radon 86	[222] <b>Rn</b> radon 86	[222] <b>Rn</b> radon 86	[222] <b>Rn</b> radon 86	[222] <b>Rn</b> radon 86
Elements with atomic numbers 112 – 116 have been reported but not fully authenticated							

\* The Lanthanides (atomic numbers 58 – 71) and the Actinides (atomic numbers 90 – 103) have been omitted.

**Cu** and **Cl** have not been rounded to the nearest whole number.

1	<b>H</b>	hydrogen	1
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#### Key

relative atomic mass
<b>atomic symbol</b>
name
atomic (proton) number