

AS: 2812

**January 2001**

**Chains and Rings  
Mark Scheme**

1. (a) (i) OH/hydroxy/hydroxyl/ROH ✓  
(ii)  $C_nH_{2n+1}OH$  ✓  
(iii)  $C_7H_{15}OH$  /  $C_7H_{16}O$  ✓
- (b) 88. ✓
- (c)(i) 1 mark for plotting the points ✓  
1 mark for the line extended to enable b.pt of  $C_8H_{17}OH$  to be estimated. ✓
- (ii) I butan-1-ol 115 -125 °C ✓  
II  $C_8H_{17}OH$  190 -205 °C ✓
- (ii) Boiling point increases as the  $M_r$  increases/ proportional to  $M_r$  . ✓

[Total : 9]

2 (a) contains carbon and hydrogen **only** ✓

separates due to differences in boiling point ✓

(b) works out/uses  $M_r = 156$  ✓

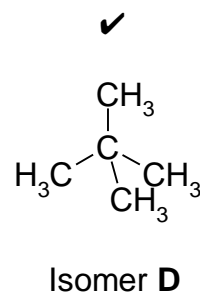
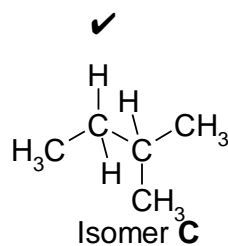
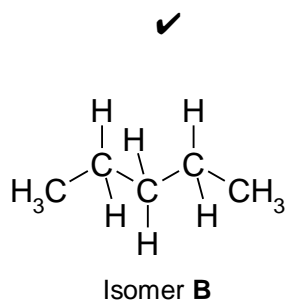
(ii) 132/156 method mark ✓

84.6% C ✓

(c)  $C_{11}H_{24} \rightarrow C_9H_{20} \checkmark + C_2H_4 \checkmark$

Ethene ✓

(d) (i) Draw the isomers of pentane.



(ii) **D, C, B** to match as drawn in (d)(i) ✓

(iii) less van der Waals' forces in **D**/ as chain length increases so does b pt./greater the branching~lower the boiling point ✓

(iv)  $C_5H_{12} + 8O_2 \rightarrow 5CO_2 + 6H_2O$  ✓✓

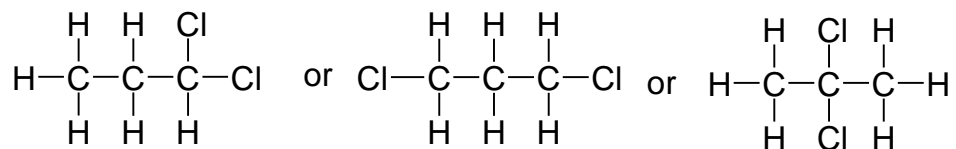
( $CO_2 + H_2O$  gets ✓)

(v) branched chains burn more efficiently/ add it to petrol ✓

[Total : 16]

		4		
3	(a)	Initiation	$\text{Cl}_2 \rightarrow 2\text{Cl}\bullet$	✓
		Propagation 1	$\text{C}_3\text{H}_8 + \text{Cl}\bullet \rightarrow \text{HCl} + \text{C}_3\text{H}_7\bullet$	✓
		Propagation 2	$\text{C}_3\text{H}_7\bullet + \text{Cl}_2 \rightarrow \text{C}_3\text{H}_7\text{Cl} + \text{Cl}\bullet$	✓
		Termination	Any two free radicals	✓
				[4]

- (b) (i) Compound **H** = 1,2-dichloropropane ✓
- (ii) 1 mark for each correct structure ✓✓

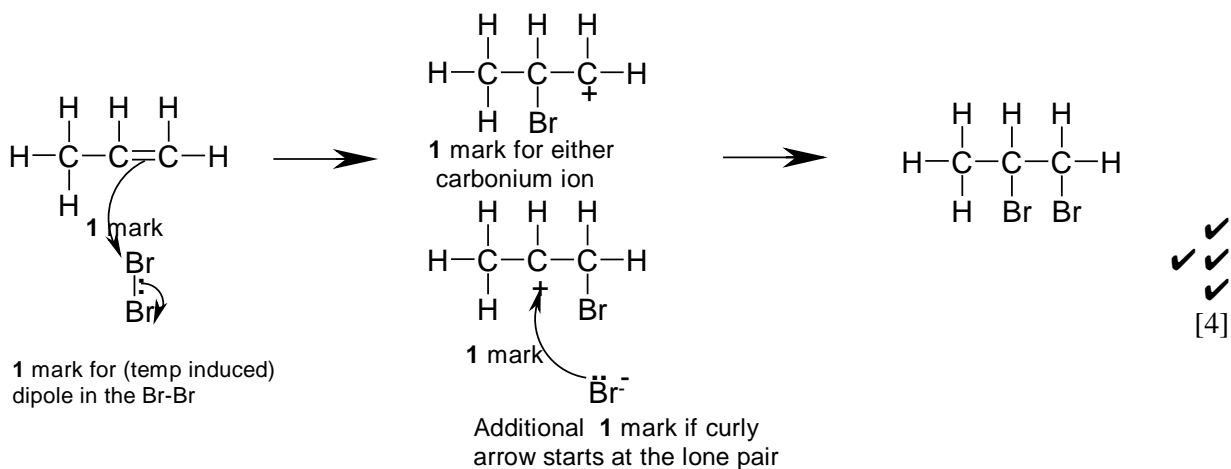


- (c) (i) water ✓
- (ii)  $\text{OH}^-$  behaves as a nucleophile ✓
- $\text{OH}^-$  has a lone pair of electrons/ seeks out electron deficient areas/attracted to  $\text{C}^{\delta+}$  ✓
- (d) (i) reflux is the **continuous** process of **evaporation** followed by **condensation**/ description of what would be seen to indicate that the process is continuous ✓
- (ii) orange ✓
- to green ✓
- (iii)  $\text{C}_3\text{H}_7\text{OH}/\text{C}_3\text{H}_8\text{O} + 2[\text{O}] \rightarrow \text{C}_2\text{H}_5\text{COOH}/\text{C}_3\text{H}_6\text{O}_2 + \text{H}_2\text{O}$  ✓✓
- (All three formulae correct gets one mark)

- (e) wavenumber 1680 – 1750  $\text{cm}^{-1}$  ✓
- bond C=O ✓
- wavenumber 2500 – 3300  $\text{cm}^{-1}$  ✓
- bond O-H ✓

[Total : 19]

4 (a)

5 marking points for **max** of 4 marks

- (b) (i)
- (ii) reagent =  $\text{H}_2$  ✓
- conditions = Ni/Pt as catalyst ✓
- (iii) ✓✓
- $\begin{array}{c} \text{H} & \text{H} & \text{H} \\ | & | & | \\ \text{H}-\text{C}-\text{C}-\text{C}-\text{H} \\ | & | & | \\ \text{H} & \text{H} & \text{OH} \end{array}$  and  $\begin{array}{c} \text{H} & \text{H} & \text{H} \\ | & | & | \\ \text{H}-\text{C}-\text{C}-\text{C}-\text{H} \\ | & | & | \\ \text{H} & \text{OHH} & \end{array}$
- (c) (i) Addition polymer ✓
- (ii)
- $\left( \begin{array}{cccc} \text{H} & \text{CH}_3 & \text{H} & \text{CH}_3 \\ | & | & | & | \\ -\text{C}- & \text{C}- & \text{C}- & \text{C}- \\ | & | & | & | \\ \text{H} & \text{H} & \text{H} & \text{H} \end{array} \right)$
- (iii) non-biodegradable or words to that effect ✓
- when burnt they release toxic fumes ✓

[Total : 13]

5 (a) (i)  $\text{C}_4\text{H}_9\text{OH}/\text{C}_4\text{H}_{10}\text{O}$ 

✓



(b) The upper layer because the organic compounds have a **lower density** than water. ✓

(c) (i)  $\text{CO}_2$  ✓

(ii)  $\text{HCl}$  ✓

(d) (i)  $51\text{ }^\circ\text{C}$  ✓

(ii)  $4/74 = 0.05(4)$  ✓

(iii)  $3.75/92.5 = 0.04(1)$  ✓

(iv) 75% (allow 80% if (d) (ii) given as 0.05 / mark ecf for (d)(ii)/(d)(i) \*100) ✓

[Total : 9]

6 (a) functional group 1 alkene ✓

test add bromine ✓

observation decolourised ✓

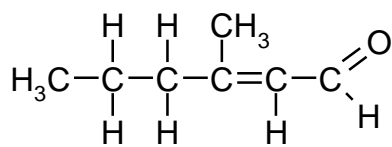
functional group 2 alcohol ✓

test Na/  $\text{PCl}_5$ /  $\text{RCO}_2\text{H}$  ✓

observation bubbles/ $\text{H}_2$  white fumes/ $\text{HCl}$  smell ✓

[6]

(b)



Compound F

✓

[1]

[Total : 7]

7	Fermentation		✓
	Yeast/enzyme		✓
	Temperature about 30 °C		✓
	$C_6H_{12}O_6 \rightarrow 2C_2H_5OH + 2CO_2$		✓
	Batch process		✓
	Hydration of ethene.		✓
	Reagent	steam/water at > 100 °C	✓
	Temp/press	300 °C & 70 –100 atm	✓
	Catalyst	phosphoric acid	✓
	$C_2H_4 + H_2O \rightarrow C_2H_5OH$		✓
	Continuous process		✓
	1 mark available for <i>Quality of written communication</i> ..... base the award of the mark on the ability to communicate the essential chemistry		✓

[Total : 12 max = 9]

8	Ethane	saturated/single bonds only/ $\sigma$ -bond	✓
		tetrahedron	✓
		109° 28'	✓
	Ethene	unsaturated/double bonds/contains a $\pi$ -bond	✓
		draws or explains overlap of adjacent p-orbitals at right angle to the plane of the molecule	✓
		trigonal planar	✓
		approx 120°	✓

1 mark available for *Quality of written communication* ..... base the award of the mark on the ability to use essential technical language such as *saturated/unstaurated/tetrahedron, trigonal planar/ overlap of adjacent p-orbitals*

[Total : 8]