

CANDIDATE  
NAME

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CENTRE  
NUMBER

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CANDIDATE  
NUMBER

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**CHEMISTRY**

**9701/33**

Paper 3 Advanced Practical Skills 1

**May/June 2018**

**2 hours**

Candidates answer on the Question Paper.

Additional Materials: As listed in the Confidential Instructions

**READ THESE INSTRUCTIONS FIRST**

Write your Centre number, candidate number and name on all the work you hand in.  
Give details of the practical session and laboratory where appropriate, in the boxes provided.  
Write in dark blue or black pen.  
You may use an HB pencil for any diagrams or graphs.  
Do not use staples, paper clips, glue or correction fluid.  
**DO NOT WRITE IN ANY BARCODES.**

Answer **all** questions.  
Electronic calculators may be used.  
You may lose marks if you do not show your working or if you do not use appropriate units.  
Use of a Data Booklet is unnecessary.

Qualitative Analysis Notes are printed on pages 10 and 11.  
A copy of the Periodic Table is printed on page 12.

At the end of the examination, fasten all your work securely together.  
The number of marks is given in brackets [ ] at the end of each question or part question.

<b>Session</b>	
<b>Laboratory</b>	

<b>For Examiner's Use</b>	
<b>1</b>	
<b>2</b>	
<b>3</b>	
<b>Total</b>	

This document consists of **11** printed pages and **1** blank page.

## Quantitative Analysis

Read through the whole method before starting any practical work. Where appropriate, prepare a table for your results in the space provided.

Show your working and appropriate significant figures in the final answer to **each** step of your calculations.

- 1 Acids are defined as substances that can donate hydrogen ions,  $\text{H}^+$ , to bases. Monoprotic acids contain one  $\text{H}^+$  that can be donated per molecule. Diprotic acids contain two  $\text{H}^+$  that can be donated per molecule.

You will determine by a titration method whether acid **Z** is monoprotic or diprotic.

**FA 1** is a solution containing  $6.10 \text{ g dm}^{-3}$  of acid **Z**.

**FA 2** is  $0.105 \text{ mol dm}^{-3}$  aqueous sodium hydroxide,  $\text{NaOH}$ .  
bromophenol blue indicator

### (a) Method

- Pipette  $25.0 \text{ cm}^3$  of **FA 1** into a conical flask.
- Fill a burette with **FA 2**.
- Add several drops of bromophenol blue indicator to the conical flask.
- Carry out a rough titration and record your burette readings in the space below.

The rough titre is .....  $\text{cm}^3$ .

- Carry out as many accurate titrations as you think necessary to obtain consistent results.
- Make sure any recorded results show the accuracy of your practical work.
- Record, in a suitable form below, all of your burette readings and the volume of **FA 2** added in each accurate titration.

I	
II	
III	
IV	
V	
VI	
VII	

[7]

- (b) From your accurate titration results, obtain a suitable value for the volume of **FA 2** to be used in your calculations. Show clearly how you obtained this value.

25.0 cm<sup>3</sup> of **FA 1** required ..... cm<sup>3</sup> of **FA 2**. [1]

(c) **Calculations**

- (i) Calculate the number of moles of sodium hydroxide present in the volume of **FA 2** calculated in (b).

moles of NaOH = ..... mol

Then deduce the number of moles of H<sup>+</sup> present in 25.0 cm<sup>3</sup> of **FA 1**.

moles of H<sup>+</sup> in 25.0 cm<sup>3</sup> of **FA 1** = ..... mol [1]

- (ii) Calculate the number of moles of H<sup>+</sup> present in 1 dm<sup>3</sup> of **FA 1**.

moles of H<sup>+</sup> in 1 dm<sup>3</sup> of **FA 1** = ..... mol [1]

- (iii) **FA 1** contains 6.10 g dm<sup>-3</sup> of acid **Z**. The relative molecular mass of **Z** is 126.

Calculate the number of moles of **Z** in 1 dm<sup>3</sup> of **FA 1**.

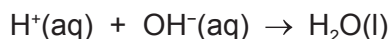
moles of **Z** in 1 dm<sup>3</sup> of **FA 1** = ..... mol [1]

- (iv) Use your answers to (ii) and (iii) to determine whether **Z** is a monoprotic or a diprotic acid. Explain your answer.

.....  
 ..... [1]

[Total: 12]

- 2 When an acid reacts with an alkali the neutralisation reaction is always exothermic.



You will determine the enthalpy change of neutralisation,  $\Delta H$ , for a monoprotic acid **X**.

**FA 3** is aqueous sodium hydroxide, NaOH.

**FA 4** is a  $2.00 \text{ mol dm}^{-3}$  solution of monoprotic acid **X**.

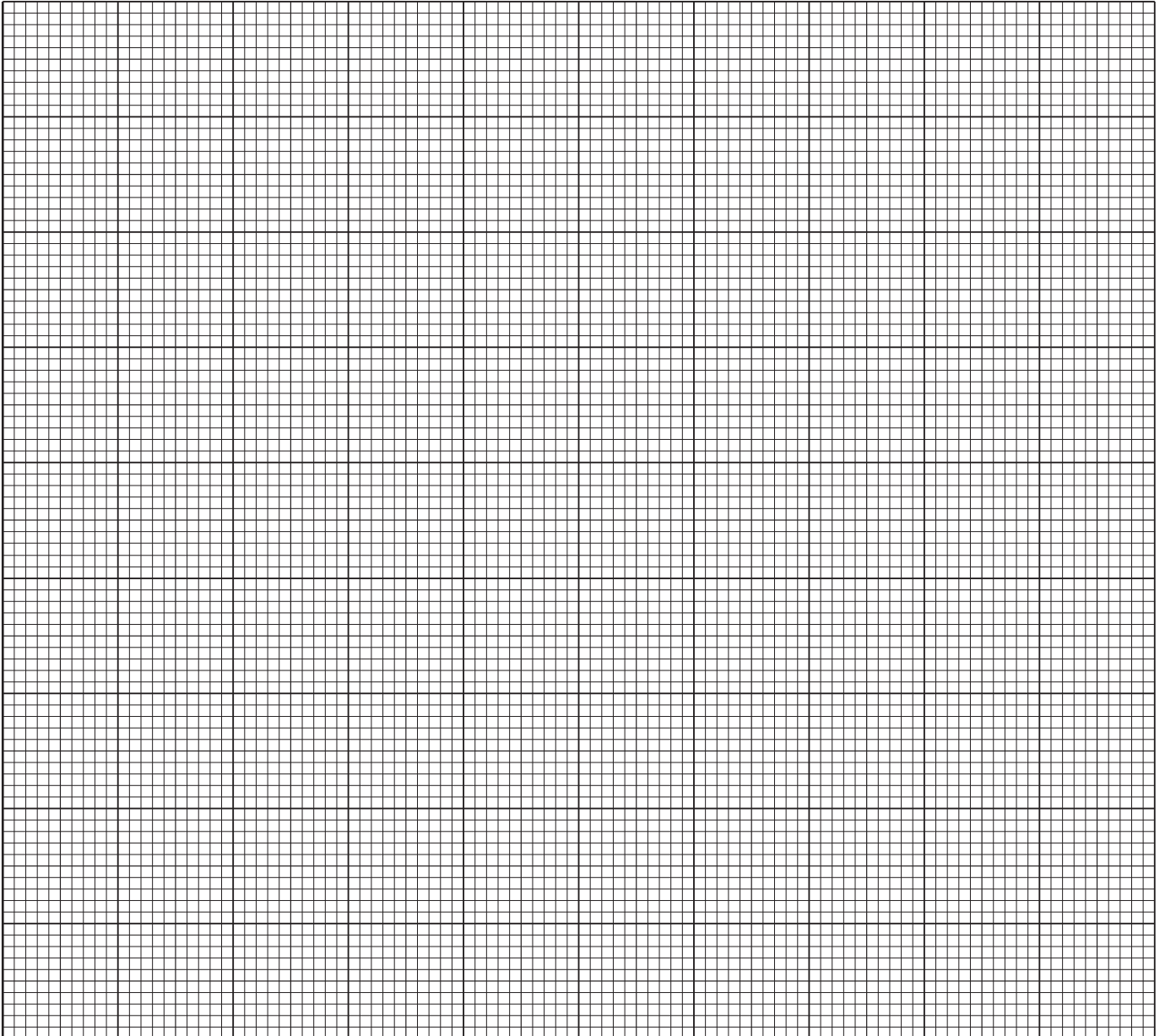
**(a) Method**

- Support the plastic cup in the  $250 \text{ cm}^3$  beaker.
- Fill the second burette with **FA 3**.
- Use the measuring cylinder to transfer  $25.0 \text{ cm}^3$  of **FA 4** into the plastic cup.
- Measure and record the temperature of **FA 4** in the plastic cup.
- Add  $5.00 \text{ cm}^3$  of **FA 3** from the burette into the plastic cup. Stir the contents of the cup. Read and record the maximum temperature of the solution.
- Add a further  $5.00 \text{ cm}^3$  of **FA 3** from the burette into the cup. Stir the contents of the cup. Read and record the maximum temperature of the solution.
- Repeat the addition of **FA 3**, in  $5.00 \text{ cm}^3$  portions, until  $50.00 \text{ cm}^3$  have been added. Read and record the maximum temperature of the solution after each addition.

I	
II	
III	
IV	

[4]

- (b) On the grid plot a graph of temperature, ( $y$ -axis), against volume of **FA 3** added, ( $x$ -axis). Your scale should allow a temperature of  $2^{\circ}\text{C}$  above the maximum measured to be plotted.



On your graph draw two lines of best fit. One line should be for when the temperature was rising and the other for after the maximum temperature had been reached. You should indicate clearly any points you consider to be anomalous.

Extrapolate the lines so that they intersect.

[4]

- (c) From your graph, determine the maximum temperature reached in the experiment and the volume of **FA 3** added to produce this maximum temperature.

maximum temperature reached = ..... °C

volume of **FA 3** added to reach maximum temperature = ..... cm<sup>3</sup> [1]

- (d) (i) Calculate the energy released during this experiment.

[Assume that 4.2J of heat energy changes the temperature of 1.0cm<sup>3</sup> of solution by 1.0°C.]

energy released = ..... J [1]

- (ii) Calculate the number of moles of acid **X** in 25.0cm<sup>3</sup> of **FA 4**.

moles of **X** = ..... mol [1]

- (iii) Calculate, in kJ mol<sup>-1</sup>, the enthalpy change of neutralisation for acid **X**.

enthalpy change = ..... kJ mol<sup>-1</sup> [1]  
(sign) (value)

- (e) Without changing the apparatus or solutions used, suggest **one** way in which the experiment could be modified to make the values obtained in (c) more accurate.

.....  
..... [1]

- (f) The enthalpy change of neutralisation of hydrochloric acid with aqueous sodium hydroxide is more exothermic than the enthalpy change of neutralisation of acid **X**.

Explain what this tells you about acid **X**.

.....  
.....  
.....  
..... [2]

[Total: 15]

### Qualitative Analysis

Where reagents are selected for use in a test, the **name** or **correct formula** of the element or compound must be given.

At each stage of any test you are to record details of the following:

- colour changes seen;
- the formation of any precipitate and its solubility in an excess of the reagent added;
- the formation of any gas and its identification by a suitable test.

You should indicate clearly at what stage in a test a change occurs.

If any solution is warmed, a **boiling tube** must be used.

Rinse and reuse test-tubes and boiling tubes where possible.

**No additional tests for ions present should be attempted.**

- 3 (a) **FA 5**, **FA 6** and **FA 7** are  $1.0 \text{ mol dm}^{-3}$  sulfuric acid,  $0.1 \text{ mol dm}^{-3}$  sulfuric acid and  $1.0 \text{ mol dm}^{-3}$  hydrochloric acid but not necessarily in that order.

You are to plan and carry out tests to determine the identities of **FA 5**, **FA 6** and **FA 7**. You should record, in a suitable form in the space below, your tests and observations. You should show clearly the observations for each of **FA 5**, **FA 6** and **FA 7** with all test reagents.

**FA 5** is .....

**FA 6** is .....

**FA 7** is .....

[4]

(b) **FA 8** and **FA 9** each contain one anion and one cation.

(i) Carry out the following tests and record your observations.

For each test you should use a **small** spatula measure of **FA 8** or **FA 9**.

test	observations	
	FA 8	FA 9
Add a small spatula measure to a 1 cm depth of distilled water in a test-tube and shake. Add a few drops of universal indicator. Record the pH of the mixture.		
Heat a small spatula measure in a hard-glass test-tube.		
Dissolve a small spatula measure in a 2 cm depth of dilute hydrochloric acid in a boiling tube (heat gently if necessary). Place a 1 cm depth of this solution into two test-tubes.	X	X
To one test-tube, add aqueous sodium hydroxide.		
To the second test-tube, add aqueous ammonia.		

[6]

(ii) Mix a spatula measure of **FA 8** with a spatula measure of **FA 9**. Heat the mixture in a hard-glass test-tube. Record your observations.

.....  
 .....  
 ..... [1]

(iii) From your observations in (b)(i) and (b)(ii) identify the ions present in **FA 8** and **FA 9**. If you are unable to identify an ion, write 'unknown'.

**FA 8** cation ..... anion .....

**FA 9** cation ..... anion .....

[2]

[Total: 13]





## Qualitative Analysis Notes

## 1 Reactions of aqueous cations

ion	reaction with	
	NaOH(aq)	NH <sub>3</sub> (aq)
aluminium, Al <sup>3+</sup> (aq)	white ppt. soluble in excess	white ppt. insoluble in excess
ammonium, NH <sub>4</sub> <sup>+</sup> (aq)	no ppt. ammonia produced on heating	–
barium, Ba <sup>2+</sup> (aq)	faint white ppt. is nearly always observed unless reagents are pure	no ppt.
calcium, Ca <sup>2+</sup> (aq)	white ppt. with high [Ca <sup>2+</sup> (aq)]	no ppt.
chromium(III), Cr <sup>3+</sup> (aq)	grey-green ppt. soluble in excess	grey-green ppt. insoluble in excess
copper(II), Cu <sup>2+</sup> (aq)	pale blue ppt. insoluble in excess	blue ppt. soluble in excess giving dark blue solution
iron(II), Fe <sup>2+</sup> (aq)	green ppt. turning brown on contact with air insoluble in excess	green ppt. turning brown on contact with air insoluble in excess
iron(III), Fe <sup>3+</sup> (aq)	red-brown ppt. insoluble in excess	red-brown ppt. insoluble in excess
magnesium, Mg <sup>2+</sup> (aq)	white ppt. insoluble in excess	white ppt. insoluble in excess
manganese(II), Mn <sup>2+</sup> (aq)	off-white ppt. rapidly turning brown on contact with air insoluble in excess	off-white ppt. rapidly turning brown on contact with air insoluble in excess
zinc, Zn <sup>2+</sup> (aq)	white ppt. soluble in excess	white ppt. soluble in excess

## 2 Reactions of anions

<i>ion</i>	<i>reaction</i>
carbonate, $\text{CO}_3^{2-}$	$\text{CO}_2$ liberated by dilute acids
chloride, $\text{Cl}^-(\text{aq})$	gives white ppt. with $\text{Ag}^+(\text{aq})$ (soluble in $\text{NH}_3(\text{aq})$ )
bromide, $\text{Br}^-(\text{aq})$	gives cream ppt. with $\text{Ag}^+(\text{aq})$ (partially soluble in $\text{NH}_3(\text{aq})$ )
iodide, $\text{I}^-(\text{aq})$	gives yellow ppt. with $\text{Ag}^+(\text{aq})$ (insoluble in $\text{NH}_3(\text{aq})$ )
nitrate, $\text{NO}_3^-(\text{aq})$	$\text{NH}_3$ liberated on heating with $\text{OH}^-(\text{aq})$ and <i>Al</i> foil
nitrite, $\text{NO}_2^-(\text{aq})$	$\text{NH}_3$ liberated on heating with $\text{OH}^-(\text{aq})$ and <i>Al</i> foil
sulfate, $\text{SO}_4^{2-}(\text{aq})$	gives white ppt. with $\text{Ba}^{2+}(\text{aq})$ (insoluble in excess dilute strong acids)
sulfite, $\text{SO}_3^{2-}(\text{aq})$	gives white ppt. with $\text{Ba}^{2+}(\text{aq})$ (soluble in excess dilute strong acids)

## 3 Tests for gases

<i>gas</i>	<i>test and test result</i>
ammonia, $\text{NH}_3$	turns damp red litmus paper blue
carbon dioxide, $\text{CO}_2$	gives a white ppt. with limewater (ppt. dissolves with excess $\text{CO}_2$ )
chlorine, $\text{Cl}_2$	bleaches damp litmus paper
hydrogen, $\text{H}_2$	'pops' with a lighted splint
oxygen, $\text{O}_2$	relights a glowing splint

## The Periodic Table of Elements

		Group															
1	2	3	4	5	6	7	8	9	10	11	12	13	14	15	16	17	18
		<div style="display: flex; justify-content: space-between; align-items: center;"> <div style="border: 1px solid black; padding: 2px;">1 H hydrogen 1.0</div> <div style="border: 1px solid black; padding: 2px;">2 He helium 4.0</div> </div>															
		<div style="display: flex; justify-content: space-between; align-items: center;"> <div style="border: 1px solid black; padding: 2px;">3 Li lithium 6.9</div> <div style="border: 1px solid black; padding: 2px;">4 Be beryllium 9.0</div> </div>															
		<div style="display: flex; justify-content: space-between; align-items: center;"> <div style="border: 1px solid black; padding: 2px;">5 B boron 10.8</div> <div style="border: 1px solid black; padding: 2px;">6 C carbon 12.0</div> <div style="border: 1px solid black; padding: 2px;">7 N nitrogen 14.0</div> <div style="border: 1px solid black; padding: 2px;">8 O oxygen 16.0</div> <div style="border: 1px solid black; padding: 2px;">9 F fluorine 19.0</div> </div>															
		<div style="display: flex; justify-content: space-between; align-items: center;"> <div style="border: 1px solid black; padding: 2px;">11 Na sodium 23.0</div> <div style="border: 1px solid black; padding: 2px;">12 Mg magnesium 24.3</div> </div>															
		<div style="display: flex; justify-content: space-between; align-items: center;"> <div style="border: 1px solid black; padding: 2px;">13 Al aluminium 27.0</div> <div style="border: 1px solid black; padding: 2px;">14 Si silicon 28.1</div> <div style="border: 1px solid black; padding: 2px;">15 P phosphorus 31.0</div> <div style="border: 1px solid black; padding: 2px;">16 S sulfur 32.1</div> <div style="border: 1px solid black; padding: 2px;">17 Cl chlorine 35.5</div> </div>															
		<div style="display: flex; justify-content: space-between; align-items: center;"> <div style="border: 1px solid black; padding: 2px;">19 K potassium 39.1</div> <div style="border: 1px solid black; padding: 2px;">20 Ca calcium 40.1</div> </div>															
		<div style="display: flex; justify-content: space-between; align-items: center;"> <div style="border: 1px solid black; padding: 2px;">21 Sc scandium 45.0</div> <div style="border: 1px solid black; padding: 2px;">22 Ti titanium 47.9</div> <div style="border: 1px solid black; padding: 2px;">23 V vanadium 50.9</div> <div style="border: 1px solid black; padding: 2px;">24 Cr chromium 52.0</div> <div style="border: 1px solid black; padding: 2px;">25 Mn manganese 54.9</div> <div style="border: 1px solid black; padding: 2px;">26 Fe iron 55.8</div> <div style="border: 1px solid black; padding: 2px;">27 Co cobalt 58.9</div> <div style="border: 1px solid black; padding: 2px;">28 Ni nickel 58.7</div> <div style="border: 1px solid black; padding: 2px;">29 Cu copper 63.5</div> <div style="border: 1px solid black; padding: 2px;">30 Zn zinc 65.4</div> </div>															
		<div style="display: flex; justify-content: space-between; align-items: center;"> <div style="border: 1px solid black; padding: 2px;">37 Rb rubidium 85.5</div> <div style="border: 1px solid black; padding: 2px;">38 Sr strontium 87.6</div> </div>															
		<div style="display: flex; justify-content: space-between; align-items: center;"> <div style="border: 1px solid black; padding: 2px;">39 Y yttrium 88.9</div> <div style="border: 1px solid black; padding: 2px;">40 Zr zirconium 91.2</div> <div style="border: 1px solid black; padding: 2px;">41 Nb niobium 92.9</div> <div style="border: 1px solid black; padding: 2px;">42 Mo molybdenum 95.9</div> <div style="border: 1px solid black; padding: 2px;">43 Tc technetium —</div> <div style="border: 1px solid black; padding: 2px;">44 Ru ruthenium 101.1</div> <div style="border: 1px solid black; padding: 2px;">45 Rh rhodium 102.9</div> <div style="border: 1px solid black; padding: 2px;">46 Pd palladium 106.4</div> <div style="border: 1px solid black; padding: 2px;">47 Ag silver 107.9</div> <div style="border: 1px solid black; padding: 2px;">48 Cd cadmium 112.4</div> </div>															
		<div style="display: flex; justify-content: space-between; align-items: center;"> <div style="border: 1px solid black; padding: 2px;">55 Cs caesium 132.9</div> <div style="border: 1px solid black; padding: 2px;">56 Ba barium 137.3</div> </div>															
		<div style="display: flex; justify-content: space-between; align-items: center;"> <div style="border: 1px solid black; padding: 2px;">57–71 lanthanoids</div> <div style="border: 1px solid black; padding: 2px;">72 Hf hafnium 178.5</div> <div style="border: 1px solid black; padding: 2px;">73 Ta tantalum 180.9</div> <div style="border: 1px solid black; padding: 2px;">74 W tungsten 183.8</div> <div style="border: 1px solid black; padding: 2px;">75 Re rhenium 186.2</div> <div style="border: 1px solid black; padding: 2px;">76 Os osmium 190.2</div> <div style="border: 1px solid black; padding: 2px;">77 Ir iridium 192.2</div> <div style="border: 1px solid black; padding: 2px;">78 Pt platinum 195.1</div> <div style="border: 1px solid black; padding: 2px;">79 Au gold 197.0</div> <div style="border: 1px solid black; padding: 2px;">80 Hg mercury 200.6</div> </div>															
		<div style="display: flex; justify-content: space-between; align-items: center;"> <div style="border: 1px solid black; padding: 2px;">87 Fr francium —</div> <div style="border: 1px solid black; padding: 2px;">88 Ra radium —</div> </div>															
		<div style="display: flex; justify-content: space-between; align-items: center;"> <div style="border: 1px solid black; padding: 2px;">89–103 actinoids</div> <div style="border: 1px solid black; padding: 2px;">104 Rf rutherfordium —</div> <div style="border: 1px solid black; padding: 2px;">105 Db dubnium —</div> <div style="border: 1px solid black; padding: 2px;">106 Sg seaborgium —</div> <div style="border: 1px solid black; padding: 2px;">107 Bh bohrium —</div> <div style="border: 1px solid black; padding: 2px;">108 Hs hassium —</div> <div style="border: 1px solid black; padding: 2px;">109 Mt meitnerium —</div> <div style="border: 1px solid black; padding: 2px;">110 Ds darmstadtium —</div> <div style="border: 1px solid black; padding: 2px;">111 Rg roentgenium —</div> <div style="border: 1px solid black; padding: 2px;">112 Cn copernicium —</div> </div>															

lanthanoids	57 La lanthanum 138.9	58 Ce cerium 140.1	59 Pr praseodymium 140.9	60 Nd neodymium 144.4	61 Pm promethium —	62 Sm samarium 150.4	63 Eu europium 152.0	64 Gd gadolinium 157.3	65 Tb terbium 158.9	66 Dy dysprosium 162.5	67 Ho holmium 164.9	68 Er erbium 167.3	69 Tm thulium 168.9	70 Yb ytterbium 173.1	71 Lu lutetium 175.0
actinoids	89 Ac actinium —	90 Th thorium 232.0	91 Pa protactinium 231.0	92 U uranium 238.0	93 Np neptunium —	94 Pu plutonium —	95 Am americium —	96 Cm curium —	97 Bk berkelium —	98 Cf californium —	99 Es einsteinium —	100 Fm fermium —	101 Md mendelevium —	102 No nobelium —	103 Lr lawrencium —