



Rewarding Learning

General Certificate of Education
2011

Chemistry

Assessment Unit A2 3
Internal Assessment
Practical Examination 2

[AC232]

THURSDAY 19 MAY, MORNING

MARK SCHEME

Annotation

1. Please do all marking in red ink.
2. All scripts are checked for mathematical errors. Please adopt the system of one tick (✓) equals [1] mark e.g. if you have awarded 4 marks for part of a question then 4 ticks (✓) should be on this candidate's answer.
3. As candidates have access to scripts please do not write any inappropriate comments on their scripts.

- 1 (a) Rinse burette and pipette with appropriate solution [1] place 25.0 cm³ (or specified volume) in conical flask using a pipette filler [1] add acid (20 cm³ of dilute sulfuric acid) [1] swirl flask during test [1] add permanganate solution quickly at first [1] then dropwise [1] do a rough titration [1] do 2(3) accurate titrations/until concordant/±0.1 cm³ until end-point/colour change [1]

maximum of [6]

(b)

Titration	Volume at start/cm ³	Volume at end/cm ³	Added volume/cm ³
1			
2			
3			

Average titre = cm³

Clearly labelled and neatly presented table, including units and the recording of the average titre. [1]

Significant figures: All accurate titration readings recorded to one decimal place (including initial burette reading a 0.0 if used). Accept, however, 0.00 and 0.05 but penalise by [-1] if other readings are given to two or more decimal places. The use of 0 is penalised by [-1] if used (only penalise once). [2]

Titration consistency: This is the difference between the accurate readings.

Difference	Mark
±0.1	2
±0.2	1
>0.3	0

[2]

2 (a)

Test	Observation	Deduction
1 Describe the appearance of X.	green solid/crystals [1]	transition metal compound/nickel ion or compound/hydrate or water of crystallisation/crystals stated [1]
2 Add a spatula measure of X to approximately 50 cm ³ of water.	dissolves/soluble/disappears [1] blue-green/green [1] solution [1] Max. [2]	<u>soluble</u> /confirms transition metal compound/nickel [1] award only once
3 Add 10 drops of silver nitrate solution to 2 cm ³ of the solution of X in a test tube. Allow to stand.	white [1] precipitate [1]	chloride [1]
4 Add 5 drops of sodium hydroxide solution to 2 cm ³ of the solution of X in a test tube.	green [1] precipitate [1]	insoluble hydroxide [1] Na(OH) ₂ etc.
5 In a fume cupboard, add 6 cm ³ of concentrated ammonia, slowly, to 2 cm ³ of the solution of X in a test tube.	blue [1] solution [1]	No deduction required
6 Add 2 cm ³ of edta solution to 2 cm ³ of the solution of X in a test tube.	blue solution [1]	No deduction required

X is (hydrated) nickel chloride/nickel(II) chloride [1]

AVAILABLE MARKS

(b)	Test	Observation	Decision
	1 Describe the appearance of Y.	white solid crystals [1]	high RMM/long strong (intermolecular forces) named white organic solid no chromophore glycine or alanine [1]
	2 Heat one spatula measure of Y in a test tube. Heat gently at first and then more strongly. Test any fumes with a glass rod dipped in concentrated hydrochloric acid.	White smoke	(smoke is) NH_4Cl molecules contain N/ammonia [1]
	3 (i) Dissolve 2 spatula measures of Y in approximately 20 cm^3 of water. (ii) Use Universal Indicator paper to determine the pH of the solution of Y.	green colour [1]	pH 6/7 [1]
	4 Add 6 drops of copper(II) sulfate solution, dropwise, to a test tube half-full of a solution of Y.	blue colour [1]	contains – NH_2 /– NH /glycine/ any amino acid/correct comment on complex formed with amino acid [1]
	5 Add 3 cm^3 of acidified potassium dichromate(VI) solution. Add one spatula measure of Y and warm gently.	No change/stays orange [1]	not a primary or secondary alcohol/ could be – COOH – NH_2 – tert alcohol – ketone dichromate not reduced [1]

Identity of Y: glycine [1]

Maximum [25]

AVAILABLE
MARKS

25

3 (a) (i) $\text{Na}_2\text{SO}_3(\text{aq}) + \text{S}(\text{s}) \rightarrow \text{Na}_2\text{S}_2\text{O}_3(\text{aq})$ [1]
state symbols dependent on correct formula [1]

(ii) $\text{Na}_2\text{S}_2\text{O}_3 \cdot 5\text{H}_2\text{O}$ [1]

(b) $\text{Na}_2\text{S}_2\text{O}_3 \cdot 5\text{H}_2\text{O} = 2 \times 23 + 2 \times 32 + 3 \times 16 + 5 \times 18$
 $= 46 \quad 64 \quad 48 \quad 90$
 $= 248$

$2.5\text{g} = 2.5/248 = 0.01\text{ mol}$

Volume of 0.06 mol dm^{-3} sulfate (IV)

$V \times 0.06 = 0.01$

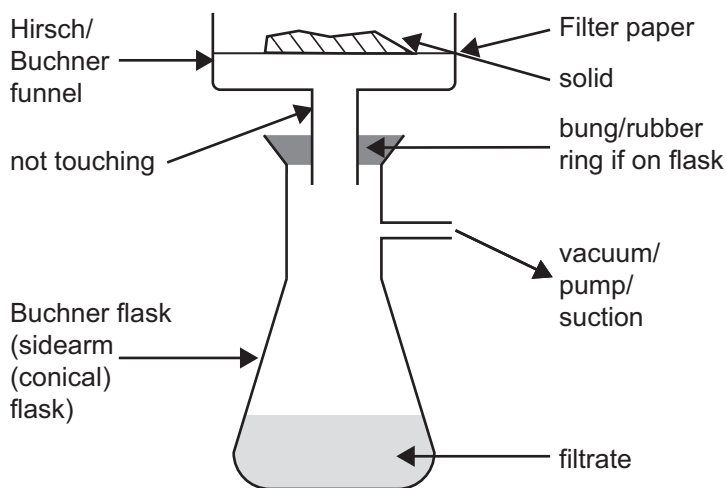
$V = \frac{0.01}{0.06} = 0.167\text{ dm}^3$

$= 167\text{ cm}^3$

[4]

80% yield needs $\frac{100}{80} \times 167 = 210\text{ cm}^3$

(c)



[4]

(d) ethanol mixes with the impurities [1] removes them/
which are not dissolved in water [1] [2]

or

sodium thiosulfate does not dissolve in ethanol/does dissolve in
water [1] to dry the sodium thiosulfate

(e) (i) $2\text{Na}_2\text{S}_2\text{O}_3 + \text{I}_2 \rightarrow 2\text{NaI} + \text{Na}_2\text{S}_4\text{O}_6$ [2]

(ii) 25 cm^3 of $0.1\text{ mol dm}^{-3} \text{ I}_2 = 25 \times 10^{-3} \times 10^{-1}\text{ mol}$
 $= 2.5 \times 10^{-3}\text{ mol}$

$1\text{ mol I}_2 \equiv 2\text{ mol Na}_2\text{S}_2\text{O}_3 \therefore 5.0 \times 10^{-3}\text{ mol}$

$\text{Na}_2\text{S}_2\text{O}_3 = 46 + 64 + 48 = 158$

$\therefore 5.0 \times 10^{-3} \times 158\text{g} = 0.79\text{g}$

$\% \text{ purity} = \frac{0.79}{1.2} \times 100 = 65.8\%$

[3]

Quality of written communication:

- 2 marks The candidate expresses ideas clearly and fluently through well-linked sentences and paragraphs. Arguments are generally relevant and well-structured. There are few errors of grammar, punctuation and spelling.

- 1 mark The candidate expresses ideas clearly, if not always fluently. Arguments may sometimes stray from the point. There may be some errors of grammar, punctuation and spelling, but not such as to suggest a weakness in these areas.

- 0 marks The candidate expresses ideas satisfactorily, but without precision. Arguments may be of doubtful relevance or obscurely presented. Errors in grammar, punctuation and spelling are sufficiently intrusive to disrupt the understanding of the passage.

[2]

20

Total

70