

EXEMPLAR PORTFOLIO WORK UNIT 8				Commentary on Mark Allocation																																
Record of data collected: Task 2b, AO2																																				
<b>Results</b>  EDTA /standard solution of calcium ions <table border="1"> <tr> <td>Burette reading</td><td>1</td><td>2</td><td>3</td></tr> <tr> <td>Final reading cm<sup>3</sup></td><td>3.10</td><td>5.80</td><td>8.60</td></tr> <tr> <td>Start reading cm<sup>3</sup></td><td>0.00</td><td>3.10</td><td>5.80</td></tr> <tr> <td>Volume used in cm<sup>3</sup></td><td>3.10</td><td>2.70</td><td>2.80</td></tr> </table> $\text{Average volume used} = \frac{2.70 + 2.80}{2} = 2.75 \text{ cm}^3$ EDTA /whole milk <table border="1"> <tr> <td>Burette reading</td><td>1</td><td>2</td><td>3</td></tr> <tr> <td>Final reading cm<sup>3</sup></td><td>8.30</td><td>16.40</td><td>24.40</td></tr> <tr> <td>Start reading cm<sup>3</sup></td><td>0.00</td><td>8.30</td><td>16.40</td></tr> <tr> <td>Volume used in cm<sup>3</sup></td><td>8.30</td><td>8.10</td><td>8.00</td></tr> </table> $\text{Average volume used} = \frac{8.00 + 8.10}{2} = 8.05 \text{ cm}^3$ Similar tables given for <ul style="list-style-type: none"> <li>• skimmed milk</li> <li>• semi-skimmed milk</li> <li>• long life semi-skimmed milk</li> <li>• goats milk</li> <li>• soya milk (dairy alternative)</li> <li>• Flora proactive chloesterol lowering milk</li> <li>• tea milk</li> <li>• Yoghurt drink ( munch bunch)</li> </ul> The titration values were used to calculate the mass of calcium in mg. in 25cm <sup>3</sup> milk  Calculation for whole milk  $\text{No. of moles EDTA} = \frac{8.05}{1000} \times 0.1 = 0.000805$ $= \text{No. moles calcium ions}$				Burette reading	1	2	3	Final reading cm <sup>3</sup>	3.10	5.80	8.60	Start reading cm <sup>3</sup>	0.00	3.10	5.80	Volume used in cm <sup>3</sup>	3.10	2.70	2.80	Burette reading	1	2	3	Final reading cm <sup>3</sup>	8.30	16.40	24.40	Start reading cm <sup>3</sup>	0.00	8.30	16.40	Volume used in cm <sup>3</sup>	8.30	8.10	8.00	Recording results AO2a Mark band 2 2 marks.  If an explanation of why the first titration result is ignored, then Mark band 2 3 marks.  For Mark band 3 see comments at end of this task.
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Limited processing and calculation contributes to AO2b and AO2c.  AO2c for calculations is Mark Band 1 2 marks.																																				

Mass of calcium ions =  $0.000805 \times 40 = 0.0322 \text{ g} = 32.2\text{mg}$   
 in  $25\text{cm}^3$   
 Mass of calcium ions in  $100 \text{ cm}^3$   
 Mass of calcium ions in  $100 \text{ cm}^3 = 32.2 \times 4 = 128.8\text{mg}$

Table comparing experimental and published data

Sample	From experiment	Published
Whole milk	128.8mg	119mg
Skimmed milk	133.6mg	124mg
Semi-skimmed	132mg	120mg
Long life semi skimmed	134.4 mg	120mg
Goats milk	132mg	120mg
Soya milk	141.6 mg	148mg
Tea milk	132mg	Not given
Flora	152mg	Not given
Yoghurt milk	114.1 mg	Not given

More variety of recording / processing and selection of data is needed to reach mark band 3

Further results of the investigation - could include

- a survey on daily intake of milk etc
- types of milk people drink etc
- need for calcium
- statistical data on milk intake (type/amount/cohort)

Possibly further experimental work on the contents of milk to include a different type of calculation /method of recording could also be added

Complex calculations etc. could involve error calculations for AO2c.

Data has been related to the objectives of the investigation

AO2 b Mark Band 1  
2 marks have been awarded.

Comments on differences in results needed to obtain Mark band 2 AO2b.

Also Mark band 2 could be awarded if the results of the weighings and the calculation for the reference solution containing  $\text{Ca}^{2+}$ .

**Total Marks AO2 = 6**