

**Modified Enlarged 24pt**  
**OXFORD CAMBRIDGE AND RSA EXAMINATIONS**

**Friday 23 June 2023 – Morning**

**A Level Chemistry A**

**H432/03 Unified chemistry**

**Time allowed: 1 hour 30 minutes**  
**plus your additional time allowance**

**YOU MUST HAVE:**

**the Data Sheet for Chemistry A**

**YOU CAN USE:**

**a scientific or graphical calculator**  
**an HB pencil**

**Please write clearly in black ink.**

**Centre number**

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**Candidate number**

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**First name(s)** \_\_\_\_\_

**Last name** \_\_\_\_\_

**READ INSTRUCTIONS OVERLEAF**



# **INSTRUCTIONS**

**Use black ink. You can use an HB pencil, but only for graphs and diagrams.**

**Write your answer to each question in the space provided. If you need extra space use the lined pages at the end of this booklet. The question numbers must be clearly shown.**

**Answer ALL the questions.**

**Where appropriate, your answer should be supported with working. Marks might be given for using a correct method, even if your answer is wrong.**

# **INFORMATION**

**The total mark for this paper is 70.**

**The marks for each question are shown in brackets [ ].**

**Quality of extended response will be assessed in questions marked with an asterisk (\*).**

## **ADVICE**

**Read each question carefully before you start your answer.**

- 1 This question is about the first 36 elements in the periodic table opposite:**

**From these 36 elements ONLY, write the symbol for the element(s) that fit each description.**

- (a) The THREE elements that, when in a molecule, can form hydrogen bonds with other suitable molecules.**

\_\_\_\_\_ **[1]**

H																	He
Li	Be											B	C	N	O	F	Ne
Na	Mg											Al	Si	P	S	Cl	Ar
K	Ca	Sc	Ti	V	Cr	Mn	Fe	Co	Ni	Cu	Zn	Ga	Ge	As	Se	Br	Kr

**(b) The element in Period 2 with the successive ionisation energies shown below.**

<b>Ionisation number</b>	<b>Ionisation energy / kJ mol<sup>-1</sup></b>
<b>1st</b>	<b>1000</b>
<b>2nd</b>	<b>2251</b>
<b>3rd</b>	<b>3361</b>
<b>4th</b>	<b>4564</b>
<b>5th</b>	<b>7012</b>
<b>6th</b>	<b>8496</b>
<b>7th</b>	<b>27 107</b>
<b>8th</b>	<b>31 671</b>

\_\_\_\_\_ [1]

**(c) An element that is a solid at RTP with a simple molecular lattice structure.**

\_\_\_\_\_ [1]

**(d) The TWO elements with atoms containing five unpaired d electrons.**

\_\_\_\_\_ [2]

**(e) The element in Period 3 that exists in the solid state as a network of atoms bonded by strong covalent bonds.**

\_\_\_\_\_ [1]

**(f) The p-block element in Period 3 that forms a compound with fluorine with octahedral molecules.**

\_\_\_\_\_ [1]

**(g) The element that forms 1– ions most readily.**

\_\_\_\_\_ [1]

**(h) The element with an average atomic mass of  $1.244 \times 10^{-22}$  g.**

\_\_\_\_\_ **[1]**



**2 These questions are from different areas of chemistry.**

**(a) This question is about two salts of rubidium (atomic number 37):  $\text{RbClO}_3$  and  $\text{RbClO}_4$ .**

**(i) The oxidation number of chlorine is different in the two rubidium salts,  $\text{RbClO}_3$  and  $\text{RbClO}_4$ .**

**What is the name of  $\text{RbClO}_4$ ?**

**\_\_\_\_\_ [1]**

- (ii) A student carries out an experiment to determine the enthalpy change of solution of  $\text{RbClO}_3$  using the method below.

A 2.00 g sample of solid  $\text{RbClO}_3$  is added to water in a well-insulated container. The initial temperature is  $23.0^\circ\text{C}$ .

The mixture is stirred until all the  $\text{RbClO}_3$  has dissolved. The final temperature is  $21.5^\circ\text{C}$ . The final solution has a mass of 102 g.

Determine the enthalpy change of solution,  $\Delta_{\text{sol}}H$ , of  $\text{RbClO}_3$  in  $\text{kJ mol}^{-1}$ .

Assume that the specific heat capacity of the solution is the same as that of pure water.

$$\Delta_{\text{sol}}H (\text{RbClO}_3) = \underline{\hspace{2cm}} \text{ kJ mol}^{-1} \text{ [3]}$$

**(b) A student investigates the rate of a reaction that is 1st order with respect to hydrochloric acid,  $\text{HCl(aq)}$ .**

**The student carries out a reaction using  $0.680 \text{ mol dm}^{-3} \text{ HCl(aq)}$ . The initial rate is  $9.52 \times 10^{-4} \text{ mol dm}^{-3} \text{ s}^{-1}$ .**

**The student dilutes a different sample of  $0.680 \text{ mol dm}^{-3} \text{ HCl(aq)}$  with water. The pH of this diluted acid is 1.50.**

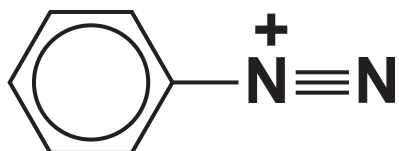
**The student repeats the reaction using the same volume of this diluted acid.**

**Determine the initial rate using this diluted acid.**

**initial rate = \_\_\_\_\_ mol dm<sup>-3</sup> s<sup>-1</sup> [3]**

- (c) The benzenediazonium ion, shown below, is stable at temperatures below 10 °C.

### Benzenediazonium



Above 10 °C, the benzenediazonium ion reacts with water to form phenol.

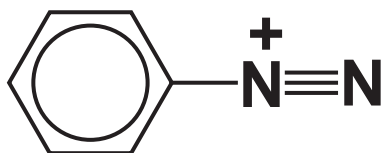
The reaction proceeds in a three-step mechanism.

**STEP 1** Elimination of nitrogen gas to form a carbocation.

**STEP 2** Nucleophilic attack by water.

**STEP 3** Proton loss to form the organic product.

Complete the boxes opposite with intermediates and curly arrows to show the mechanism for this reaction. [4]

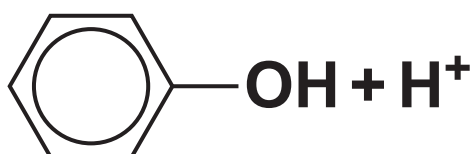


**STEP 1**

+ \_\_\_\_\_

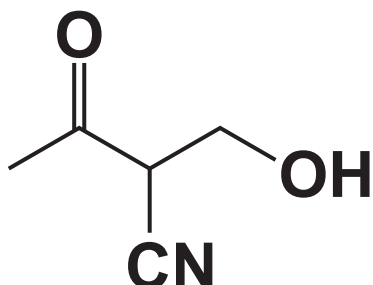
**STEP 2**

**STEP 3**



**3\*** A chemist is investigating compound A, shown below, as a potential organic intermediate.

**Compound A**



**Describe the type of stereoisomerism shown by compound A and suggest THREE reactions of compound A, one for each of the three functional groups using reagents of your choice.**

**In your answer, show stereoisomers of compound A, your chosen reactants and conditions, and the structures for the organic products produced.**

**Mechanisms and equations are NOT required.** **[6]**

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[illegible]

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**Additional answer space if required.**

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**4** **Tutton's salts are 'double salts' with the formula  $X_2Y(ZO_4)_2 \cdot 6H_2O$ .**

**A Tutton's salt contains two cations:  $X^+$  and  $Y^{2+}$ .**

**$X^+$  can be an ion of the Group 1 elements K, Rb, Cs or Fr, or an ammonium ion.**

**$Y^{2+}$  can be a  $2+$  ion of magnesium or an ion of most of the transition elements in Period 4.**

**Z can be S or Cr.**

**$(NH_4)_2Cu(SO_4)_2 \cdot 6H_2O$  is an example of a Tutton's salt.**

**(a) Predict the formula of a Tutton's salt containing different ions from  $(NH_4)_2Cu(SO_4)_2 \cdot 6H_2O$ .**

**\_\_\_\_\_ [1]**

**(b) A student prepares a sample of the Tutton's salt,  $(\text{NH}_4)_2\text{Cu}(\text{SO}_4)_2 \cdot 6\text{H}_2\text{O}$  using the method shown below.**

**STEP 1 Dissolve 0.025 mol of ammonium sulfate and 0.025 mol of hydrated copper(II) sulfate,  $\text{CuSO}_4 \cdot 5\text{H}_2\text{O}$ , in water and make up to  $50 \text{ cm}^3$ .**

**STEP 2 Boil the resulting mixture for 2 minutes and allow to cool.**

**STEP 3 Allow the solvent to evaporate slowly. Pale blue crystals of the Tutton's salt form.**

**(i) What masses are needed  
of ammonium sulfate and  
 $\text{CuSO}_4 \cdot 5\text{H}_2\text{O}$ ?**

**mass of ammonium sulfate \_\_\_\_\_ g**

**mass of  $\text{CuSO}_4 \cdot 5\text{H}_2\text{O}$  \_\_\_\_\_ g  
[2]**

**(ii) In STEP 3, why does the student  
allow the solvent to evaporate  
and NOT boil off all the solvent  
in STEP 2?**

\_\_\_\_\_  
\_\_\_\_\_  
\_\_\_\_\_ [1]

- (c) The student dissolves their Tutton's salt in water. A pale blue solution forms.**

**The student carries out two tests on this aqueous solution.**

- (i) The student adds an excess of aqueous ammonia to their aqueous solution of Tutton's salt. A deep blue solution forms.**

**The complex ion responsible for the deep blue solution has a molar mass of  $167.5 \text{ g mol}^{-1}$ .**

**Suggest the formula of this complex ion.**

**\_\_\_\_\_ [1]**

**(ii) The student adds NaOH(aq) to the aqueous solution of Tutton's salt and warms the mixture.**

**A precipitate and a gas are formed.**

**Write the formulae of the precipitate and gas and suggest a test that could confirm the identity of the gas.**

**Formula of precipitate \_\_\_\_\_**

**Formula of gas \_\_\_\_\_**

**Test to confirm the identity of the gas**

\_\_\_\_\_  
\_\_\_\_\_  
\_\_\_\_\_ **[3]**



**(iii) How could the student carry out a test-tube test to confirm the anion in the Tutton's salt?**

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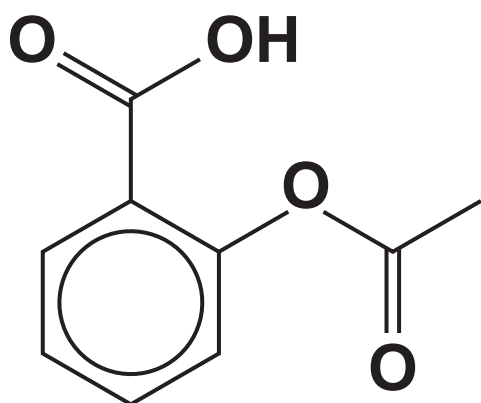
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**[2]**

**5 Aspirin tablets are used for pain relief.**

**The structure of aspirin is shown below.**

**Aspirin**



- (a) A student uses the reaction of aspirin with cold NaOH(aq) to determine the mass of aspirin in ONE tablet.**

**In this reaction, 1 mol of aspirin reacts with 1 mol of cold NaOH(aq).**

**The student's method is outlined below.**

**STEP 1** The student reacts  
**THREE** aspirin tablets  
with  $100\text{ cm}^3$  of  
 $0.500\text{ mol dm}^{-3}\text{ NaOH(aq)}$ .  
The NaOH is in excess.  
A colourless solution forms.

**STEP 2** The colourless solution  
from STEP 1 is made up  
to  $250.0\text{ cm}^3$  with distilled  
water.

**STEP 3** A  $25.00\text{ cm}^3$  sample of  
the diluted solution from  
STEP 2 is titrated with  
 $0.200\text{ mol dm}^{-3}\text{ HCl(aq)}$  in  
the burette.

The  $\text{HCl(aq)}$  reacts with excess  
 $\text{NaOH(aq)}$  that remains in STEP 1:



The student repeats the titration  
to obtain concordant (consistent)  
titres.

## **TITRATION RESULTS**

**The trial titre has been omitted.**

**The burette readings have been read to the nearest  $0.05\text{ cm}^3$ .**

	<b>1</b>	<b>2</b>	<b>3</b>
<b>Final reading / <math>\text{cm}^3</math></b>	<b>23.10</b>	<b>45.40</b>	<b>27.40</b>
<b>Initial reading / <math>\text{cm}^3</math></b>	<b>0.00</b>	<b>23.10</b>	<b>5.00</b>

## **ANALYSIS OF RESULTS**

**From the results, the student can determine the following.**

- 1. The amount, in mol, of excess NaOH(aq) that remains after the reaction of aspirin with NaOH(aq).**
- 2. The amount, in mol, of NaOH(aq) that reacted with the aspirin.**

**Use the results to determine the mass, in mg, of aspirin in ONE aspirin tablet.**

mass of aspirin  
in ONE tablet = \_\_\_\_\_ mg [6]

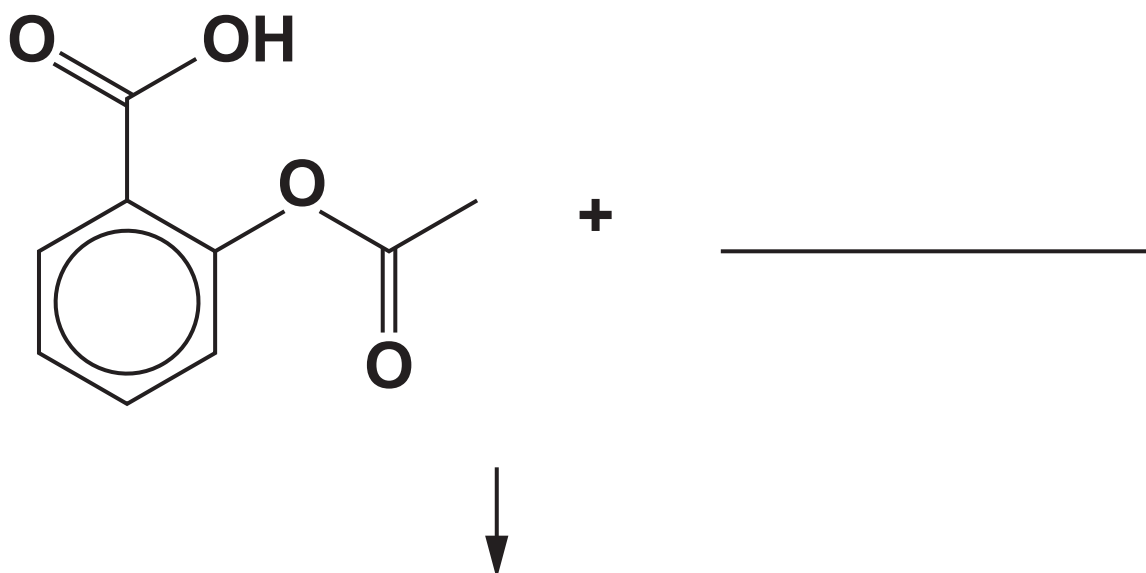
**(b) Aspirin reacts with hot NaOH(aq), under reflux.**

**(i) Draw a labelled diagram of suitable apparatus for reflux. Use the space below. [2]**

(ii) In this reaction, 1 mol of aspirin reacts with 3 mol of hot NaOH(aq).

Complete the equation for the reaction of aspirin with an excess of hot NaOH(aq). [3]

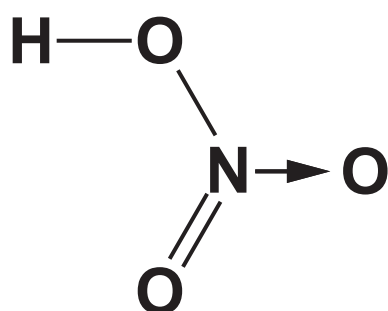
Show structures for organic compounds.



**6 This question is about nitric acid, hydrochloric acid and sulfuric acid.**

**(a) Nitric acid has 2 single covalent bonds, 1 double covalent bond and 1 dative covalent bond as shown below.**

**Nitric acid**



**Predict the H–O–N and O–N–O bond angles in nitric acid.**

**Explain your reasoning.**

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[4]

**(b) Dilute nitric acid reacts with aluminium oxide to form a solution of aluminium nitrate.**

**(i) Write an equation for this reaction.**

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[2]

**(ii) The solution contains nitrate ions,  $\text{NO}_3^-$ .**

**Draw a 'dot-and-cross' diagram for the  $\text{NO}_3^-$  ion.**

**Use a different symbol for the extra electron. Use the space below. [2]**

- (c) A mixture of concentrated nitric and hydrochloric acid is called 'aqua regia'. Aqua regia can dissolve gold.

The reaction of aqua regia with gold is a redox reaction which forms chlorauric acid,  $\text{HAuCl}_4$ .

- (i) Balance the half-equation for the oxidation process in this reaction. [1]



- (ii) In the reduction process in this reaction,  $\text{HNO}_3$  and  $\text{H}^+$  react together to form 2 oxides: X ( $M_r = 30$ ) and Z ( $M_r = 18$ ).

Determine the formulae of X and Z and write the half-equation for this reduction.

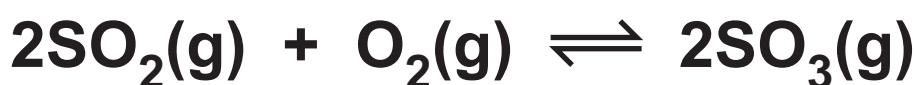
X = \_\_\_\_\_

Z = \_\_\_\_\_

half-equation \_\_\_\_\_ [3]

**(d) In the UK, most sulfuric acid,  $\text{H}_2\text{SO}_4$ , is manufactured by the Contact process.**

**One stage in the Contact process involves the equilibrium between sulfur dioxide, oxygen and sulfur trioxide.**



**This equilibrium is investigated:**

**STEP 1**  $5.82 \times 10^{-2}$  mol of  $\text{SO}_2$  is mixed with  $7.40 \times 10^{-2}$  mol of  $\text{O}_2$  in a  $2.00 \text{ dm}^3$  container.

**STEP 2** The container is sealed and allowed to reach equilibrium at constant temperature.

**STEP 3** At equilibrium,  $5.20 \times 10^{-2}$  mol of  $\text{SO}_3$  is formed.

**Determine the equilibrium concentrations and calculate  $K_c$ , including units.**

$K_c =$  \_\_\_\_\_ units \_\_\_\_\_ [5]

**(e)\*Three reactions involving sulfuric acid are shown below.**

### **REACTION 1**

**Dilute sulfuric acid is reacted with nickel(II) hydroxide to form a green solution.**

**The solvent is allowed to evaporate leaving hydrated crystals of compound D, with the percentage composition by mass: Ni, 22.33%; S, 12.20%; O, 60.87%; H, 4.60%.**

### **REACTION 2**

**Concentrated sulfuric acid is reacted with hydrogen bromide, HBr, to form three products:  
an element which exists as diatomic molecules  
a gaseous compound E  
a liquid.**

**At RTP, 1.00 dm<sup>3</sup> of compound E has a mass of 2.67 g.**



### REACTION 3

Concentrated sulfuric acid acts as a catalyst when 2-hydroxypropanoic acid reacts to form compound F ( $M_r = 144$ ).

In this reaction, 2 mol of 2-hydroxypropanoic acid forms 1 mol of compound F and 2 mol of water.

Identify compounds D, E and F and construct equations for the reactions. [6]

Show structures for any organic compounds.

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**Additional answer space if required.**

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**END OF QUESTION PAPER**

**ADDITIONAL ANSWER SPACE**

**If additional space is required, you should use the following lined page(s). The question number(s) must be clearly shown in the margin(s).**


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[illegible]




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