

**OXFORD CAMBRIDGE AND RSA EXAMINATIONS
AS GCE**

F321/01

CHEMISTRY A

Atoms, Bonds and Groups

FRIDAY 27 MAY 2016: Morning

**DURATION: 1 hour
plus your additional time allowance**

MODIFIED ENLARGED

Candidate forename		Candidate surname	
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Centre number						Candidate number				
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Candidates answer on the Question Paper.

OCR SUPPLIED MATERIALS:

Data Sheet for Chemistry A (inserted)

OTHER MATERIALS REQUIRED:

Scientific calculator

READ INSTRUCTIONS OVERLEAF



INSTRUCTIONS TO CANDIDATES

The Insert will be found inside this document.

Write your name, centre number and candidate number in the boxes on the first page. Please write clearly and in capital letters.

Use black ink. HB pencil may be used for graphs and diagrams only.

Answer ALL the questions.

Read each question carefully. Make sure you know what you have to do before starting your answer.

Write your answer to each question in the space provided. If additional answer space is required, you should use the lined page(s) at the end of this booklet. The question number(s) must be clearly shown.

INFORMATION FOR CANDIDATES

The number of marks is given in brackets [] at the end of each question or part question.



Where you see this icon you will be awarded marks for the quality of written communication in your answer.

This means for example you should:

ensure that text is legible and that spelling, punctuation and grammar are accurate so that meaning is clear;

organise information clearly and coherently, using specialist vocabulary when appropriate.

You may use a scientific calculator.

A copy of the Data Sheet for Chemistry A is provided as an insert with this question paper.

You are advised to show all the steps in any calculations.

The total number of marks for this paper is 60.

Any blank pages are indicated.

Answer ALL the questions.

1 Nitrogen is the most common gas in the atmosphere.

(a) Atoms of nitrogen consist of protons, neutrons and electrons.

Complete the table below.

PARTICLE	RELATIVE MASS	RELATIVE CHARGE	POSITION WITHIN THE ATOM
Proton			
Neutron			
Electron			shell

[2]

(b) The electrons in the second shell of a nitrogen atom are found in an s-orbital and three p-orbitals.

(i) State, in words, the 3D shape of an s-orbital and a p-orbital.

s-orbital _____

p-orbital _____

[1]

- (ii) Describe the relative energies of the 2s orbital and EACH of the three 2p orbitals in a nitrogen atom.

[2]

- (c) Draw a 'dot-and-cross' diagram to show the bonding in a nitrogen molecule.

Show OUTER electrons only.

[1]

- (d) Calculate the amount, in mol, of nitrogen **ATOMS** in 5.117×10^{20} nitrogen **MOLECULES**.

Give your answer in standard form.

amount of nitrogen atoms = _____ mol [2]

- (e) N_2O_3 is an unstable oxide of nitrogen that decomposes in a redox reaction.



- (i) State the oxidation number of nitrogen in each oxide in the table below.

OXIDE	OXIDATION NUMBER OF NITROGEN
N_2O_3	
NO	
NO_2	

[1]

(ii) Name this type of redox reaction.

In your answer you should use appropriate technical terms spelled correctly.

_____ [1]

(f) N_2O_3 reacts with water to form an acid as the only product. This reaction is NOT a redox reaction. The empirical formula of the acid formed is the same as the molecular formula.

(i) State what is meant by the term ‘molecular formula’.

_____ [1]

(ii) Suggest the empirical formula of the acid formed.

empirical formula of acid = _____ [1]

[TOTAL: 12]

2 This question is about halogens.

(a) Solid chlorine and solid bromine have a similar structure.

Name this structure.

_____ [1]

(b) The intermolecular attractions in halogens are van der Waals' forces.

(i) Explain how van der Waals' forces arise between halogen molecules.

_____ [3]

- (ii) The boiling points of chlorine and bromine are shown in the table.

HALOGEN	BOILING POINT / °C
chlorine	−34
bromine	59

Explain why bromine has a higher boiling point than chlorine.

[2]

- (c) A student carries out test-tube experiments to prove the trend in reactivity of halogens.**

The student is provided with the following solutions:

bromine water

aqueous iodine

aqueous barium chloride

aqueous magnesium bromide

aqueous calcium iodide.

Chlorine gas and chlorine water are NOT available.

The student carries out the MINIMUM number of test-tube experiments using these solutions in the presence of cyclohexane (an organic solvent).

State the solutions that need to be added together in order to prove the trend in reactivity of the halogens, using the MINIMUM number of test-tube experiments.

Describe the colour seen in the organic solvent at the end of each test-tube experiment.

Write an ionic equation for ONE reaction that takes place.

[illegible]

[TOTAL: 11]

- 3 The elements of Period 2 and Period 3 of the Periodic Table are shown in TABLE 3.1.

TABLE 3.1

GROUP	1	2	3	4	5	6	7	0
Period 2	Li	Be	B	C	N	O	F	Ne
Period 3	Na	Mg	Al	Si	P	S	Cl	Ar

- (a) The elements in these two periods show a repeating pattern in chemical and physical properties.

What is the name given to this repeating pattern of properties?

_____ [1]

- (b) State the element in TABLE 3.1 with:

the lowest first ionisation energy _____

the lowest fourth ionisation energy _____

the lowest boiling point _____

[3]

- (c) Gallium, atomic number 31, is in Period 4 of the Periodic Table. Gallium is a Group 3 element.

Predict the formula of a gallium ion.

_____ [1]

- (d) The melting points of the Period 3 metals sodium and magnesium are shown below.

METAL	MELTING POINT / °C
sodium	98
magnesium	649

Explain the differences in the melting points of sodium and magnesium, using the model of metallic bonding.

In your answer you should use appropriate technical terms spelled correctly.

[3]

(e) When magnesium nitrate, $\text{Mg}(\text{NO}_3)_2$, is heated, it decomposes as shown.



A student heats 2.966 g of $\text{Mg}(\text{NO}_3)_2$, which decomposes as above.

Calculate the TOTAL volume of gas formed, in cm^3 , at room temperature and pressure, RTP.

total volume of gas formed = _____ cm^3 [3]

(f) Fluorine forms several compounds with sulfur and with oxygen.

(i) Give the FORMULA and the NAME of the compound formed between fluorine and sulfur which has octahedral molecules.

Formula _____

Name _____ [1]

(ii) Fluorine reacts with aqueous sodium hydroxide to form the oxide F_2O . Two other products are also formed. One product is an ionic compound with a relative formula mass of 42.0.

Construct a balanced equation for this reaction.

_____ [2]

- (g) (i) Fluorine is the most electronegative element. Indicate any dipoles on the molecule of F_2O below using partial charges.



[1]

- (ii) Suggest the SHAPE of the F_2O molecule and the F–O–F BOND ANGLE.

Shape _____

Bond angle _____

[1]

- (iii) What is the oxidation number of oxygen in F_2O ?

Include the sign in your answer.

_____ [1]

[TOTAL: 17]

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4 This question is about the chemistry of the metals zinc, magnesium, aluminium and calcium.

(a) Complete the electron configuration of a zinc atom.

1s² _____ [1]

(b) A sample of zinc was found to contain four isotopes with the abundances shown in the table.

ISOTOPE	ABUNDANCE (%)
⁶⁴Zn	49.0
⁶⁶Zn	27.9
⁶⁷Zn	4.3
⁶⁸Zn	18.8

(i) Define the term ‘relative atomic mass’.

_____ **[3]**

(ii) Calculate the relative atomic mass of zinc in this sample.

Give your answer to TWO decimal places.

relative atomic mass of zinc = _____ [2]

- (c) Zinc carbonate, ZnCO_3 , reacts with dilute hydrochloric acid.**

A student reacts a sample of ZnCO_3 with an excess of dilute hydrochloric acid in a test-tube.

- (i) Describe what the student would see during this reaction.**

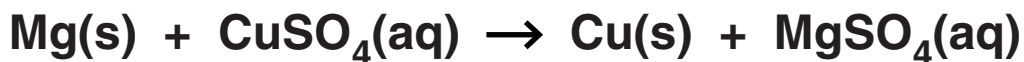
_____ [1]

- (ii) Write the equation for the reaction between ZnCO_3 and dilute hydrochloric acid.**

_____ [1]

(d) Magnesium will undergo redox reactions with aqueous salts of less reactive metals.

(i) A student reacts magnesium with aqueous copper(II) sulfate.



Explain, in terms of NUMBERS of electron transferred, the redox processes taking place in this reaction.

[2]

(ii) The student also noticed that the magnesium started fizzing.

The student thought the fizzing was due to the magnesium reacting with water in the mixture.

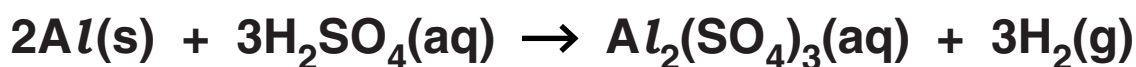
Write the equation for the reaction of magnesium with water.

Include state symbols.

[2]

- (e) A student reacts 35.0 cm^3 of $3.00 \times 10^{-2} \text{ mol dm}^{-3}$ $\text{H}_2\text{SO}_4(\text{aq})$ with an excess of Al .

An equation for this reaction is shown.



Calculate the mass, in g, of $\text{Al}_2(\text{SO}_4)_3$ formed in solution.

Give your answer to THREE significant figures.

Show your working.

mass = _____ g [4]

(f) Compounds of calcium have many uses.

- (i) Identify a compound of calcium that could be used to convert a soil pH from 5.8 to 7.5.**

_____ [1]

- (ii) Calcium phosphide, Ca_3P_2 , is an ionic compound used in rat poison.**

Calcium phosphide can be prepared by reacting calcium metal with phosphorus, P_4 .

Write the equation for the reaction of calcium with phosphorus to form calcium phosphide.

_____ [1]

- (iii) Draw a 'dot-and-cross' diagram to show the bonding in calcium phosphide, Ca_3P_2 .**

Show OUTER electrons only.

[2]

[TOTAL: 20]

END OF QUESTION PAPER

ADDITIONAL ANSWER SPACE

If additional space is required, you should use the following lined page(s). The question number(s) must be clearly shown in the margin(s).

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