

UNIVERSITY OF CAMBRIDGE INTERNATIONAL EXAMINATIONS

GCE Advanced Subsidiary Level and GCE Advanced Level

**MARK SCHEME for the May/June 2011 question paper
for the guidance of teachers**

9701 CHEMISTRY

9701/32

Paper 32 (Advanced Practical Skills 2),
maximum raw mark 40

This mark scheme is published as an aid to teachers and candidates, to indicate the requirements of the examination. It shows the basis on which Examiners were instructed to award marks. It does not indicate the details of the discussions that took place at an Examiners' meeting before marking began, which would have considered the acceptability of alternative answers.

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UNIVERSITY of CAMBRIDGE
International Examinations

Page 2	Mark Scheme: Teachers' version	Syllabus	Paper
	GCE AS/A LEVEL – May/June 2011	9701	32

Question	Sections	Indicative material	Mark
1 (a)	PDO Layout	I Volume given for Rough titre and accurate titre details tabulated. <i>Minimum of 2 × 2 boxes.</i>	1
	MMO Collection	II Initial and final (burette) (readings) and volume of FB 2 added/reading at start and finish recorded for each accurate titre (not 'difference'). and mass tube + FB 1 , mass tube + residue/empty, mass FB 1 . Ignore units. <i>Headings should match readings.</i> <i>Do not award this mark if:</i> <i>50(.00) is used as an initial burette reading;</i> <i>More than one final burette reading is 50(.00);</i> <i>Any burette reading is greater than 50(.00).</i>	1
	PDO Recording	III All accurate burette readings (initial and final) recorded to nearest 0.05 cm ³ . <i>Assessed on burette readings only (minimum of 2 readings).</i>	1
	MMO Decisions	IV Has two uncorrected accurate titres within 0.1 cm ³ . <i>Do not award this mark if, having performed two titres within 0.1 cm³, a further titration is performed that is more than 0.10 cm³ from the closer of the initial two titres, unless a fourth titre, within 0.1 cm³ of any of the previous titres, has also been carried out.</i>	1
<p>Round any burette readings to the nearest 0.05 cm³.</p> <p>Check and correct, if necessary, subtractions in the titre table and in the calculation of mass.</p> <p>Examiner then selects the 'best' titre using the hierarchy: two identical; titres within 0.05 cm³, titres within 0.1 cm³ etc.</p> <p>Calculate: candidate's titre × $\frac{\text{Supervisor mass}}{\text{candidate mass}}$ to 2 decimal places</p> <p>Calculate difference in Supervisor and candidate scaled values and award quality marks as below.</p>			
	MMO Quality	<p>V, VI and VII</p> <p>Award V, VI and VII if $\delta \leq 0.25 \text{ cm}^3$</p> <p>Award V and VI if $0.25 < \delta \leq 0.50 \text{ cm}^3$</p> <p>Award V if $0.50 < \delta \leq 0.80 \text{ cm}^3$</p> <p><i>If the 'best' titres are $\geq 0.60 \text{ cm}^3$ apart cancel one of the Q marks.</i></p>	3

[7]

Page 4	Mark Scheme: Teachers' version	Syllabus	Paper
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2 (a)	PDO Layout	I Two balance readings, one mass, two thermometer readings and one change in temperature shown in suitable layout.	1	
	PDO Recording	II Masses and temperatures recorded with correct headings and units for all data shown. <i>Acceptable units for temperature are /°C, (°C), temperature in degrees Celsius, temperature in °C., units for mass are /g, (g), mass in grams.</i>	1	
	PDO Recording	III All thermometer readings recorded to 0.0°C or 0.5°C and all balance readings recorded to same degree of accuracy.	1	
<p>Round all thermometer readings to nearest 0.5°C. Check and correct, if necessary, subtractions in the temperature change and the mass used.</p> <p>Calculate to 1 decimal place: candidate temperature change $\times \frac{\text{Supervisor mass}}{\text{candidate mass used}}$</p> <p>Calculate difference in candidate and Supervisor scaled values and award quality marks as below.</p>				
	MMO Quality	IV and V Award IV and V for changes within 0.8°C of Supervisor Award V for changes > 0.8 but within 1.6°C of Supervisor	2	[5]
(b) (i)	ACE Interpretation	I Expression for heat change in (i) = $25 \times 4.3 \times \text{temperature change from (a)}$ (answer given must correspond to units quoted).	1	
		II Expression for moles of washing soda from mass used and M_r from (a) or $M_r = 259$ or $M_r = 286$ in (ii)	1	
		III Correctly evaluates enthalpy change = heat change / (1000 \times moles of washing soda) in (iii) (if 1000 not used, must say J).	1	
	ACE Conclusions	IV Enthalpy change shown as positive and to 3 sig figs. (Answer need not be arithmetically correct). Ignore sig figs (except if approximated to 1 sig fig in rest of question.)	1	[4]
(c)	ACE Improvements	Use a more precise thermometer/a thermometer with more accurate calibrations/a thermometer that reads to 0.1°C or 0.2°C (a more accurate thermometer/a digital thermometer/thermocouple is insufficient) or use a more precise method to measure the volume of acid or use a deeper plastic cup or scaling up apparatus and quantities of chemicals used (Do not accept 'add a lid')	1	[1]
			[Total: 10]	

Page 5	Mark Scheme: Teachers' version	Syllabus	Paper
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FB 5 is $\text{MgSO}_4(\text{aq})$; FB 6 is $\text{Pb}(\text{NO}_3)_2(\text{aq})$ FB 7 is $\text{Al}_2(\text{SO}_4)_3(\text{aq})$; FB 8 is $(\text{NH}_4)_2\text{FeSO}_4(\text{aq})$			
3	(a)	(i)	<p>MMO Decisions</p> <p>MMO Collection</p> <p>I Reagents chosen $\text{KI}(\text{aq})$ or $\text{HCl}(\text{aq})$ or K_2CrO_4 or $\text{K}_2\text{Cr}_2\text{O}_7$ or H_2SO_4 and $\text{NaOH}(\text{aq})$ (penalise additional reagents) 1</p> <p>II NaOH white precipitates for all 1</p> <p>III Excess NaOH no effect FB 5, precipitate dissolves FB 6 and FB 7 1</p> <p>IV $\text{KI} / \text{HCl} / \text{K}_2\text{CrO}_4 / \text{K}_2\text{Cr}_2\text{O}_7 / \text{H}_2\text{SO}_4$ nothing/no visible reaction for (FB 5 and FB 7), yellow precipitate/white precipitate for FB 6. 1</p> <p>Ignore observations for additional reagents. [4]</p>
		(ii)	<p>ACE Conclusions</p> <p>I FB 5 contains Mg^{2+}, FB 6 contains Pb^{2+} and FB 7 contains Al^{3+} (no ecf and must follow observations in (i)) 1</p> <p>II FB 5 (white) precipitate with NaOH, insoluble in excess 1</p> <p>III FB 6 (yellow) precipitate with $\text{KI} /$ (yellow) precipitate with K_2CrO_4 or $\text{K}_2\text{Cr}_2\text{O}_7 /$ (white) precipitate with HCl or H_2SO_4. 1</p> <p>FB 7 No precipitate with $\text{KI} / \text{HCl} / \text{H}_2\text{SO}_4$ and (white) precipitate with NaOH, soluble in excess. (Both observations needed unless FB 6 already identified as Pb^{2+}). 1</p> <p>Allow ecf, based on candidate's observations, for II, III and IV. [4]</p>

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(b)	(i)	MMO Collection	Effervescence/bubbles/hydrogen produced (ignore any test for ammonia but tests for other gases negate). (Do not accept gas produced) or Black/grey solid/coating on magnesium	1	
	(ii)		Ammonia/gas turns litmus paper blue	1	
	(iii)		Green precipitate (any qualified green including grey/green but do not allow green/brown.)	1	
			Turns brown (any qualified brown) on addition of hydrogen peroxide. Allow rusty or orange/brown precipitate but not orange alone. Ignore effervescence.	1	
	ACE Conclusions		Fe ²⁺ / iron (II).	1	[5]
			(+) ² to 0 (ecf on chromium (+) ³ to 0) or (+) ³ to (+) ² .	1	
		(+) ² to (+) ³ .	1		
		Conclusions are free standing but must be Fe ²⁺ .		[2]	
				[Total: 15]	