

**OXFORD CAMBRIDGE AND RSA EXAMINATIONS**

**Wednesday 19 June 2019 – Morning**

**A Level Chemistry A**

**H432/03 Unified chemistry**

**Time allowed: 1 hour 30 minutes  
plus your additional time allowance**

**YOU MUST HAVE:  
the Data Sheet for Chemistry A**

**YOU MAY USE:  
a scientific or graphical calculator**

**Please write clearly in black ink.**

**Centre number**

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**Candidate number**

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**First name(s)** \_\_\_\_\_

**Last name** \_\_\_\_\_

**READ INSTRUCTIONS OVERLEAF**



## **INSTRUCTIONS**

**Use black ink. You may use an HB pencil for graphs and diagrams.**

**Answer ALL the questions.**

**Where appropriate, your answers should be supported with working. Marks may be given for a correct method even if the answer is incorrect.**

**Write your answer to each question in the space provided. If additional space is required, use the lined page(s) at the end of this booklet. The question number(s) must be clearly shown.**

## **INFORMATION**

**The total mark for this paper is 70.**

**The marks for each question are shown in brackets [ ].**

**Quality of extended responses will be assessed in questions marked with an asterisk (\*).**

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**Answer ALL the questions.**

**1 These short questions are from different areas of chemistry.**

**(a) Explain why a  $\text{CF}_4$  molecule has polar bonds but does NOT have an overall dipole.**

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**[2]**

**(b) Explain why a small proportion of molecules in water have a relative molecular mass of 20.**

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**[1]**

**(c) What is the partial pressure of  $\text{O}_2$  (in Pa) in a gas mixture containing 21%  $\text{O}_2$  by volume and with a total pressure of  $1.0 \times 10^5 \text{ Pa}$ ?**

**partial pressure of  $\text{O}_2$  = \_\_\_\_\_ Pa [1]**

- (d) What mass of carbon dioxide (in g) is formed by the complete combustion of  $42.0 \text{ m}^3$  (measured at RTP) of propane?

mass = \_\_\_\_\_ g [2]

- (e) A reaction is first order with respect to  $\text{H}^+$ . At a pH of 1, the initial rate is  $2.4 \times 10^{-3} \text{ mol dm}^{-3} \text{ s}^{-1}$ .

What is the initial rate at a pH of 3?

initial rate = \_\_\_\_\_  $\text{mol dm}^{-3} \text{ s}^{-1}$  [1]

- (f) What is the number of oxygen atoms in 4.26 g of  $\text{P}_2\text{O}_5$ ?

number of oxygen atoms = \_\_\_\_\_ [2]

- 2 Benzoic acid,  $\text{C}_6\text{H}_5\text{COOH}$ , is added to some foods as a preservative.

A student prepares benzoic acid as outlined below.

**STEP 1** The student mixes  $4.00\text{ cm}^3$  of phenylmethanol,  $\text{C}_6\text{H}_5\text{CH}_2\text{OH}$ , (density =  $1.04\text{ g cm}^{-3}$ ) with sodium carbonate and aqueous potassium manganate(VII), as an oxidising agent. The mixture is heated under reflux.

**STEP 2** The resulting mixture is cooled and then acidified with concentrated  $\text{HCl}$ . Impure crystals of benzoic acid appear.

**STEP 3** The student recrystallises the impure crystals to obtain  $1.59\text{ g}$  of pure benzoic acid.

- (a) In STEP 1, sodium carbonate,  $\text{Na}_2\text{CO}_3$ , makes the reaction mixture alkaline.

Write an ionic equation to show how carbonate ions form an alkaline solution in water.

\_\_\_\_\_ [1]

- (b) In STEP 2, explain why the mixture must be acidified so that crystals of benzoic acid appear.

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[1]

- (c) Write the overall equation for the preparation of benzoic acid from phenylmethanol.

Use [O] for the oxidising agent.

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[1]

- (d) Calculate the percentage yield of benzoic acid.

Give your answer to 3 significant figures.

percentage yield = \_\_\_\_\_ % [3]

**(e) In STEP 3, describe how the student can recrystallise the impure crystals to obtain pure benzoic acid.**

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**[2]**

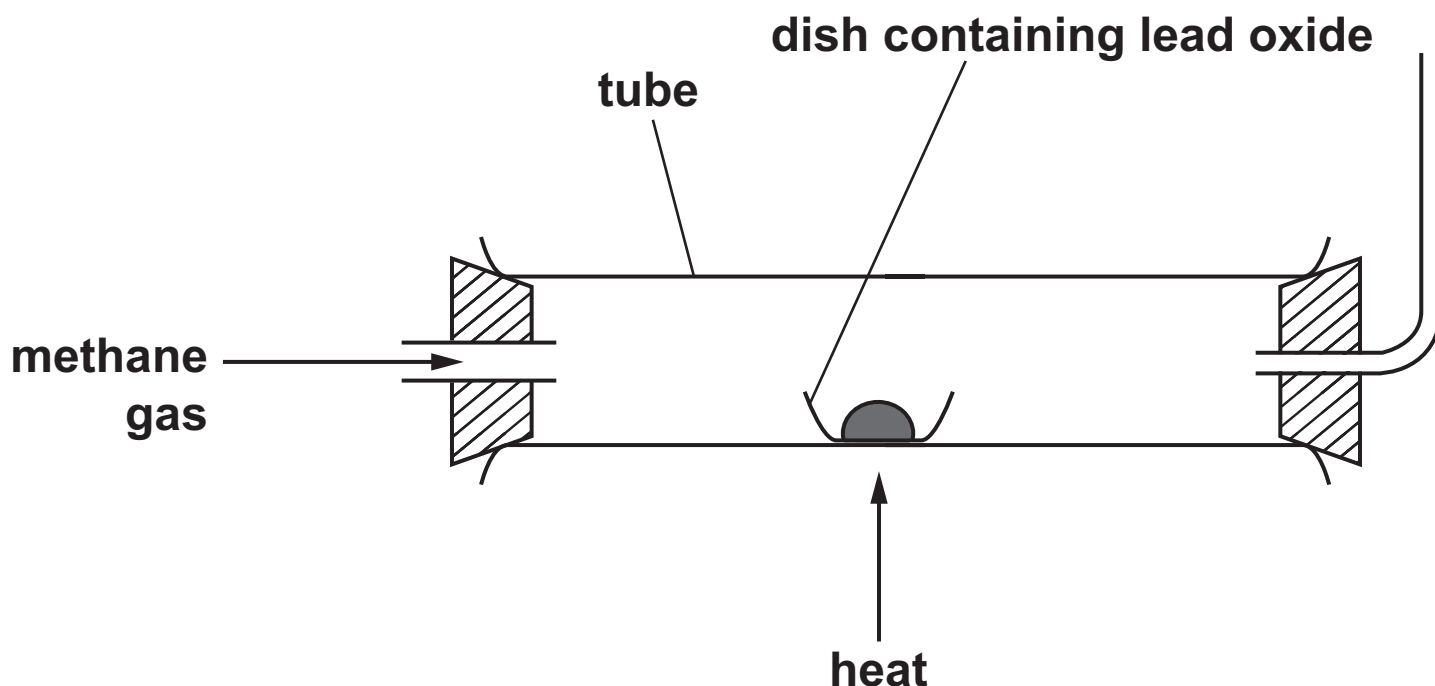
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3 This question is about elements and compounds in Group 14 (Group 4) of the periodic table.

(a) There are four oxides of lead:  $\text{PbO}$ ,  $\text{PbO}_2$ ,  $\text{Pb}_2\text{O}_3$  and  $\text{Pb}_3\text{O}_4$ .

A student carries out an experiment to identify an unknown lead oxide, which is one of the four oxides of lead shown above.

The student plans to reduce the unknown lead oxide to lead by heating the lead oxide in a stream of methane gas,  $\text{CH}_4$ . The apparatus is shown below.



## STUDENT'S METHOD

**Weigh an empty dish.**

**Add the lead oxide to the dish and reweigh.**

**Set up the apparatus and pass methane gas through the tube as shown.**

**Heat the dish for 10 minutes.**

**Pass cold air through the tube to cool the dish and contents.**

**Weigh the dish and contents.**

- (i) Write the equation for the reduction of  $\text{Pb}_2\text{O}_3$  with  $\text{CH}_4$ .**

\_\_\_\_\_ [1]

- (ii) The student uses safety glasses and a lab coat.**

**State, with a reason, ONE other important safety precaution the student should take when carrying out this experiment.**

\_\_\_\_\_  
\_\_\_\_\_  
\_\_\_\_\_ [1]

- (iii) The student was not sure that all the oxygen had been removed from the lead oxide.

Suggest TWO modifications that the student could make to their method to be confident that all the oxygen had been removed. Explain your reasoning.

1 \_\_\_\_\_

\_\_\_\_\_

2 \_\_\_\_\_

\_\_\_\_\_

[2]

- (iv) The student makes suitable modifications to the method and repeats the experiment to obtain the accurate results shown below.

Mass of dish / g	8.364
Mass of dish + lead oxide / g	11.818
Mass of dish + lead at end of experiment / g	11.496

Calculate the empirical formula of the lead oxide.

empirical formula = \_\_\_\_\_ [2]

(b)  $\text{SiO}_2$  and  $\text{CO}_2$  are oxides of other Group 14 (Group 4) elements.

Solid  $\text{SiO}_2$  melts at  $2156^\circ\text{C}$ . Solid  $\text{CO}_2$  melts at  $-56^\circ\text{C}$ .

Suggest the type of lattice structure in solid  $\text{SiO}_2$  and in solid  $\text{CO}_2$  and explain the difference in melting points in terms of the types of force within each lattice structure.

Structure in  $\text{SiO}_2(\text{s})$  \_\_\_\_\_

Structure in  $\text{CO}_2(\text{s})$  \_\_\_\_\_

Explanation \_\_\_\_\_

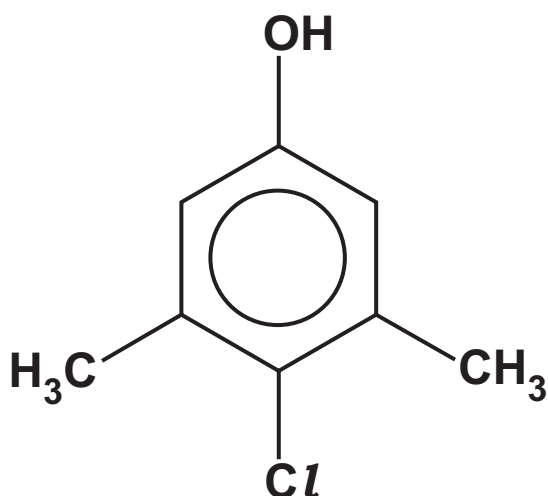
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\_\_\_\_\_ [4]

- 4 Dettol<sup>®</sup> is a disinfectant containing the antiseptic chloroxylenol, shown below.

chloroxylenol



- (a) Chloroxylenol is a weak Brønsted–Lowry acid.

- (i) What is the systematic name of chloroxylenol?

\_\_\_\_\_ [1]

- (ii) Predict the number of peaks in a  $^{13}\text{C}$  NMR spectrum of chloroxylenol.

\_\_\_\_\_ [1]

- (iii) Name the functional group responsible for the acidity of chloroxylenol and describe a simple test which would confirm the presence of this group.

Functional group \_\_\_\_\_

Test \_\_\_\_\_

\_\_\_\_\_  
\_\_\_\_\_ [2]

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- (iv) A student measures the pH of the contents in a bottle of Dettol<sup>®</sup> as 5.14.

The label on the bottle shows the percentage of chloroxylenol in Dettol<sup>®</sup> as 4.80% i.e. 100 cm<sup>3</sup> of Dettol<sup>®</sup> contains 4.80 g of chloroxylenol.

Assume the following:  
Chloroxylenol is the only acidic component in Dettol<sup>®</sup>.

Chloroxylenol is a weak monobasic acid.

The density of Dettol<sup>®</sup> is 1.00 g cm<sup>-3</sup>.

Write the equation, using molecular formulae, for the acid dissociation of chloroxylenol.

Calculate the acid dissociation constant,  $K_a$ , for chloroxylenol. [5]

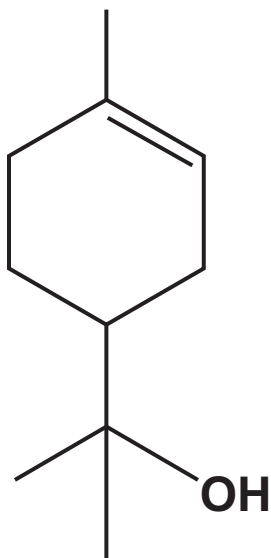
$$K_a = \text{_____} \text{ mol dm}^{-3}$$

(b) Dettol<sup>®</sup> contains other chemicals including  $\alpha$ -terpineol, shown below.

(i)  $\alpha$ -Terpineol is a chiral compound.

Show with an asterisk, (\*), the chiral centre(s) in the structure of  $\alpha$ -terpineol. [1]

$\alpha$ -terpineol



- (ii)  $\alpha$ -Terpineol meets the requirements for *E/Z* isomerism.  
However, only one *E/Z* isomer of  $\alpha$ -terpineol exists.

Explain

why  $\alpha$ -terpineol meets the requirements for *E/Z* isomerism

whether  $\alpha$ -terpineol is an *E*- or *Z*- isomer

why only one *E/Z* isomer of  $\alpha$ -terpineol exists.

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[4]

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**(iii)  $\alpha$ -Terpineol contains two functional groups.**

**For each functional group, choose a reagent that reacts with that group ONLY.**

**Draw the structures for the organic products of the reactions.**

**Show structures for organic compounds. [4]**

**Reagent(s) \_\_\_\_\_**

**Name of functional group that reacts \_\_\_\_\_**

**Structure of organic product**

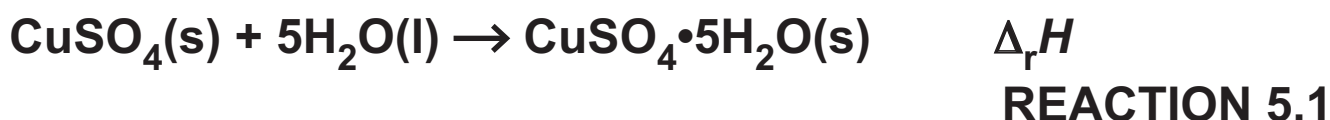
**Reagent(s)** \_\_\_\_\_

**Name of functional group that reacts** \_\_\_\_\_

**Structure of organic product**

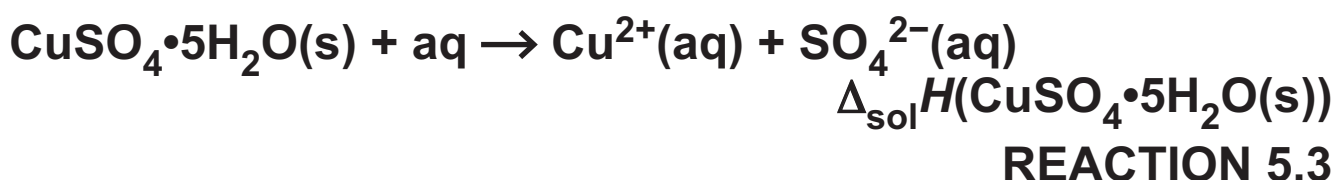
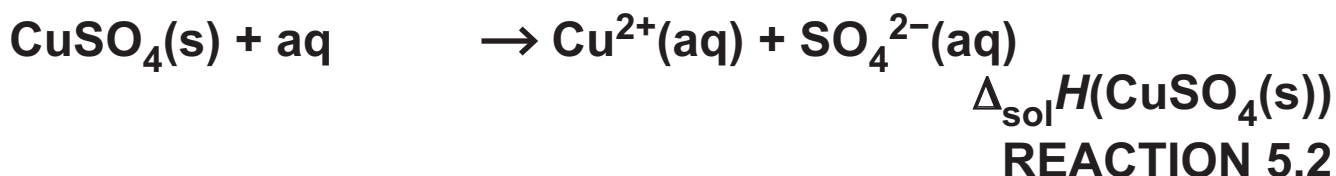
5 This question is about copper(II) sulfate,  $\text{CuSO}_4$ , and sodium thiosulfate,  $\text{Na}_2\text{S}_2\text{O}_3$ .

- (a) The enthalpy change of reaction,  $\Delta_r H$ , for converting anhydrous copper(II) sulfate to hydrated copper(II) sulfate is difficult to measure directly by experiment.



The enthalpy changes of solution of anhydrous and hydrated copper(II) sulfate can be measured by experiment. The reactions are shown below.

In the equations, 'aq' represents an excess of water.



## EXPERIMENT 1

A student carries out an experiment to find  $\Delta_{\text{sol}} H(\text{CuSO}_4(\text{s}))$  for REACTION 5.2.

## STUDENT'S METHOD

Weigh a bottle containing  $\text{CuSO}_4(\text{s})$  and weigh a polystyrene cup.

Add about  $50\text{ cm}^3$  of water to the polystyrene cup and measure its temperature.

Add the  $\text{CuSO}_4(\text{s})$ , stir the mixture, and measure the final temperature.

Weigh the empty bottle and weigh the polystyrene cup with final solution.

## MASS READINGS

Mass of bottle + $\text{CuSO}_4(\text{s})/\text{g}$	28.04
Mass of empty bottle/g	20.06
Mass of polystyrene cup/g	23.43
Mass of polystyrene cup + final solution/g	74.13

## TEMPERATURE READINGS

Initial temperature of water/ $^{\circ}\text{C}$	20.5
Temperature of final solution/ $^{\circ}\text{C}$	34.0

## EXPERIMENT 2

The student carries out a second experiment with  $\text{CuSO}_4 \cdot 5\text{H}_2\text{O}$  (REACTION 5.3). The student uses the same method as in EXPERIMENT 1.

The student calculates  $\Delta_{\text{sol}}H(\text{CuSO}_4 \cdot 5\text{H}_2\text{O}(\text{s}))$  as  $+8.43\text{ kJ mol}^{-1}$ .

**(i)\* Calculate  $\Delta_{\text{sol}}H(\text{CuSO}_4(\text{s}))$  for REACTION 5.2 and determine the enthalpy change of REACTION 5.1,  $\Delta_{\text{r}}H$ .**

**Assume that the specific heat capacity,  $c$ , of the solution is the same as for water.**

**Show your working, including an energy cycle linking the enthalpy changes. [6]**

[illegible]

**Additional answer space if required**

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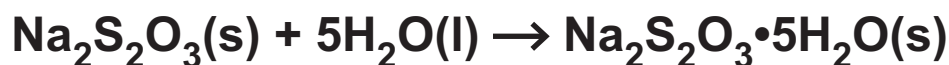
- (ii) The thermometer had an uncertainty in each temperature reading of  $\pm 0.1^\circ\text{C}$ .

The student calculates a 20% uncertainty in the temperature change in **EXPERIMENT 2**.

Calculate the temperature change in **EXPERIMENT 2**.

temperature change = \_\_\_\_\_  $^\circ\text{C}$  [1]

- (b) The standard enthalpy change of reaction,  $\Delta_r H^\ominus$ , and the standard free energy change,  $\Delta G^\ominus$ , for converting anhydrous sodium thiosulfate to hydrated sodium thiosulfate are shown below.



$$\Delta_r H^\ominus = -55.8 \text{ kJ mol}^{-1}$$

$$\Delta G^\ominus = -16.1 \text{ kJ mol}^{-1}$$

Standard entropies are given in the table.

Compound	$S^\ominus / \text{J K}^{-1} \text{ mol}^{-1}$
$\text{Na}_2\text{S}_2\text{O}_3 \cdot 5\text{H}_2\text{O}(\text{s})$	372.4
$\text{H}_2\text{O}(\text{l})$	69.9

Determine the STANDARD entropy,  $S^\ominus$ , of anhydrous sodium thiosulfate,  $\text{Na}_2\text{S}_2\text{O}_3(\text{s})$ .

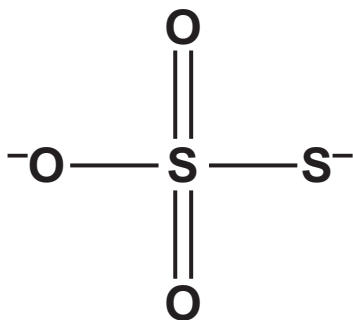
Give your answer to 3 significant figures.

$$S^\ominus = \underline{\hspace{2cm}} \text{ J K}^{-1} \text{ mol}^{-1} \text{ [4]}$$

- (c) Sodium thiosulfate contains the thiosulfate ion,  $\text{S}_2\text{O}_3^{2-}$ .

The displayed formula of  $\text{S}_2\text{O}_3^{2-}$  can be shown as below.

thiosulfate ion



- (i) Predict the O–S–S bond angle and name of the shape of the thiosulfate ion.

Bond angle \_\_\_\_\_

Name of shape \_\_\_\_\_ [1]

- (ii) In some of its reactions, the thiosulfate ion forms the tetrathionate ion,  $\text{S}_4\text{O}_6^{2-}$ .

The  $\text{S}_4\text{O}_6^{2-}$  ion is a 'dimer' of  $\text{S}_2\text{O}_3^{2-}$ .

Draw a displayed formula for the  $\text{S}_4\text{O}_6^{2-}$  ion.  
Use the space below. [1]

**6 This question is about reactions of iron compounds.**

**(a) A student carries out the reactions in the flowchart opposite, starting with iron(II) sulfide.**

**(i) In the boxes, write the formulae of A and B. [2]**

**(ii) The student thinks that the reaction of iron(II) sulfide with  $\text{H}^+/\text{H}_2\text{O}$  is a redox reaction.**

**Explain, with reasons, whether the student is correct.**

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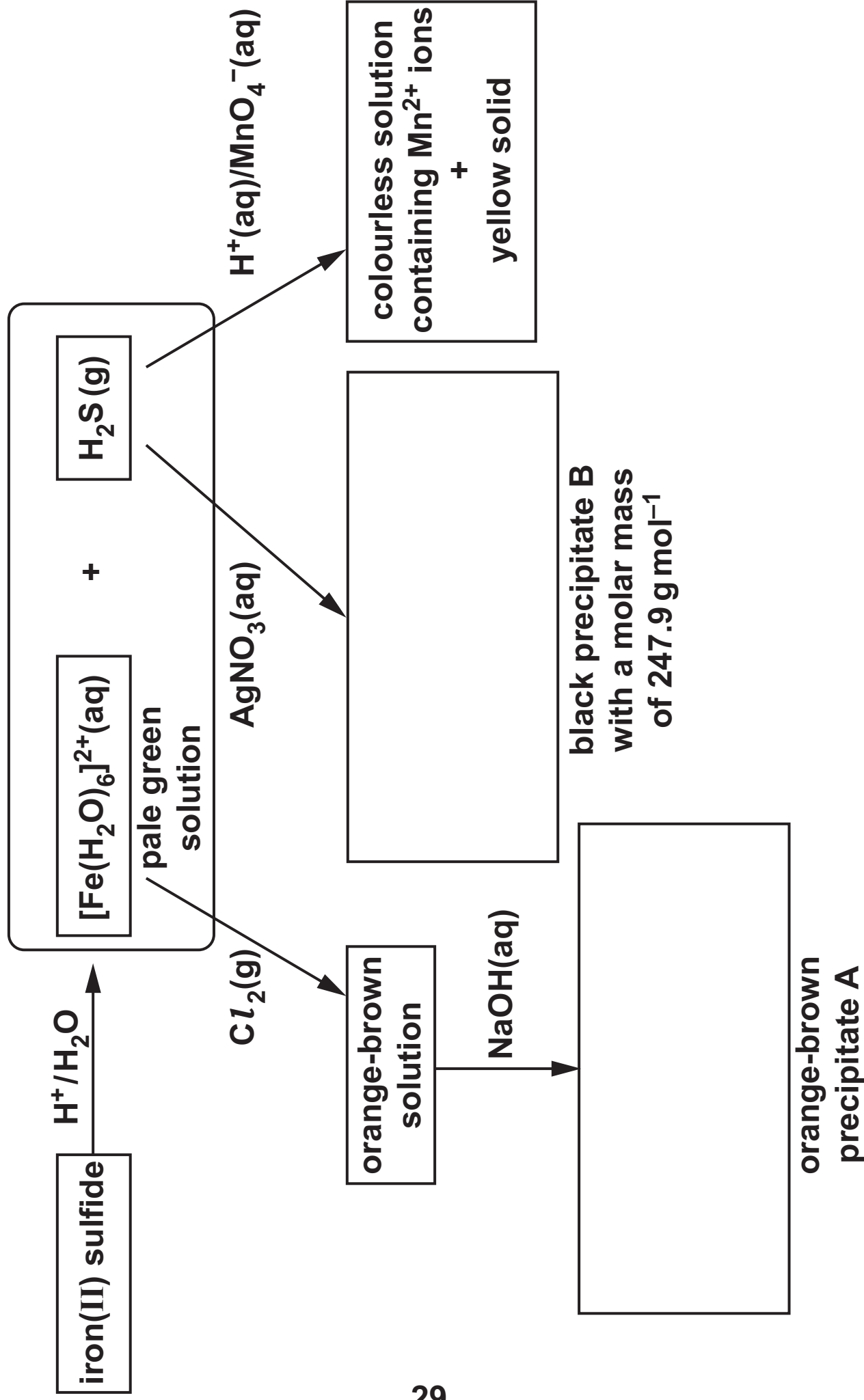
**[1]**

**(iii) Write the equation for the reaction of  $[\text{Fe}(\text{H}_2\text{O})_6]^{2+}(\text{aq})$  with  $\text{Cl}_2(\text{g})$ .**

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**[1]**

**(iv) Construct an equation for the reaction of  $\text{H}_2\text{S}(\text{g})$  with  $\text{H}^+(\text{aq})/\text{MnO}_4^-(\text{aq})$ . Use the space below. [2]**



**(b)\* Compound C is a hydrated ionic compound with the empirical formula:  $\text{FeH}_{18}\text{N}_3\text{O}_{18}$ .**

**A student investigates the thermal decomposition of compound C as outlined below.**

### **STAGE 1**

**The student gently heats 0.00300 mol of compound C to remove the water of crystallisation.**

**0.486 g of water is collected, leaving 0.00300 mol of the anhydrous compound D.**

### **STAGE 2**

**The student strongly heats 0.00300 mol of compound D, which decomposes to form a solid oxide E (molar mass of  $159.6 \text{ g mol}^{-1}$ ) and  $270 \text{ cm}^3$  of a gas mixture, measured at RTP, containing gases F and G.**

### **STAGE 3**

**The student cools the  $270 \text{ cm}^3$  gas mixture of F and G.**

**Gas F is a compound that condenses to form 0.414 g of a liquid.**

**Gas G remains and has a volume of  $54 \text{ cm}^3$ , measured at RTP.**

**Gas G is tested and it relights a glowing splint.**

**Determine the formulae of C, D, E, F and G.**

**Show all your working and equations for the reactions. [6]**

[illegible]

**Additional answer space if required.**

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**END OF QUESTION PAPER**

### ADDITIONAL ANSWER SPACE

**If additional space is required, you should use the following lined page(s). The question number(s) must be clearly shown in the margin(s).**

[illegible]



[illegible]

