



ADVANCED SUBSIDIARY (AS)
General Certificate of Education
2013

Centre Number

71

Candidate Number

MV18

Chemistry

Assessment Unit AS 1

assessing

Basic Concepts in Physical
and Inorganic Chemistry

[AC112]

WEDNESDAY 12 JUNE, AFTERNOON

TIME

1 hour 30 minutes, plus your additional time allowance.

INSTRUCTIONS TO CANDIDATES

Write your Centre Number and Candidate Number in the spaces provided at the top of this page.

Answer **all sixteen** questions.

Answer **all ten** questions in **Section A**. Record your answers by marking the appropriate letter on the answer sheet provided. Use only the spaces numbered 1 to 10. Keep in sequence when answering.

Answer **all six** questions in **Section B**. Write your answers in the spaces provided in this question paper.

INFORMATION FOR CANDIDATES

The total mark for this paper is 100.

Quality of written communication will be assessed in Question

16(b)(iii).

In Section A all questions carry equal marks, i.e. **two** marks for each question.

In Section B the figures in brackets printed at the end of each question indicate the marks awarded to each question or part question.

A Periodic Table of the Elements, containing some data, is included in this question paper.

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(Questions start overleaf)

Section A

For each of the following questions only **one** of the lettered responses (A–D) is correct.

Select the correct response in each case and mark its code letter by connecting the dots as illustrated on the answer sheet.

1 Which one of the following is the formula for nitrogen(I) oxide?

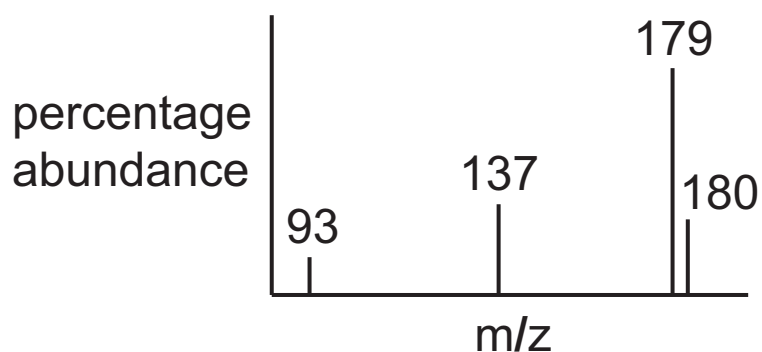
- A NO
- B NO₂
- C N₂O
- D N₂O₄

2 Which one of the following is the number of atoms present in 0.25 moles of C₁₂H₂₂O₁₁?

- A 6.8×10^{24}
- B 1.4×10^{25}
- C 2.7×10^{25}
- D 1.1×10^{26}

- 3 Which one of the following is a molecular covalent substance?
- A CaO
 - B CO
 - C Cr_2O_3
 - D CuO
- 4 A caesium atom differs from a caesium ion because the atom has a greater
- A atomic number.
 - B mass number.
 - C number of electrons.
 - D number of protons.

- 5 Part of the mass spectrum for aspirin is shown below. Which one of the following numbers is the molecular ion peak?



- A 93
B 137
C 179
D 180
- 6 In which one of the following liquids are the van der Waals forces greatest?
- A Argon
B Krypton
C Neon
D Xenon

7 Prozac tablets contain 20 mg of fluoxetine ($C_{17}H_{18}F_3NO$) in each tablet. The number of moles of fluoxetine in a tablet is

- A 6.47×10^{-5}
- B 1.39×10^{-4}
- C 6.47×10^{-2}
- D 1.39×10^{-1}

8 Which one of the following does **not** have a total of 14 electrons?

- A CO
- B Li_2O
- C N_2
- D S^{2-}

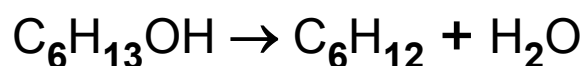
- 9 Successive ionisation energies for elements X and Y are shown below.

Ionisation energy (kJ mol ⁻¹)	1st	2nd	3rd	4th	5th	6th	7th	8th
X	578	1817	2745	11 577	14 842	18 379	23 326	27 465
Y	1314	3388	5301	7469	10 990	13 327	71 330	84 078

Which one of the following is the formula for a compound of X and Y?

- A XY₂
- B X₂Y
- C X₂Y₃
- D X₃Y₂

- 10 Hexan-1-ol can be converted to hex-1-ene as follows:



40.0 g of hexan-1-ol produced 24.7 g of hex-1-ene. Which one of the following is the percentage yield?

- A 24.7%
- B 50.8%
- C 72.0%
- D 75.0%

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(Questions continue overleaf)

Section B

Answer **all six** questions in this section.

11 (a) Atoms consist of protons, neutrons and electrons.

(i) Complete the table below giving the properties of a proton, a neutron and an electron. [3]

	Relative mass	Relative charge
Proton		
Neutron		
Electron		

(ii) Element 116, ununhexium, was added to the Periodic Table in June 2011. Complete the table below. [3]

Atomic number	116
Mass number	
Number of protons	
Number of neutrons	177
Number of electrons	

- (b) Iron is the sixth most abundant element in the universe. It has four isotopes as shown in the table.

Isotope	^{54}Fe	^{56}Fe	^{57}Fe	^{58}Fe
Percentage abundance	5.84	91.76	2.12	0.28

- (i) Explain what is meant by the term **isotope**. [2]

- (ii) Use the table to calculate the relative atomic mass of iron to **two** decimal places. [2]

12 The emission spectrum for atomic hydrogen has been used to provide evidence for discrete electron energy levels in atoms.

(a) Complete the diagram to show the electron transitions associated with the first **two** lines of the hydrogen emission spectrum in the visible region. [2]

$n = 5$ _____

$n = 4$ _____

$n = 3$ _____

$n = 2$ _____

$n = 1$ _____

- (b) The convergence limit of the hydrogen spectrum in the ultraviolet region is at 3.28×10^{15} Hz. Calculate the ionisation energy of hydrogen in kJ mol^{-1} . [3]

- (c) The emission spectra of elements give rise to characteristic flame colours. Complete the table below. [3]

Flame colour	Formula of metal ion
Blue-green	
Crimson	
Green	

13 Wood vinegar, which contains ethanoic acid, is formed when wood is heated. The percentage by mass of ethanoic acid in wood vinegar can be found by titration with standard sodium hydroxide solution.

(a) (i) What is meant by the term **standard solution**? [1]

(ii) Write the equation for the reaction between ethanoic acid and sodium hydroxide. [1]

(b) 25.0 cm^3 of wood vinegar were diluted to 250 cm^3 in a volumetric flask. 25.0 cm^3 of the diluted wood vinegar required 30.3 cm^3 of 0.1 mol dm^{-3} sodium hydroxide solution for neutralisation.

(i) How many moles of sodium hydroxide were required? [1]

(ii) How many moles of ethanoic acid were present in the 25.0 cm^3 of diluted wood vinegar? [1]

(iii) How many moles of ethanoic acid were present in 25.0 cm³ of undiluted wood vinegar? [1]

(iv) What was the mass of ethanoic acid in the 25.0 cm³ of undiluted wood vinegar? [1]

(v) What was the percentage of ethanoic acid by mass in the wood vinegar? Assume that the density of wood vinegar is 1.02 g cm⁻³. [1]

(c) Suggest a suitable indicator for the titration and state the colour change at the end point. [3]

Indicator: _____

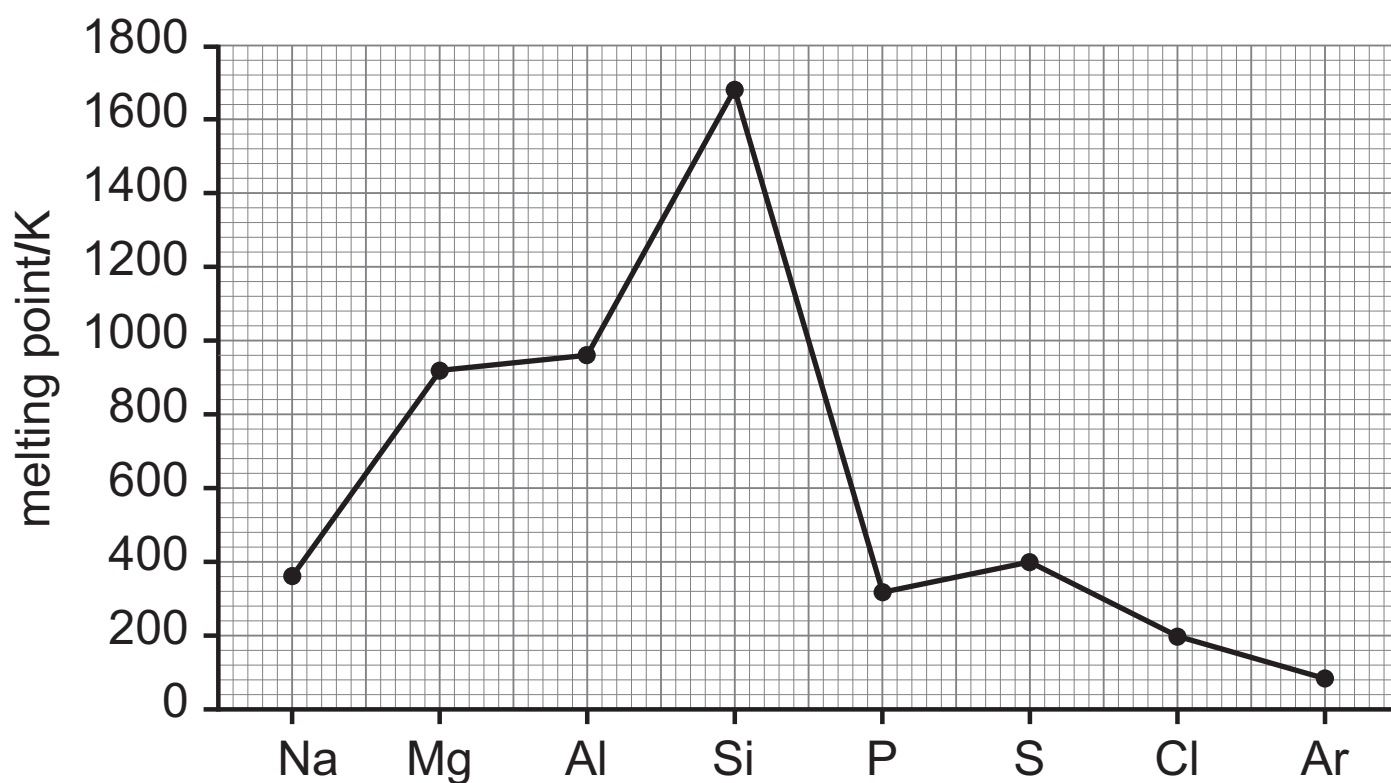
Colour change: from _____

to _____

14 The Third Period from sodium to argon can be used to illustrate trends in the Periodic Table.

(a) In which block of the Periodic Table is argon found?
Explain your answer. [2]

(b) The graph below shows the melting points of the elements in the Third Period.



(i) Explain the rise in melting point from sodium to magnesium. [2]

(ii) Explain why silicon has the highest melting point. [2]

(iii) Explain why the melting point of sulfur, S, is higher than phosphorus, P. [2]

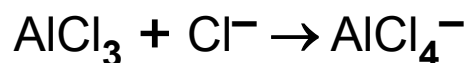
(c) State and explain the trend in atomic radius across the Third Period. [3]

15 Aluminium chloride exists as the molecule AlCl_3 in the vapour state. This molecule contains covalent bonds but does not obey the octet rule.

(a) (i) Explain what is meant by the term **covalent bond**.
[2]

(ii) Explain what is meant by the term **octet rule**. [2]

(b) Aluminium chloride reacts with chloride ions as follows:



(i) Draw dot and cross diagrams, using outer electrons only, to show the bonding in AlCl_3 and AlCl_4^- . [4]

(ii) What type of bond is formed between AlCl_3 and the Cl^- ion? [1]

(iii) Draw and name the shapes of AlCl_3 and AlCl_4^- . [4]

16 The halogens form Group VII of the Periodic Table.

(a) The table below gives some of the physical properties of the halogens.

Element	Atomic radius (nm)	Boiling point (°C)	Electronegativity value	First ionisation energy (kJ mol ⁻¹)
Fluorine	0.133	-187	4.0	1618
Chlorine	0.181	-35	3.0	1256
Bromine	0.196	59	2.8	1143
Iodine	0.219	183	2.0	1009

(i) Explain why the atomic radii of the halogens increase as the Group is descended. [1]

(ii) Explain the trend in the boiling points of the halogens. [2]

(iii) Explain what is meant by the term **electron affinity**.
[1]

(iv) Explain the trend in electronegativity values of the halogens. [2]

(v) Write an equation, including state symbols, for the first ionisation energy of fluorine. [2]

(vi) Explain the trend in the first ionisation energy of the halogens. [2]

(b) Chlorine is used to sterilise water.

(i) Write an equation for the reaction of chlorine with water. [1]

(ii) Using changes in oxidation number explain why this is considered to be a disproportionation reaction. [3]

(iii) Ultraviolet light does not react with water and is equally effective as chlorine at sterilising water. Suggest the advantages and disadvantages of storing and using chlorine to sterilise water. [3]

Quality of written communication [2]

(c) Iodide ions react with a variety of reagents. For each of the following state what you would observe and write a balanced equation for the reaction.

(i) Chlorine gas with aqueous iodide ions. [3]

Observation _____

Equation _____

(ii) A solution containing **excess** Fe^{3+} ions with aqueous iodide ions. [3]

Observation _____

Equation _____

(iii) Silver nitrate solution with aqueous iodide ions. [3]

Observation _____

Equation _____

THIS IS THE END OF THE QUESTION PAPER

For Examiner's use only	
Question Number	Marks
Section A	
1–10	
Section B	
11	
12	
13	
14	
15	
16	
Total Marks	

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